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# CBSE Class 12 Chemistry Alcohols Phenols and Ethers MCQs

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# Alcohols Phenols and Ethers Class 12 Chemistry MCQ

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Alcohols Phenols and Ethers in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

### Alcohols Phenols and Ethers MCQ Questions Class 12 Chemistry with Answers

#### Question : Cresol has

(a) Alcoholic – OH

(b) Phenolic – OH

(c) – COOH

(d) – CHO

Answer : B

Question : How many isomers of C<sub>5</sub>H<sub>11</sub>OH will be primary alcohols ?

(a) 5

(b) 4

(c) 2

(d) 3

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#### Question : How many alcohol(s) with molecular formula $C_4H_{10}O$ are chiral in nature?

(a)	1
(b)	2
(c)	З

- (C) 3
- (d) 4

#### **Answer : A**

#### Question : Ethyl alcohol is industrially prepared from ethylene by

- (a) Permanganate oxidation
- (b) Catalytic reduction
- (c) Absorbing in H<sub>2</sub>SO<sub>4</sub> followed by hydrolysis
- (d) All the three

#### Answer : C

#### Question : Sodium salt of benzene sulphonic acid on fusion with caustic soda gives

- (a) Benzene
- (b) Phenol
- (c) Thiophenol
- (d) Benzoic acid

#### Answer: B

#### Question : Acid catalyzed hydration of alkenes except ethene leads to the formation of

- (a) primary alcohol
- (b) secondary or tertiary alcohol
- (c) mixture of primary and secondary alcohols
- (d) mixture of secondary and tertiary alcohols

#### Answer : B

#### Question : Ethyl alcohol can be prepared from Grignard reagent by the reaction of :

(a) HCHO (b) R<sub>2</sub>CO(c) RCN (d) RCOCI

#### **Answer : A**

Question : Isopropyl alcohol is obtained by reacting which of the following alkenes with concentrated H<sub>2</sub>SO<sub>4</sub> followed by boiling with H<sub>2</sub>O?

(a) Ethylene

(b) Propylene

(c) 2-Methylpropene

(d) Isoprene

Answer : B

Question : Alkenes convert into alcohols by

(a) hydrolysis by dil. H<sub>2</sub>SO<sub>4</sub>

- (b) hydration of alkene by alkaline KMnO<sub>4</sub>
- (c) hydrolysis by water vapours and conc. HNO<sub>3</sub>
- (d) hydration of alkene by aqueous KOH

#### Answer : B

#### Question : Which of the following reacts with NaOH to give an alcohol?

- (a) Propene
- (b) Butene
- (c) Ethanal
- (d) Methanal

#### Answer : D

#### Question : Which of the following compounds is oxidised to prepare methyl ethyl ketone?

- (a) 2-Propanol
- (b) I-Butanol
- (c) 2-Butanol
- (d) t-Butyl alcohol

#### Answer : C

#### **Question : The compound HOCH<sub>2</sub> - CH<sub>2</sub>OH is**

- (a) ethane glycol
- (b) ethylene glycol
- (c) ethylidene alcohol
- (d) dimethyl alcohol

#### Answer : B

#### Question : HBr reacts fastest with

(a) 2-Mehtylpropan-1-ol

#### (b) 2-Methylpropene-2-ol

- (c) propan-2-ol
- (d) propan-1-ol

#### Answer : B

**Question : Which of the following is dihydric alcohol ?** 

(a) Glycerol

(b) Ethylene glycol

(c) Catechol

(d) Resorcinol

Answer : B

#### Question : An example of a compound with functional group – O – is :

(a) acetic acid

- (b) methyl alcohol
- (c) diethyl ether
- (d) acetone

#### Answer : C

#### **Question : Butane-2-ol is**

(a) primary alcohol

#### (b) secondary alcohol

(c) tertiary alcohol

(d) aldehyde

#### Answer : D

#### **Question : Cresol has**

(a) Alcoholic – OH

#### (b) Phenolic – OH

(c) – COOH

(d) – CHO

#### Answer : B

#### Question : How many isomers of C<sub>5</sub>H<sub>11</sub>OH will be primary alcohols ?

(a) 5

- (b) 4
- (c) 2
- (d) 3

#### Answer : B

Question : *n*-Propyl alcohol and isopropyl alcohol can be chemically distinguished by which reagent?

(a) PCl<sub>5</sub>

(b) Reduction

(c) Oxidation with potassium dichromate

(d) Ozonolysis

Answer : D

Question : Number of metamers represented by molecular formula C<sub>4</sub>H<sub>10</sub>O is

(a)	4
-----	---

- (b) 3
- (c) 2
- (d) 1

#### Answer : B

#### **Question : Lucas reagent is**

(a) Conc. HCl and anhydrous ZnCl<sub>2</sub>

(b) Conc. HNO<sub>3</sub> and hydrous ZnCl<sub>2</sub>

(c) Conc. HCl and hydrous ZnCl<sub>2</sub>

(d) Conc. HNO<sub>3</sub> and anhydrous ZnCl

#### Answer : A

#### Question : How many alcohol(s) with molecular formula C<sub>4</sub>H<sub>10</sub>O are chiral in nature?

(a) 1			
(b) 2			
(c) 3			
(d) 4			

#### Answer : A

#### Question : The compound which reacts fastest with Lucas reagent at room temperature is

- (a) Butan-1-ol
- (b) Butan-2-ol
- (c) 2-Methyl propan-1-ol
- (d) 2-Methylpropan-2-ol

#### Answer : D

#### Question : IUPAC name of *m*-cresol is \_\_\_\_\_

(a) 2-methylphenol

(b) 3-chlorophenol

(c) 3-methoxyphenol

(d) benzene-1, 3-diol

Answer : A

#### **Question : When phenol is treated with excess bromine water, it gives:**

(a) *m*-Bromophenol

(b) *o*- and *p*-Bromophenol

(c) 2, 4-Dibromophenol

(d) 2, 4, 6-Tribromophenol

#### CBSE Class 12 Chemistry Alcohols Phenols and Ethers MCQs, Multiple Choice Questions for Chemistry

#### **Answer : D**

#### Question : When phenol is heated with CHCl<sub>3</sub> and alcoholic KOH when salicyladehyde is produced. This reaction is known as

(a) Rosenmund's reaction

- (b) Reimer-Tiemann reaction
- (c) Friedel-Crafts reaction
- (d) Sommelet reaction

#### **Answer: B**

#### **Question : On distilling phenol with Zn dust, one gets :**

- (a) Toluene
- (b) Benzaldehyde + ZnO
- (c) ZnO + benzene
- (d) Benzoic acid

#### Answer : C

#### Question : Which of the following is an example of unsymmetrical ether?

- (a)  $C_2H_5OC_2H_5$
- (b)  $C_6H_5OC_6H_5$
- (c)  $C_6H_5OC_2H_5$
- (d) CH<sub>3</sub>OCH

#### **Answer : C**

#### Question : Which of the following will not form phenol or phenoxide ?

- (a)  $C_6H5N_2CI$
- (b)  $C_6H_5SO_3Na$
- (c)  $C_6H_5CI$

#### (d) $C_6H_5CO_2H$



#### **Answer : D**

#### Question : Benzyl alcohol is obtained from benzaldehyde by

(a) Fittig's reaction

(b) Cannizzaro's reaction

(c) Kolbe's reaction

(d) Wurtz's reaction

#### **Answer : B**

#### Question : In the reduction R - CHO+ $H_2 \! \rightarrow \! RCH_2OH$ the catalyst used is :

(a) Ni

(b) Pd

(c) Pt

#### (d) Any of these

#### Answer : D

#### Question : Ethylene reacts with Baeyer's reagent to give

(a) ethane

(b) ethyl alcohol

- (c) ethylene glycol
- (d) None of these

#### Answer : C

#### Question : Ethyl alcohol is industrially prepared from ethylene by

- (a) Permanganate oxidation
- (b) Catalytic reduction
- (c) Absorbing in H<sub>2</sub>SO<sub>4</sub> followed by hydrolysis
- (d) All the three

#### Answer : C

#### Question : Sodium salt of benzene sulphonic acid on fusion with caustic soda gives

(a) Benzene

#### (b) Phenol

- (c) Thiophenol
- (d) Benzoic acid

#### Answer: B

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(a) primary alcohol

(b) secondary or tertiary alcohol

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(d) mixture of secondary and tertiary alcohols

#### Answer : C

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#### (b) R<sub>2</sub>CO

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Question : Isopropyl alcohol is obtained by reacting which of the following alkenes with concentrated H<sub>2</sub>SO<sub>4</sub> followed by boiling with H<sub>2</sub>O?

(a) Ethylene

(b) Propylene

(c) 2-Methylpropene

(d) Isoprene

**Answer : D** 

#### **Question : Alkenes convert into alcohols by**

(a) hydrolysis by dil. H<sub>2</sub>SO<sub>4</sub>

(b) hydration of alkene by alkaline KMnO<sub>4</sub>

(c) hydrolysis by water vapours and conc. HNO<sub>3</sub>

(d) hydration of alkene by aqueous KOH

#### **Answer: B**

#### Question : Which of the following reacts with NaOH to give an alcohol?

(a) Propene

(b) Butene

- (c) Ethanal
- (d) Methanal

#### **Answer : D**

#### Question : By which of the following methods alcohol can be prepared in excellent yield?

- (a) From alkenes
- (b) By hydroboration-oxidation
- (c) From carbonyl compounds

(d) From Grignard reagent

**Answer : D** 

Question : Which of the following are used to convert RCHO into RCH2OH?

(i) H<sub>2</sub>/Pd

(ii) LiAlH<sub>4</sub>

(iii) NaBH<sub>4</sub>

(iv) Reaction with RMgX followed by hydrolysis

#### (b) (i), (ii) and (iii)

(c) (ii), (iii) and (iv)

(d) (i) and (iii)

#### **Answer : D**

Question : Commercially carboxylic acids are reduced to alcohols by converting them to the \_\_\_\_\_.

(a) esters

(b) aldehydes

(c) ketones

(d) amines

#### **Answer: A**

Question : Phenols do not react with one of the following :

(a) Alkali metals

(b) Sodium hydroxide

(c) Potassium hydroxide

(d) Sodium bi-carbonate

#### **Answer : D**

#### **Question : Dehydration of 2-butanol yields**

(a) 1-butene

(b) 2-butene

(c) 2-butyne

(d) Both (a) and (b)

#### **Answer : D**

Question : The correct order of boiling points for primary (1°), secondary (2°) and tertiary alcohol (3°) is

(a) 1° > 2° > 3°

(b) 3° > 2° > 1°

(c) 2° > 1° > 3°

(d) 2° > 3° > 1°

**Answer : A** 

#### Question : Alcohols of low molecular weight are

(a) soluble in water

(b) soluble in water on heating

- (c) insoluble in water
- (d) insoluble in all solvents

#### **Answer : A**

#### Question : Which of the following has lowest boiling point ?

- (a) *p*-Nitrophenol
- (b) *m*-Nitrophenol
- (c) *o*-Nitrophenol
- (d) Phenol

#### Answer : C

#### **Question : Which statement is not correct about alcohol?**

- (a) Molecular weight of alcohol is higher than water
- (b) Alcohol of less no. of carbon atoms is less soluble in water than alcohol of more no. of carbon atoms
- (c) Alcohol evaporates quickly
- (d) All of the above

#### Answer : B

#### Question : Which one of the following alcohols is least soluble in water?

- (a) CH<sub>3</sub>OH
- (b) C<sub>3</sub>H<sub>7</sub>OH
- (c)  $C_4H_9OH$
- (d) C<sub>10</sub>H<sub>21</sub>OH

#### Answer : D

#### Question : Methanol and ethanol are miscible in water due to

(a) covalent character

- (b) hydrogen bonding character
- (c) oxygen bonding character

#### Answer : B

#### Question : If ethanol dissolves in water, then which of the following would be observed

(a) absorption of heat and contraction in volume

(b) emission of heat and contraction in volume

(c) absorption of heat and increase in volume

(d) emission of heat and increase in volume

#### Answer : B

#### Question : Which of the following is correct?

- (a) On reduction of any aldehyde, secondary alcohol is formed
- (b) Reaction of vegetable oil with H2SO4 gives glycerine
- (c) Sucrose on reaction with NaCl gives invert sugar
- (d) Alcoholic iodine gives iodoform with NaOH

#### **Answer : D**

#### Question : Which of the following is not true in case of reaction with heated copper at 300°C?

- (a) Phenol  $\rightarrow$  Benzyl alcohol
- (b) Secondary alcohol  $\rightarrow$  Ketone
- (c) Primary alcohol  $\rightarrow$  Aldehyde
- (d) Tertiary alcohol  $\rightarrow$  Olefin

#### Answer : A

#### Question : Phenol is more acidic than alcohol because

- (a) phenol is more stable than water
- (b) phenol is aromatic and alcohol is aliphatic
- (c) phenoxide ion is resonance stabilised
- (d) None of these

#### Answer : C

#### Question : Acidity of phenol is due to

- (a) hydrogen bonding
- (b) phenolic group
- (c) benzene ring
- (d) resonance stabilisation of its anion

#### **Answer : D**

#### Question : Lucas test is done to differentiate between

(a) alcohol and ketone

(b) alcohol and aromatic ketones

(c) 1°, 2° and 3° alcohols

Answer : C

**Question : The ionization constant of phenol is higher than that of ethanol because :** 

(a) phenoxide ion is bulkier than ethoxide

(b) phenoxide ion is stronger base than ethoxide

#### (c) phenoxide ion is stabilized through delocalization

(d) phenoxide ion is less stable than ethoxide

#### Answer : C

#### Question : Which one of the following on oxidation gives a ketone ?

- (a) Primary alcohol
- (b) Secondary alcohol
- (c) Tertiary alcohol
- (d) All of these

#### **Answer : B**

#### Question : Primary and secondary alcohols on action of reduced copper give

- (a) Aldehydes and ketones respectively
- (b) Ketones and aldehydes respectively
- (c) Only aldehydes
- (d) Only ketones

#### **Answer : A**

#### Question : When ethyl alcohol reacts with acetic acid, the products formed are

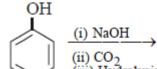
(a) Sodium ethoxide + hydrogen

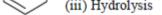
#### (b) Ethyl acetate + water

- (c) Ethyl acetate + soap
- (d) Ethyl alcohol + water
- (d) None of these

#### **Answer : B**

#### Question : The product obtained from the reaction is





(a) Benzene(b) Toluene(c) Salicylic acid(d) Benzoic acid

#### Answer : C

#### **Question : Picric acid is**

(a) Trinitrophenol(b) Trinitrotoluene(c) Trinitrobenzene(d) Tribromobenzene

#### **Answer : A**

#### **Question : Lucas reagent is**

(a) anhy. AlCl<sub>3</sub> + conc. HCl
(b) anhy. AlC<sub>13</sub> + conc.HNO<sub>3</sub>
(c) anhy. ZnCl<sub>2</sub>
(d) anhy. ZnCl<sub>2</sub> + conc. HCl

#### Answer : D

#### Question : Lucas test is used for the detection of

(a) alcohols (b) alkyl halides(c) phenols (d) aldehydes

#### **Answer : A**

#### Question : Intermolecular hydrogen bonding is strongest in

(a) Methylamine(b) Phenol(c) Formaldehyde(d) Methanol

#### **Answer : A**

#### Question : Among the following the one which reacts most readily with ethanol is

(a) p-nitrobenzyl bromide

(b) p-chlorobenzyl bromide

(c) p-methoxybenzyl bromide

(d) p-methylbenzyl bromide

#### Answer : C

#### Question : The compound HOCH<sub>2</sub> - CH<sub>2</sub>OH is

(a) ethane glycol(b) ethylene glycol(c) ethylidene alcohol (d) dimethyl alcohol

#### Answer : B

#### Question : Which of the following is dihydric alcohol ?

(a) Glycerol(b) Ethylene glycol(c) Catechol(d) Resorcinol

#### **Answer: B**

Question : An example of a compound with functional group – O – is :

(a) acetic acid(b) methyl alcohol(c) diethyl ether (d) acetone

Answer : C

#### **Question : Butane-2-ol is**

(a) primary alcohol (b) secondary alcohol(c) tertiary alcohol (d) aldehyde

#### Answer : B

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# **Aldehydes Ketones and Carboxylic Acids Class 12 Chemistry MCQ**

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Aldehydes Ketones and Carboxylic Acids in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

### Aldehydes Ketones and Carboxylic Acids MCQ Questions **Class 12 Chemistry with Answers**

Question : Which of the following forces explain the boiling point of aldehydes and ketones?

(a) Hydrogen bonding (b) van der Waal's forces (c) Dipole-dipole attraction (d) None of these

**Answer : C** 

#### **Question : Which is highly soluble in water?**

(a) Methanal

(b) Propanal

(c) Propanone

(d) Butanone

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#### Answer : A

# Question : Propanal and propanone, both have same molecular formula(C<sub>3</sub>H<sub>6</sub>O), what do you expect about their boiling points?

- (a) Both have same boiling point
- (b) Boiling point of propanal is higher than the boiling point of propanone.
- (c) Boiling point of propanal is lower than the boiling point of propanone
- (d) Nothing can be predicted

#### Answer : C

#### Question : Less reactivity of ketone is due to

(a) + I inductive effect decrease positive charge on carbonyl carbon atom
(b) steric effect of two bulky alkyl groups
(c) sp<sup>2</sup> hybridised carbon atom of carbonyl carbon atom
(d) Both (a) and (b)

#### Answer : D

#### **Question : Acetaldehyde reacts with**

(a) Electrophiles only
(b) Nucleophiles only
(c) Free radicals only
(d) Both electrophiles and nucleophiles

#### **Answer : B**

#### Question : Carbonyl compounds undergo nucleophilic addition because of

(a) electronegativity difference of carbon and oxygen atoms(b) electromeric effect(c) more stable anion with negative charge on oxygen atom and less stable carbonium ion(d) None of the above

#### Answer : C

#### **Question : Which of the following statement is false ?**

- (a) Cannizzaro reaction is given by aldehydes in presence of alkali
- (b) Aldol condensation is given by aldehydes in presence of alkali
- (c) Aldol condensation is given by aldehydes and ketones in presence of acids
- (d) None of the above

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#### Question : If formaldehyde and KOH are heated, then we get

(a) methane

(b) methyl alcohol

(c) ethyl formate

(d) acetylene

Answer : B

Question : Aldehydes can be oxidised by :

- (a) Tollen's reagent
- (b) Fehling solution
- (c) Benedict solution
- (d) All the above

#### **Answer : D**

#### **Question : 2-pentanone and 3-pentanone can be distinguished by :**

- (a) Cannizaro's reaction
- (b) Aldol condensation
- (c) lodoform reaction
- (d) Clemmensen's reduction

#### Answer : C

#### Question : Which of the following compounds will give but anone on oxidation with alkaline ${\sf KMnO}_4$

#### solution?

- (a) Butan-1-ol
- (b) Butan-2-ol
- (c) Both of these
- (d) None of these

#### Answer : B

#### Question : Schiff's reagent gives pink colour with

(a) acetaldehyde

(b) acetone

(c) acetic acid

(d) methyl acetate

Answer : A

#### Question : The IUPAC name of CH<sub>3</sub>COCH(CH<sub>3</sub>)<sub>2</sub> is

(a) 2-methyl-3-butanone
(b) 4-methylisopropyl ketone
(c) 3-methyl-2-butanone
(d) Isopropylmethyl ketone
Answer : C

#### Question : IUPAC name of ethyl isopropyl ketone is

(a) 4-methyl pent-3-one(b) 2-methyl pent-3-one(c) 4-methyl pent-2-one(d) 2-methyl pent-2-one

#### Answer : B

#### Question : In > C = O group sigma bond is formed by

(a) sp<sub>2</sub>-p-overlapping
(b) sp3-p-overlapping
(c) sp-p-overlapping
(d) s-p-overlapping
Answer : A

Question : The π-bond in carbonyl group is formed by(a) s-s-overlapping(b) p-p-overlapping(c) s-p-overlapping(d) p-d-overlappingAnswer : B

#### Question : Which of the following contain an aldehyde?

- (a) Vanilla beans
- (b) Meadow sweet
- (c) Cinnamon
- (d) All of these

#### Answer : D

#### Question : Which of the following have pleasant smell?

- (a) Methanal
- (b) Propanal
- (c) Ethanal
- (d) Hexanal

#### Answer : D

#### Question : Which one of the following can be oxidised to the corresponding carbonyl compound?

- (a) 2-hydroxy-propane(b) Ortho-nitrophenol(c) Phenol
- (d) 2-methyl-2 hydroxy-propane

#### Answer : A

#### Question : Which one of the following on oxidation gives a ketone ?

- (a) Primary alcohol
- (b) Secondary alcohol
- (c) Tertiary alcohol
- (d) All of these

#### Answer : B

#### Question : What is formed when a primary alcohol undergoes catalytic ehydrogenation ?

- (a) Aldehyde
- (b) Ketone
- (c) Alkene
- (d) Acid

#### **Answer : A**

#### Question : Primary and secondary alcohols on action of reduced copper give

- (a) Aldehydes and ketones respectively
- (b) Ketones and aldehydes respectively
- (c) Only aldehydes
- (d) Only ketones

#### Answer : A

#### Question : Which of the following does not represent the natural source of the corresponding acids

- ?
- (a) Formic acid : Red ant
- (b) Acetic acid : Vinegar
- (c) Butyric acid : Rancid butter
- (d) Isobutyric acid : Automobile exhausts
- . .

#### Answer : D

#### Question : The catalyst used in Rosenmund's reduction is

(a) HgSO<sub>4</sub>(b) Pd/BaSO<sub>4</sub>

(c) anhydrous AlCl3

(d) anhydrous ZnCl<sub>2</sub>

#### Answer : B

#### Question : Vinegar is a solution of acetic acid which is :

(a) 15 – 20%

(b) 20 –25%

(c) 6 – 8%

(d) 2 – 4%

#### Answer : C

#### Question : Benzaldehyde can be prepared by oxidation of toluene by

(a) Acidic KMnO<sub>4</sub>
(b) K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> / H+
(c) CrO<sub>2</sub>Cl<sub>2</sub>
(d) All of these

#### Answer : C

#### Question : The oxidation of toluene to benzaldehyde by chromyl chloride is called

(a) Rosenmund reaction(b) Wurtz reaction

(c) Etard reaction

(d) Fittig reaction

Answer : C

#### Question : An aldehyde group can be present

- (a) in between carbon chain
- (b) at any position in carbon atom
- (c) only at the end of carbon chain
- (d) at the second carbon atom of the carbon chain

#### Answer : C

#### Question : Methyl cyanide can be converted into acetic acid by one of the following reactions.

- (a) Reduction
- (b) Hydrolysis
- (c) Electrolysis
- (d) Decarboxylation

#### Answer : B

#### Question : Which of the following is not used in the preparation of ketone?

- (a) Oxidation of secondary alcohols
- (b) Dehydrogenation of 2° alcohol
- (c) Pyrolysis of calcium acetate
- (d) Acid hydrolysis of alkyl cyanide
- **Answer : D**

#### Question : Which reaction is used for detecting the presence of carbonyl group?

- (a) Reaction with hydrazine
- (b) Reaction with phenyl hydrazine
- (c) Reaction with hydroxylamine
- (d) All of the above

#### **Answer : D**

#### Question : The difference between aldol condensation and Cannizzaro's reaction is that:

- (a) the former takes place in the presence of -H-atom.
- (b) the former takes place in the absence of -H-atom.

(c) the former takes place in the presence of -H-atom.

(d) none of the above

#### Answer : A

#### Question : The product obtained by the reaction of an aldehyde and hydroxylamine is

(a) hydrazone

(b) aldoxime

(c) primary amine

(d) alcohol

Answer : B

#### **Question : Ketone upon treatment with Grignard Reagent gives** (a) primary alcohol

(b) secondary alcohol

(c) tertiary alcohol(d) aldehydeAnswer : C

# Question : Two compounds benzyl alcohol and benzoic acid are formed from this compound, when this compound is heated in the presence of conc.NaOH, this compound is.

(a) Benzaldehyde(b) Benzylalcohol(c) Acetophenone

#### (d) Benzophenone

#### Answer : A

#### Question : Benzophenone can be converted into benzene by using

(a) fused alkali

(b) anhydrous AlCl<sub>3</sub>

(c) sodium amalgam in water

(d) acidified dichromate

#### Answer : A

#### Question : Which of the following forces explain the boiling point of aldehydes and ketones?

(a) Hydrogen bonding

(b) van der Waal's forces

(c) Dipole-dipole attraction

(d) None of these

#### Answer : C

#### **Question : Which is highly soluble in water?**

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#### Question : Which of the following statement is false ?

- (a) Cannizzaro reaction is given by aldehydes in presence of alkali
- (b) Aldol condensation is given by aldehydes in presence of alkali
- (c) Aldol condensation is given by aldehydes and ketones in presence of acids

#### (d) None of the above

#### Question : If formaldehyde and KOH are heated, then we get

- (a) methane
- (b) methyl alcohol
- (c) ethyl formate
- (d) acetylene

#### Answer : B

#### Question : The reagent which can be used to distinguish acetophenone from benzophenone is

- (a) 2,4- dinitrophenylhydrazine
- (b) aqueous solution of NaHSO<sub>3</sub>
- (c) benedict reagent
- (d) I2and Na<sub>2</sub>CO<sub>3</sub>

#### **Question : Wolf-Kishner reduction is**

- (a) reduction of carbonyl compound into alcohol(b) reduction of carbonyl compound into alkene(c) reduction of carboxyl compound into alkane
- (d) reduction of nitro compound into aniline

#### Answer : C

#### Question : Acetone reacts with iodine (I2) to form iodoform in the presence of

(a) CaCO<sub>3</sub>
(b) NaOH
(c) KOH
(d) MgCO<sub>3</sub>
Answer : B

#### Question : Imine derivatives of aldehyde and ketone is called as

- (a) Schiff's reagent
- (b) Fehling's reagent
- (c) Schiff's base
- (d) Schiff's acid
- Answer : C

#### **Question : Tollen's reagent is**

- (a) ammonical CuSO<sub>4</sub>
  (b) ammonical AgNO<sub>3</sub>
  (c) alkaline solution containing complex of copper nitrate
  (d) none of these
  Answer : B

#### Question : lodoform test is not given by

(a) 2-Pentanone

(b) Ethanol

(c) Ethanal

(d) 3-Pentanone

Answer : D

Question : Phenylmethyl ketone can be converted into ethylbenzene in one step by which of the following reagents?

(a) LiAlH4

(b) Zn-Hg/HCl

(c) NaBH4

(d) CH3MgI

Answer : B

#### Question : When acetaldehyde is heated with Fehling's solution it gives a precipitate of

(a) Cu
(b) CuO
(c) Cu<sub>2</sub>O
(d) Cu(OH)<sub>2</sub>
Answer : C

#### Question : The reagent which does not react with both, acetone and benzaldehyde.

- (a) Sodium hydrogensulphite
- (b) Phenyl hydrazine
- (c) Fehling's solution
- (d) Grignard reagent

#### Answer : C

# Question : Which of the following pairs of compounds will undergo aldol and Cannizzaro reaction respectively ?

(i) acetone; benzaldehyde
(ii) acetaldehyde; butan-2-one
(iii) propanone; formaldehyde.
(iv) cyclopentanone, benzaldehyde
(a) (i), (ii) and (iii) (b) (ii) and (iii)
(c) (ii), (iii) and (iv) (d) (iii) and (iv)
Answer : A

#### Question : Which of the following compound will show positive silver mirror test ?

(a) HCOOH (b)  $CH_3$  (CHOH)<sub>3</sub>CHO (c)  $CH_3CO(CHOH)CH_3$ (d) Both (a) and (b) Answer : D

#### Question : Aldehydes and ketones are distinguished by which of the following test ?

(a) Lucas test
(b) Tollen's test
(c) KMnO<sub>4</sub> solution (Baeyer's test)
(d) None of these
Answer : B

#### Question : Aldehydes and ketones are generally reduced by :

(a) Clemmensen reduction
(b) H<sub>2</sub>S
(c) H2 / Ni
(d) None of these
Answer : A

#### Question : In which reaction, > C = O can be reduced to $> CH_2$ ?

- (a) Wolf-Kishner reaction

(b) Reimer-Tiemann reaction(c) Wurtz reaction(d) None of theseAnswer : A

Question : Which of the following contain an aldehyde?

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- (b) Ketone
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- (a) Hydrogen bonding
- (b) van der Waal's forces
- (c) Dipole-dipole attraction
- (d) None of these

#### Answer : C

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# **Amines Class 12 Chemistry MCQ**

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Amines in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

### Amines MCQ Questions Class 12 Chemistry with Answers

Question : The total number of electrons around the nitrogen atom in amines are

(a) 8 (b) 7 (c) 4 (d) 3

#### **Answer : A**

#### Question : Amines play an important role in the survival of life. Naturally they are found in

(a) proteins(b) vitamins(c) alkaloids(d) All of these

Answer : D

Question : Propionamide on Hofmann degradation gives -

(a) methyl amine (b) ethyl amine(c) propyl amine (d) ethyl cyanide



<u>VBQs</u>

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≺=

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#### Answer : B

#### **Question : Secondary amines could be prepared by**

(a) reduction of nitriles

- (b) Hofmann bromamide reaction
- (c) reduction of amides
- (d) reduction of isonitriles

#### Answer : D

#### Question : Gabriel's phthalimide synthesis is used for the preparation of

- (a) Primary aromatic amines
- (b) Secondary amines
- (c) Primary aliphatic amines
- (d) Tertiary amines

#### Answer : C

#### Question : Treatment of ammonia with excess of ethyl iodide will yield

(a) diethylamine(b) ethylamine(c) triethylamine(d) tetraethylammonium iodide

#### **Answer : D**

#### Question : The reduction of nitro compounds is most preferred in the presence of

(a) Pd/H<sub>2</sub> in ethanol (b) Sn + HCl(c) finely divided Ni (d) iron scrap and HCl.

#### Answer : D

#### Question : An alkyl or benzyl halide on reaction with an ethanolic solution of ammonia undergoes

(a) electrophilic substitution reaction

- (b) nucleophilic substitution reaction.
- (c) free radical mechanism.
- (d) nucleophilic addition reaction.

**Answer: B** 

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**Question : Which of the following will give primary amine only ?** (i) ammonia + propylchloride (ii) potassium pthalimide + ethylchloride

(iii) potassium pthalimide + chlorobenzene

(a) (i) and (ii) (b) (i) and (iii) (c) (ii) and (iii) (d) (i), (ii) and (iii)

#### Answer : A

#### **Question : Amines have**

(a) Garlic odour(b) Fishy odour(c) Jasmine odour (d) Bitter almonds odour

#### **Question : Amines behave as**

(a) lewis acids (b) lewis bases (c) aprotic acids (d) amphoteric compounds

#### **Answer: B**

#### Question : The basic character of amines is due to

- (a) presence of nitrogen atom
- (b) lone pair of electrons on nitrogen atom
- (c) tetrahedral structure
- (d) high electronegativity of nitrogen

#### **Answer: B**

#### Question : Aliphatic amines are.....basic than NH<sub>3</sub> but aromatic amines are.....basic than NH<sub>3</sub>.

(a) more, less (b) less, more (c) both (a) and (b) (d) None of these

#### Answer: A

#### Question : Substitution of one alkyl group by replacing hydrogen of primary amines

- (a) increases the base strength
- (b) decreases the base strength
- (c) remains the same
- (d) None of the above

#### **Answer: A**

#### Question : Which of the following is not characteristic of amines?

- (a) They smell like ammonia
- (b) They are inflammable in air
- (c) They show the property of hydrogen bonding
- (d) They are amphoteric in nature

#### Answer : D

#### **Question : Which statement is not true among the following?**

(a) Amines are bases

(b) They turn red litmus blue

(c) Trimethyl amine is less basic than dimethyl amine (d) Amines yield alcohols on aqueous hydrolysis.

#### **Answer : D**

#### **Question : Aniline is used**

(a) in crimping of wool (b) in dyeing industry (c) in making of glue (d) in fast drying vanish

#### **Answer: B**

#### Question : Which of the following statements about primary amines is False' ?

(a) Alkyl amines are stronger bases than aryl amines

- (b) Alkyl amines react with nitrous acid to produce alcohols
- (c) Aryl amines react with nitrous acid to produce phenols
- (d) Alkyl amines are stronger bases than ammonia

#### Answer : C

#### **Question : Mark the correct statement**

- (a) Methylamine is slightly acidic
- (b) Methylamine is less basic than ammonia
- (c) Methylamine is a stronger base than ammonia
- (d) Methylamine forms salts with alkalies.

#### Answer : C

#### Question : For carbylamine reaction, we need hot alcoholic KOH and

- (a) any primary amine and chloroform
- (b) chloroform and silver powder
- (c) a primary amine and an alkyl halide
- (d) a monoalkylamine and trichloromethane.

#### **Answer : A**

#### **Question : A secondary amine is**

- (a) a compound with two carbon atoms and an  $-NH_2$  group.
- (b) a compound containing two  $-NH_2$  groups.

#### (c) a compound in which hydrogens of NH<sub>3</sub> have been replaced by two alkyl groups.

(d) a compound with an –NH2 group on carbon atom in number two position.

#### Question : The general formula of quaternary ammonium compound is

(a) R–NH<sub>2</sub>

(b) R<sub>3</sub>N

#### (c) R<sub>4</sub>N+ X–

(d)  $NH_4X$ 

(b) 7

(c) 4

(d) 3

#### Question : (CH<sub>3</sub>)<sub>2</sub>CHNH<sub>2</sub> is reacted with excess acetic anhydride, the compound formed is

(a) (CH<sub>3</sub>)<sub>2</sub>CHNCOCH<sub>3</sub>

(b) (CH<sub>3</sub>)<sub>2</sub>CN(COCH<sub>3</sub>)<sub>2</sub>

(c) (CH<sub>3</sub>)<sub>2</sub>CHOH

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(d) (CH<sub>3</sub>)<sub>2</sub>CN(COOCH<sub>3</sub>)<sub>2</sub>

#### Question : The number of primary amines of formula $C_4H_{11}N$ is :

1-1	1
(a)	_ I
(u)	

- (b) 3
- (c) 4
- (d) 2

#### **Question : Ethylamine reacts with HNO<sub>2</sub> giving :**

#### (a) $C_2H_5OH$

(b)  $C_2H_5NO_2$ 

(c) NH<sub>3</sub>

(d) C<sub>2</sub>H<sub>6</sub>

#### Question : The correct IUPAC name for CH<sub>2</sub> = CHCH<sub>2</sub>NHCH<sub>3</sub> is

- (a) Allylmethylamine
- (b) 2-amino-4-pentene
- (c) 4-aminopent-1-ene
- (d) N-methylprop-2-en-1-amine

#### Question : Amines play an important role in the survival of life. Naturally they are found in

- (a) proteins
- (b) vitamins
- (c) alkaloids

#### (d) All of these

#### Question : Intermediates formed during reaction of RCONH<sub>2</sub> with Br<sub>2</sub> and KOH are

#### (a) **RCONHBr** and **RNCO**

(b) RNHCOBr and RNCO

(c) RNHBr and RCONHBr

#### (d) RCONBr<sub>2</sub>

#### Question : Primary amines can be distinguished from secondary and tertiary amines by reacting with

#### (a) Chloroform and alcoholic KOH

(b) Methyl iodide

(c) Chloroform alone

(d) Zinc dust

#### Question : Propionamide on Hofmann degradation gives -

(a) methyl amine

#### (b) ethyl amine

(c) propyl amine

(d) ethyl cyanide

#### Question : Secondary amines could be prepared by

- (a) reduction of nitriles
- (b) Hofmann bromamide reaction
- (c) reduction of amides

#### (d) reduction of isonitriles

#### Question : Gabriel's phthalimide synthesis is used for the preparation of

- (a) Primary aromatic amines
- (b) Secondary amines

#### (c) Primary aliphatic amines

(d) Tertiary amines

#### Question : Ethyl amine can be obtained by the

- (a) Action of  $NH_3$  on ethyl iodide.
- (b) Action of  $NH_3$  on ethyl alcohol.

#### (c) Both (a) and (b)

(d) Neither (a) nor (b)

#### Question : Treatment of ammonia with excess of ethyl iodide will yield

- (a) diethylamine
- (b) ethylamine
- (c) triethylamine

Question : For alkylation of ammonia which of the following is not used?

(a) CH<sub>3</sub>–X

(b) CH<sub>3</sub>–CH<sub>2</sub>–X

(c) (CH3)2CH–X

(d) (CH<sub>3</sub>)<sub>3</sub>C–X

#### **Question : Which of the following amines can be prepared by Gabriel method ?**

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- (i)  $CH_3CH_2NH_2$
- (ii)  $(CH_3)_2CHNH_2$
- (iii) (CH<sub>3</sub>)<sub>3</sub>CNH<sub>2</sub>
- (iv)  $C_6H_5NH_2$
- (a) (i) and (iii)
- (b) (ii) and (iv)
- (c) (i), (ii) and (iii)

#### (d) (i) and (ii)

Question : Amongst the given set of reactants, the most appropriate for preparing 2° amine is

- (a)  $2^{\circ}R-Br + NH_3$
- (b)  $2^{\circ}R$ –Br + NaCN followed by H<sub>2</sub>/Pt

#### (c) $1^{\circ}R-NH_2 + RCHO$ followed by $H_2/Pt$

(d)  $1^{\circ}R$ –Br (2 mol) + Potassium phthalimide followed by H<sub>3</sub>O+/heat

#### Question : The best reagent for converting 2 – phenylpropanamide into 2-phenylpropanamine is

(a) excess H<sub>2</sub>

------•

(b) Br<sub>2</sub>in aqueous NaOH

(c) iodine in the presence of red phosphorus

#### (d) LiAlH<sub>4</sub>in ether

#### Question : Which of the following methods of preparation of amines will give same number of carbon atoms in the chain of amines as in the reactant?

(a) Reaction of nitrite with LiAlH<sub>4</sub>.

(b) Reaction of amide with LiAlH<sub>4</sub> followed by treatment with water.

#### (c) Heating alkylhalide with potassium salt of phthalimide followed by hydrolysis.

(d) Treatment of amide with bromine in aquesous solution of sodium hydroxide.

Question : The reduction of nitro compounds is most preferred in the presence of

(a)  $Pd/H_2$  in ethanol

(b) Sn + HCl

(c) finely divided Ni

(d) iron scrap and HCl.

Question : An alkyl or benzyl halide on reaction with an ethanolic solution of ammonia undergoes

(a) electrophilic substitution reaction

#### (b) nucleophilic substitution reaction.

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(c) free radical mechanism.

(d) nucleophilic addition reaction.

#### Question : In the ammonolysis of alkyl halides the halogen atom is replaced by an amino(–NH2) group which of the following represent the correct order of reactivity of halides with amines.

(a) RBr > RI > RCI

(b) RI > RCI > RBr

#### (c) RI > RBr > RCl

(d) RCI > RBr > RI

#### **Question : Which of the following will give primary amine only ?**

(i) ammonia + propylchloride

(ii) potassium pthalimide + ethylchloride

(iii) potassium pthalimide + chlorobenzene

#### (a) (i) and (ii)

(b) (i) and (iii)

(c) (ii) and (iii)

(d) (i), (ii) and (iii)

#### **Question : Amines have**

(a) Garlic odour

#### (b) Fishy odour

(c) Jasmine odour

(d) Bitter almonds odour

#### Question : Aniline is less soluble in water than ethyl amine due to

(a) resonance stablization of benzene ring

(b) resonance stabilization of anilium ion

#### (c) more hydrophobic nature of C<sub>6</sub>H<sub>5</sub> group than C<sub>2</sub>H<sub>5</sub> group

(d) more hydrophobic nature of  $C_6H_5$  group than  $C_2H_5$  Group

#### **Question : Which of the following is not correct ?**

(a) Ethyl amine and aniline both have –  $\mathsf{NH}_2$  group

(b) Ethyl amine and aniline dissolve in HCl

(c) Ethyl amine and aniline both react with CHCl<sub>3</sub> and KOH to form unpleasant smelling compound

(d) Ethyl amine and aniline both react with HNO<sub>2</sub> in cold to give hydroxy compounds

#### **Question : Amines behave as**

(a) lewis acids

#### (b) lewis bases

- (c) aprotic acids
- (d) amphoteric compounds

#### Question : The basic character of amines is due to

(a) presence of nitrogen atom

#### (b) lone pair of electrons on nitrogen atom

- (c) tetrahedral structure
- (d) high electronegativity of nitrogen

#### Question : Aliphatic amines are.....basic than NH3 but aromatic amines are.....basic than NH3.

#### (a) more, less

- (b) less, more
- (c) both (a) and (b)
- (d) None of these

#### Question : Substitution of one alkyl group by replacing hydrogen of primary amines

#### (a) increases the base strength

- (b) decreases the base strength
- (c) remains the same
- (d) None of the above

#### Question : Which of the following is not characteristic of amines?

- (a) They smell like ammonia
- (b) They are inflammable in air
- (c) They show the property of hydrogen bonding

#### (d) They are amphoteric in nature

Question : The correct order of basicity in amines

(i)  $C_2H_5NH_2$ 

(ii) CH<sub>3</sub>NH<sub>2</sub>

(iii) (CH<sub>3</sub>)<sub>2</sub>NH

(iv) (CH<sub>3</sub>)<sub>3</sub>N

(a) (i) < (iv) < (ii) < (iii)

(b) (iv) < (ii) < (iii) < (i)

(c) (i) < (ii) < (iii) < (iv)

(d) (ii) < (iii) < (iv) < (i)

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#### Question : The conjugate base of $(CH_3)_2 NH^{+2}$ is

#### (a) (CH<sub>3</sub>)<sub>2</sub>NH

(b) (CH<sub>3</sub>)<sub>2</sub>N+\

(c) (CH<sub>3</sub>)<sub>3</sub>N+

(d) (CH<sub>3</sub>)<sub>2</sub>N-

#### Question : High basicity of Me<sub>2</sub>NH relative to Me<sub>3</sub>N is attributed to:

#### (a) effect of solvent

(b) inductive effect of Me\

(c) shape of Me<sub>2</sub>NH

(d) shape of  $Me_3N$ 

#### **Question : Hinsberg reagent is**

(a)  $C_6H_5SO_3H$ 

(b)  $C_6H_5NO$ 

#### (c) $C_6H_5SO_2CI$

(d)  $C_6H_5N_2CI$ 

#### **Question : Which of the following statement is correct ?**

(a) Ammonia is more basic than methylamine.

#### (b) Methylamine is more basic than ammonia.

(c) Dimethylamine is less basic than methylamine.

(d) Dimethylamine is less basic than trimethylamine.

#### Question : Reaction of aniline with benzaldehyde is

(a) substitution

(b) addition

#### (c) condensation

#### (d) polymerization

Question : The amine that does not react with acetyl chloride is

(a) CH<sub>3</sub>NH<sub>2</sub>

(b) (CH<sub>3</sub>)<sub>2</sub>NH

#### (c) (CH<sub>3</sub>)<sub>3</sub>N

(d) None of these

Question : The correct decreasing order of basic strength of the following species is \_ H2O, NH3, OH, NH2 -

#### (a) $NH_2 - > OH - > NH_3 > H_2O$

(b)  $OH - > NH_2 - > H_2O > NH_3$ 

(c)  $NH_3 > H_2O > NH_2 - > OH_-$ 

(d)  $H_2O > NH_3 > OH -> NH_2 -$ 

#### Question : Which of the following factors affect the basic strength of amine?

- (i) Inductive effect
- (ii) Steric hinderance
- (iii) Solvation effect
- (iv) Solubility in organic solvents.
- (a) (i) and (iv)

#### (b) (i), (ii) and (iii)

- (c) (ii) and (iii)
- (d) (ii) and (iv)

#### **Question : Which statement is not true among the following?**

- (a) Amines are bases
- (b) They turn red litmus blue
- (c) Trimethyl amine is less basic than dimethyl amine

#### (d) Amines yield alcohols on aqueous hydrolysis.

#### **Question : Aniline is used**

(a) in crimping of wool

#### (b) in dyeing industry

- (c) in making of glue
- (d) in fast drying vanish

Question : Which of the following statements about primary amines is 'False' ?

(a) Alkyl amines are stronger bases than aryl amines

(b) Alkyl amines react with nitrous acid to produce alcohols

(c) Aryl amines react with nitrous acid to produce phenols

(d) Alkyl amines are stronger bases than ammonia

#### **Question : Mark the correct statement**

(a) Methylamine is slightly acidic

(b) Methylamine is less basic than ammonia

#### (c) Methylamine is a stronger base than ammonia

(d) Methylamine forms salts with alkalies.

#### Question : For carbylamine reaction, we need hot alcoholic KOH and

#### (a) any primary amine and chloroform

- (b) chloroform and silver powder
- (c) a primary amine and an alkyl halide
- (d) a monoalkylamine and trichloromethane.

# Question : The compound obtained by heating a mixture of a primary amine and chloroform with ethanolic potassium hydroxide (KOH) is

- (a) an alkyl cyanide
- (b) a nitro compound

#### (c) an alkyl isocyanide

(d) an amide

#### Question : Which of the following compounds cannot be identified by carbylamine test?

- (a)  $CH_3CH_2NH_2$
- (b)  $CHCl_3$
- (c)  $C_6H_5NH_2$

#### (d) $C_6H_5-NH-C_6H_5$

#### **Question : Carbylamine reaction is used for the detection of**

- (a) aliphatic 2° amines
- (b) aliphatic 1° amines
- (c) aromatic 1° amines

#### (d) Both (b) and (c)

# Question : An organic amino compound reacts with aqueous nitrous acid at low temperature to produce an oily nitrosoamine. The compound is

(a)  $CH_3NH_2$ 

(b) CH<sub>3</sub>CH<sub>2</sub>NH<sub>2</sub>

(c) CH<sub>3</sub>CH<sub>2</sub>NHCH<sub>2</sub>CH<sub>3</sub>

(d)  $(CH_3CH_2)_3N$ 

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CBSE vide Circular No.Acad-51/2021 dated 5th July, 2021, notified that in the session 2021-2022, Board Examinations would be conducted in two terms, i.e.. Term I and Term II. This decision was taken due to the uncertainty arising out of COVID

### <u>National Youth Day and</u> <u>Birth Anniversary of Swami</u> <u>Vivekananda</u>

Ministry of Education, Govt. of India vide D.O No. 12-4/2021-IS.4 dated 04.01.2022 intimated that 12 January 2022 will be celebrated as "National Youth Day" and "Birth Anniversary of Swami Vivekananda". All Schools affiliated to CBSE may celebrate 12 January 2022 as...

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The 5th edition of Pariskhas Pe Charcha the unique interactive program of Hon'ble Prime Minister with students teaches and parents will be held through virtual mode in February, 2022. In order to select participants who will be featured in Pariksha Pe Charcha programme...

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You are aware that session 2022 has also been affected severely by Covid and the session has already been delayed, now there is a need for the remaining activities of this session to be completed on time. As per the earlier circular, students of

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# **CBSE Class 12 Chemistry Biomolecules MCQs**

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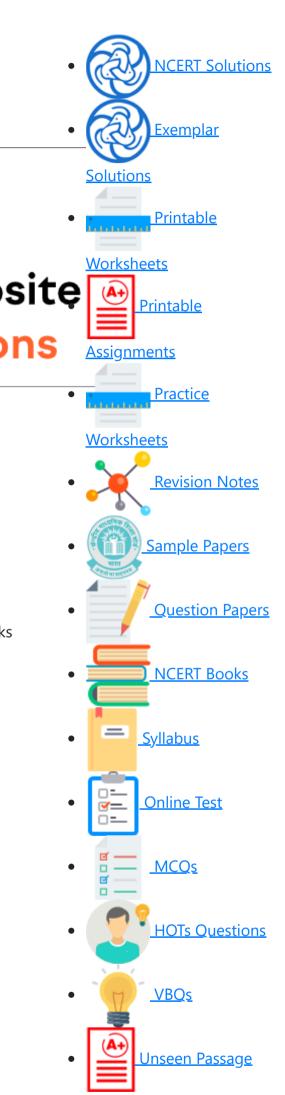
CBSE Class 12 Chemistry Biomolecules MCQs with answers available in Pdf for free download. The MCQ Questions for Class 12 Chemistry with answers have been prepared as per the latest syllabus, NCERT books and examination pattern suggested in Standard 12 by CBSE, NCERT and KVS. Multiple Choice Questions are an important part of Term 1 and Term 2 exams for Grade 12 Chemistry and if practiced properly can help you to get higher marks. Refer to more Chapter-wise MCQs for NCERT Class 12 Chemistry and also download more latest study material for all subjects

# **Biomolecules Class 12 Chemistry MCQ**

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Biomolecules in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

### **Biomolecules MCQ Questions Class 12 Chemistry with** Answers

**Question : Which of the following gives positive Fehling solution test?** 



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(a) Protein

(b) Sucrose

(c) Glucose

(d) Fats

Answer : C

#### Question : Which of the following statements is incorrect regarding glucose?

(a) It is an aldohexose. (b) It is also known as dextrose (c) It is monomer of cellulose.

(d) It is the least abundant organic compound on earth.



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#### Answer : D

#### Question : Glucose gives silver mirror test with Tollen's reagent. It shows the presence of

- (a) acidic group
- (b) alcoholic group
- (c) ketonic group
- (d) aldehyde group

#### **Answer : D**

#### **Question : The symbols D and L represents**

- (a) the optical activity of compounds.
- (b) the relative configuration of a particular stereoisomer.
- (c) the dextrorotatory nature of molecule.
- (d) the levorotatory nature of molecule

#### Answer : B

#### Question : The function of glucose is to

- (a) provides energy
- (b) promote growth
- (c) prevent diseases
- (d) perform all above

#### **Answer: A**

#### Question : Which one of the following compounds is different from the rest?

(a) Sucrose(b) Maltose(c) Lactose(d) Glucose

#### Answer : D

#### Question : The two functional groups present in a typical carbohydrate are:

(a) – CHO and – COOH
(b) > C = O and – OH
(c) – OH and – CHO
(d) – OH and – COOH

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#### Answer : C

#### Question : When glucose reacts with bromine water, the main product is

(a) gluconic acid(b) glyceraldehyde(c) saccharic acid

(d) acetic acid

#### Answer : A

**Question : Glucose does not react with** 

(a)  $Br_2/H_2O$ (b) H<sub>2</sub>NOH (c) HI (d) NaHSO<sub>3</sub>

#### **Answer : D**

#### Question : Glucose reacts with acetic anhydride to form

- (a) monoacetate
- (b) tetra-acetate
- (c) penta-acetate
- (d) hexa-acetate

#### **Answer : C**

#### **Question : Biomolecules are**

- (a) aldehydes and ketones
- (b) acids and esters
- (c) carbohydrates, proteins and fats
- (d) alcohols and phenols

#### Answer : C

#### Question : Which of the following is a disaccharide ?

- (a) Lactose
- (b) Starch
- (c) Cellulose
- (d) Fructose

#### **Answer: A**

#### Question : The sugar that is characteristic of milk is

- (a) maltose
- (b) ribose
- (c) lactose
- (d) galactose



#### Answer : C

#### **Question : Which one is a disaccharide ?**

(a) Glucose

(b) Fructose

(c) Xylose

(d) Sucrose

#### **Answer : D**

#### Question : Which of the following monosaccharide is pentose ?

(a) Glucose

- (b) Fructose
- (c) Arabinose
- (d) Galactose

#### Answer : C

#### Question : The commonest disaccharide has the molecular formula

- (a) C<sub>10</sub>H<sub>18</sub>O<sub>9</sub>
- (b) C<sub>10</sub>H<sub>20</sub>O<sub>10</sub>
- (c) C<sub>18</sub>H<sub>22</sub>O<sub>11</sub>
- (d) C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>

#### Answer : D

#### Question : Monosaccharides usually contains ... carbon atoms.

- (a)  $C_3$  to  $C_{10}$
- (b) C<sub>1</sub> to C<sub>6</sub>
- (c) C<sub>4</sub> to C<sub>10</sub>
- (d)  $C_5$  to  $C_8$

#### Answer : A

#### Question : Which one of the following compounds is found abudnantly in nature?

- (a) Fructose
- (b) Starch
- (c) Glucose
- (d) Cellulose
- Answer : D

#### Question : A carbohydrate that cannot be hydrolysed into simpler units is called

(a) polysaccharides

(b) trisaccharides

(c) disachharides

(d) monosaccharides

Answer : D

**Question : Which of the following statements is incorrect ?** 

(a) Maltose gives two molecules of glucose only.

- (b) Cellulose and sucrose are polysaccharide.
- (c) Polysaccharides are not sweet in taste.

(d) Polysaccharides are also known as non-sugars.

#### **Answer: B**

#### **Question : Reducing sugars reduce.**

- (a) only Fehling's solution
- (b) only Tollen's solution.
- (c) both (a) & (b)
- (d) neither (a) nor (b)

#### **Answer : C**

#### Question : Which among the following is the simplest sugar?

- (a) Glucose
- (b) Starch
- (c) Cellulose
- (d) None of these

#### **Answer: A**

#### Question : Glucose can't be classified as

- (a) hexose
- (b) carbohydrate
- (c) aldose
- (d) oligosaccharide

#### **Answer : D**

#### Question : Which of the following properties of glucose cannot be explained by its open chain structure?

- (i) Glucose does not form hydrogen sulphite with NaHSO3
- (ii) On oxidation with  $\mathsf{HNO}_3$  glucose gives saccharic acid.

(iii) Glucose is found to exist in two different crystalline forms which are named as  $\alpha$  and  $\beta$ .

(a) (ii) only

(b) (i) and (iii)

(c) (ii) and (iii)

(d) (i) and (ii)

**Answer : B** 

**Question : Which of the following gives positive Fehling solution test?** 

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- (a) Protein
- (b) Sucrose
- (c) Glucose
- (d) Fats

#### **Answer : C**

#### Question : Which of the following statements is incorrect regarding glucose?

- (a) It is an aldohexose.
- (b) It is also known as dextrose
- (c) It is monomer of cellulose.
- (d) It is the least abundant organic compound on earth.

#### **Answer : D**

#### Question : Glucose gives silver mirror test with Tollen's reagent. It shows the presence of

- (a) acidic group
- (b) alcoholic group
- (c) ketonic group
- (d) aldehyde group

#### **Answer : D**

#### **Question : The symbols D and L represents**

- (a) the optical activity of compounds.
- (b) the relative configuration of a particular stereoisomer.
- (c) the dextrorotatory nature of molecule.
- (d) the levorotatory nature of molecule

#### **Answer: B**

#### Question : Glucose is found to exist in two different $\alpha$ and $\beta$ crystalline forms. These forms can be obtained by.

(i) The  $\alpha$  form of glucose is obtained by crystallisation from concentrated solution of glucose at 303 K.

(ii) The  $\beta$  form of glucose is obtained by crystallisation from concentrated solution of glucose at 303 K.

(iii) The  $\beta$  form is obtained by crystallisation from hot and saturated aqueous solution at 371 K.

(iv) The  $\alpha$  form is obtained by crystallisation from hot and saturated aqueous solution at 371 K.

(a) (i) and (iii)

(b) (ii) and (iv)

(c) (ii) and (iii)

(d) (i) only

#### **Answer: B**

#### Question : The function of glucose is to

- (a) provides energy
- (b) promote growth
- (c) prevent diseases
- (d) perform all above

#### **Answer : A**

#### Question : Which one of the following compounds is different from the rest?

- (a) Sucrose
- (b) Maltose
- (c) Lactose
- (d) Glucose

#### Answer : B

#### Question : The two functional groups present in a typical carbohydrate are:

- (a) CHO and COOH
- (b) > C = O and -OH
- (c) OH and CHO
- (d) OH and COOH

#### Answer : C

#### Question : When glucose reacts with bromine water, the main product is

- (a) gluconic acid
- (b) glyceraldehyde
- (c) saccharic acid
- (d) acetic acid

#### **Answer: A**

#### **Question : Glucose does not react with**

(a)  $Br_2/H_2O$ 

(b) H<sub>2</sub>NOH

(c) HI

(d) NaHSO<sub>3</sub>

Answer : D

**Question : Glucose reacts with acetic anhydride to form** 

(a) monoacetate

- (b) tetra-acetate
- (c) penta-acetate
- (d) hexa-acetate

#### Answer : C

#### Question : Reduction of glucose by HI suggest that

- (a) presence of OH groups
- (b) presence of –CHO group
- (c) cyclic structure of glucose
- (d) six carbon atoms are arranged in straight chain

#### Answer : D

#### Question : The reaction of glucose with red P + HI is called

- (a) Sandmeyer's reaction
- (b) Reformatsky reaction
- (c) Gattermann's reaction
- (d) Reduction

#### Answer : D

#### Question : Which of the following reactions of glucose can be explained only by its cyclic structure?

- (a) Glucose forms pentaacetate
- (b) Glucose reacts with hydroxylamine to form an oxime
- (c) Pentaacetate of glucose does not react with hydroxylamine
- (d) Glucose is oxidised by nitric acid to gluconic acid

#### Answer : C

#### **Question : Which is the least stable form of glucose ?**

- (a)  $\alpha$ -D-Glucose
- (b) **β**-D-Glucose

(c) Open chain structure

(d) All are equally stable

Answer : C

**Question : Isomerization of glucose produces** 

(a) galactose

(b) fructose

(c) mannose

(d) allose

#### **Answer : B**

#### Question : A solution of D-glucose in water rotates the plane polarised light

- (a) to the right
- (b) to the left
- (c) to either side
- (d) None of these

#### **Answer: A**

#### Question : The number of chiral carbon atoms present in cyclic structure $\alpha$ -D(+) glucose

(a) 3			
(b) 4			
(c) 5			
(d) 6			

#### Answer : C

#### Question : The $\alpha$ -D glucose and $\beta$ -D glucose differ from each other due to difference in carbon atom with respect to its

- (a) conformation
- (b) configuration
- (c) number of OH groups
- (d) size of hemiacetal ring

#### **Answer: B**

#### Question : The two forms of D-glucopyranose obtained from the solution of D-glucose are called

- (a) isomers
- (b) anomers
- (c) epimers
- (d) enantiomers

#### Answer : B



#### Question : Which of the following carbohydrates are branched polymer of glucose?

(i) Amylose (ii) Amylopectin

(iii) Cellulose (iv) Glycogen

(a) (i) and (ii)

(b) (ii) and (iv)

(c) (iii) and (iv)

(d) (i), (ii) and (iii)

#### Question : The number of chiral carbon atoms present in cyclic structure $\alpha$ -D(+) glucose

(a) 3			
(b) 4			
(c) 6			
(d) 5			

#### Answer : C

#### Question : Which of the following reagent cannot distinguish between glucose and fructose?

- (a) Fehling's solution
- (b) Tollen's reagent
- (c) Benedict's solution
- (d) All of these

#### Answer : C

#### **Question : Maltose and glucose are**

- (a) oxidising sugar
- (b) reducing sugar
- (c) first is oxidising and second is reducing sugar
- (d) both are non-reducing sugar

#### Answer : B

#### Question : Choose the correct relationship for glucose and fructose

- (a) these are functional isomers
- (b) these are chain isomers
- (c) these are position isomers
- (d) All of these

#### Answer : A

Question : The pair of compounds in which both the compounds give positive test with Tollen's reagent is

(a) Glucose and Sucrose

(b) Fructose and Sucrose

(c) Acetophenone and Hexanal

(d) Glucose and Fructose

**Question : The letter D and L in carbohydrates represent** 

- (a) its optical rotation
- (b) its mutarotation
- (c) its direct synthesis
- (d) its configuration

#### **Answer : D**

#### Question : Which of the following statement is correct about fructose?

- (a) It is dextrorotatory compound
- (b) It exists in the two cyclic forms which is obtained by the addition of OH at C-5 to the >C=O group
- (c) It exists as six membered ring
- (d) It is named as furanose as it contain one oxygen and six carbon atom

#### **Answer: B**

#### **Question : Fructose is**

- (a) a hemiacetal
- (b) an acetal
- (c) a hemiketal
- (d) a ketal

#### **Answer : C**

#### **Question : The sugar present in fruits is**

- (a) fructose
- (b) glucose
- (c) sucrose
- (d) galactose

#### **Answer: A**

Question : Which of the following carbohydrate does not correspond to the general formula  $Cx(H_2O)y$ ?

(a) Glucose

(b) 2-Deoxyribose

(c) Fructose

(d) Arabinose

**Answer : B** 

Question : The sugar present in honey is

(a) sucrose

(b) glucose

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(c) fructose

#### (d) maltose

#### Answer : C

#### Question : Which of the following is the sweetest sugar?

- (a) Sucrose
- (b) Glucose
- (c) Fructose
- (d) Maltose

#### Answer : C

#### **Question : Cellulose is a polymer of**

- (a) Glucose
- (b) Fructose
- (c) Ribose
- (d) Sucrose

#### **Answer: A**

#### **Question : Sucrose on hydrolysis gives**

- (a) fructose+ribose
- (b) glucose + fructose
- (c) glucose+glucose
- (d) fructose + fructose

#### **Answer: B**

#### Question : The presence or absence of hydroxyl group on which carbon atom of sugar differentiates **RNA and DNA?**

(a) 1st

(b) 2nd

(c) 3rd

(d) 4th

#### **Answer: B**

#### **Question : Biomolecules are**

(a) aldehydes and ketones (b) acids and esters (c) carbohydrates, proteins and fats (d) alcohols and phenols

#### Answer : C

#### Question : Which of the following is a disaccharide ?

(a) Lactose

- (b) Starch
- (c) Cellulose
- (d) Fructose

#### Answer : A

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(c) lactose

(d) galactose

#### Answer : C

#### **Question : Which one is a disaccharide ?**

(a) Glucose

(b) Fructose

(c) Xylose

(d) Sucrose

#### Answer : D

#### Question : Which of the following monosaccharide is pentose ?

(a) Glucose

- (b) Fructose
- (c) Arabinose
- (d) Galactose

#### Answer : C

#### Question : Which one of the following compounds is found abudnantly in nature?

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(b) Starch

(c) Glucose

(d) Cellulose

Answer : D

Question : A carbohydrate that cannot be hydrolysed into simpler units is called

(a) polysaccharides

(b) trisaccharides

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(c) disachharides

(d) monosaccharides

#### **Answer : D**

#### **Question : Which of the following statements is incorrect ?**

(a) Maltose gives two molecules of glucose only.

- (b) Cellulose and sucrose are polysaccharide.
- (c) Polysaccharides are not sweet in taste.
- (d) Polysaccharides are also known as non-sugars

#### Answer : B

#### **Question : Reducing sugars reduce.**

(a) only Fehling's solution(b) only Tollen's solution.(c) both (a) & (b)(d) neither (a) nor (b)

#### Answer : C

#### Question : Which among the following is the simplest sugar?

(a) Glucose

(b) Starch

(c) Cellulose

(d) None of these

#### Answer : A

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# CBSE Class 12 Chemistry Chemical Kinetics MCQs Set A

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# **Chemical Kinetics Class 12 Chemistry MCQ**

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Chemical Kinetics in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

### Chemical Kinetics MCQ Questions Class 12 Chemistry with Answers

Question : The rate constant of first order reaction is  $3 \times 10^{-6}$  per second. The initial concentration is 0.10 M. The initial rate is:

(a)  $3 \times 10^{-7}$  mol/litre/sec (b)  $3 \times 10^{-8}$  mol/litre/sec (c)  $3 \times 10^{-5}$  mol/litre/sec (d)  $3 \times 10^{-8}$  mol/litre/sec

#### **Answer : A**

**Question : Choose the correct options** 

#### For the reaction :

 $H_2 + Cl_2 \xrightarrow{\text{sunlight}} 2HCl$ the order of reaction is



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(a) 0

(b) 2

(c) 1

(d) 3

**Answer: A** 

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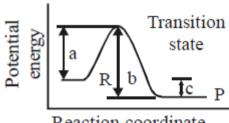
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#### **Question : Choose the correct options**

The potential energy diagram for a reaction

$$R \longrightarrow P$$
 is given below

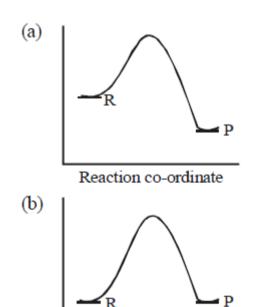


Reaction coordinate

- $\Delta H^{\circ}$  of the reaction corresponds to the energy :
- (a) a (b) b (c) c
- (d) a + b

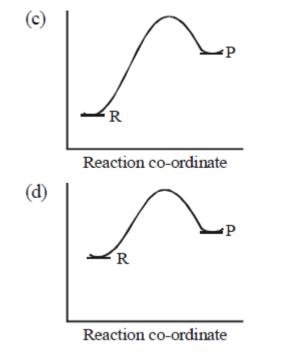
#### **Answer : C**

by the diagram :





Reaction co-ordinate



#### Answer: B

#### **Question : Choose the correct option**

For the reaction,

 $2N_2O_5 \longrightarrow 4NO_2 + O_2$ the rate of reaction is :

(a) 
$$\frac{1}{2} \frac{d}{dt} [N_2 O_5]$$
 (b)  $2 \frac{d}{dt} [N_2 O_5]$   
(c)  $\frac{1}{4} \frac{d}{dt} [NO_2]$  (d)  $4 \frac{d}{dt} [NO_2]$ 

#### Answer : C

Question : For a first order reaction, to obtain a positive slope, we need to plot {where [A] is the concentration of reactant A}

(a)  $-\log_{10}[A]$  vs t (b)  $-\log_{e}[A]$  vs t (c)  $\log_{10}[A]$  vs log t (d) [A] vs t

#### Answer : C

Question : In most cases, for a rise of 10K temperature the rate constant is doubled to tripled. This is due to the reason that

(a) collision frequency increases by a factor of 2 to 3.

(b) fraction of molecules possessing threshold energy increases by a factor of 2 to 3

(c) Activation energy is lowered by a factor of 2 to 3.

(d) none of these

#### Answer : A

Question : Half-lives of a first order and a zero order reaction are same. Then the ratio of the initial rates of first order reaction to that of the zero order reaction is

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(a) 
$$\frac{1}{0.693}$$
 (b)  $2 \times 0.693$   
(c)  $0.693$  (d)  $\frac{2}{0.693}$ 

#### Answer : B

Question : Which of the following relation represents correct relation between standard electrode potential and equilibrium constant?

I.  $\log K = \frac{nFE^{\circ}}{2.303 \text{ RT}}$ II.  $K = e^{\frac{nFE^{\circ}}{RT}}$ III.  $\log K = \frac{nFE^{\circ}}{2.303 \text{ RT}}$ IV.  $\log K = 0.4342 \frac{nFE^{\circ}}{RT}$ Choose the correct statement(s).

(a) I, II and III are correct

(b) II and III are correct

(c) I, II and IV are correct

(d) I and IV are correct

#### Answer : D

Question : Collision theory is used to explain how chemical species undergo a reaction. Using this theory and the kinetic molecular model, which of the following does NOT influence the rate of a chemical reaction?

(a) The temperature of the system

- (b) The geometry or orientation of the collision
- (c) The velocity of the reactants at the point of collision
- (d) All of the above influence the rate

#### **Answer : A**

#### Question : The rate of a chemical reaction tells us about

(a) the reactants taking part in reaction

(b) the products formed in the reaction

(c) how slow or fast the reaction is taking place

(d) none of the above

#### Answer : C

#### Question : In the rate equation, when the concentration of reactants is unity then rate is equal to

- (a) specific rate constant
- (b) average rate constant
- (c) stantaneous rate constant
- (d)) None of the above

#### Answer : A

#### Question : The role of a catalyst is to change

(a) gibbs energy of reaction.

- (b) enthalpy of reaction.
- (c) activation energy of reaction.

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(d) equilibrium constant.

### Answer : C

### Question : In the presence of a catalyst, the heat evolved or absorbed during the reaction\_\_\_\_\_

(a) increases.

(b) decreases.

(c) remains unchanged.

(d) may increase or decrease.

### Answer : C

### Question : Which of the following statements is not correct about order of a reaction.

(a) The order of a reaction can be a fractional number.

(b) Order of a reaction is experimentally determined quantity.

(c) The order of a reaction is always equal to the sum of the stoichiometriccoefficients of reactants in the balanced chemical equation for a reaction.

(d) The order of a reaction is the sum of the powers of molar concentration of the reactants in the rate law expression

### Answer : C

### **Question : Which of the following statements is correct?**

(a) The rate of a reaction decreases with passage of time as the concentration of reactants dereases.

- (b) The rate of a reaction is same at any time during the reaction.
- (c) The rate of a reaction is independent of temperature change.

(d) The rate of a reaction decreases with increase in concentration ofreactant(s).

### Answer : A

# Question : Which of the following statements is incorrect about the collison theory of chemical reaction?

(a) It considers reacting molecules or atoms to be hard spheres and ignorestheir structural features.

(b) Number of effective collisions determines the rate of reaction.

(c) Collision of atoms or molecules possessing sufficient threshold energyresults into the product formation.

(d) Molecules should collide with sufficient threshold energy and properorientation for the collision to be effective.

### Answer : C

### Question : Which of the following statement is not correct for the catalyst?

(a) It catalyses the forward and backward reaction to the same extent.

(b) It alters  $\Delta G$  of the reaction.

(c) It is a substance that does not change the equilibrium constant of areaction.

(d) It provides an alternate mechanism by reducing activation energybetween reactants and products.

### Answer : B

### Question : The value of rate constant of a pseudo first order reaction \_\_\_\_\_\_.

(a) depends on the concentration of reactants present in small amount.(b) depends on the concentration of reactants present in excess.(c) is independent of the concentration of reactants.(d) depends only on temperature.

### Answer : B

### Question : The term – dc/dt in a rate equation refers to :

- (a) the conc. of a reactant
- (b) the decrease in conc. of the reactant with time
- (c) the velocity constant of reaction
- (d) None of these

### **Answer : B**

### Question : The rate law for the single- step reaction $2A + B \rightarrow 2C$ , is given by:

(a) rate = k [A].[B](b) rate = k [A]2.[B](c) rate = k [2A].[B](d) rate = k [A]2.[B]o

### **Answer: B**

### **Question : Rate of which reaction increases with temperature :**

- (a) of any type of reactions
- (b) of exothemic reactions
- (c) of endothemic reactions
- (d) of none

### **Answer: A**

### Question : In a slow reaction, rate of reaction generally ...... with time:

- (a) decreases
- (b) increases
- (c) sometimes increases and sometimes decreases.
- (d) remains constant

### **Answer: A**

### Question : The rate of a chemical reaction tells us about,

- (a) the reactants taking part in reaction

(b) the products formed in the reaction

(c) how slow or fast the reaction is taking place

(d) None of the above

### **Answer : C**

### **Question : Point out the wrong statement: For a first order reaction**

(a) time for half-change  $(t_{1/2})$  is independent of initial concentration

(b) change in the concentration unit does not change the rate constant (k)

(c) time for half-change  $\times$  rate constant = 0.693

(d) the unit of k is  $mole^{-1} min^{-1}$ 

### **Answer : D**

Question : Consider the reaction,  $2A + B \rightarrow$  products. When concentration of B alone was doubled, the alf-life did notchange. When the concentration of A alone was doubled, the rate increased by two imes. The unit of rate constant for this reaction is

(a)  $s^{-1}$ 

(b) L mol-1 s $^{-1}$ 

(c) no unit

(d) mol  $L^{-1} s^{-1}$ .

**Answer: B** 

Question : The decomposition of  $N_2O_5$  occurs as  $2N_2O_5 \rightarrow 4NO_2 + O_2$  and follows lst order kinetics, hence:

(a) the reaction is unimolecular

(b) the reaction is bimolecular

(c)  $t_{1/2} \propto a$ 

(d) None of these

### Answer : C

### Question : Which of the following reaction does not occur fastly ?

(a) Precipitation of AgCl by mixing aqueous solutions of  $AgNO_3$  and NaCl.

(b) Burning of gasoline

(c) Rusting of iron

(d) Burning of LPG for cooking

Answer : C

### Question : Chemical kinetics is a study to find out

(a) the feasibility of a chemical reaction

(b) extent to which a reaction will proceed

(c) speed of a reaction

(d) All of the above

Answer : C

### Question : Rate of a reaction can be defined as

(a) the rate of decrease in concentration of any one of the reactants

(b) the rate of increase in concentration of any one of the products

(c) the rate of decrease in concentration of any one of the reactants or the rate of increase in concentration of any one of the products

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(d) the sum of rate of decrease in concentration of all the reactants or the rate of increase in concentration of all the products

### Answer : C

### **Question : The rate of reaction**

- (a) increases as the reaction proceeds
- (b) decreases as the reaction proceeds
- (c) remains the same as the reaction proceeds
- (d) may decrease or increase as the reaction proceeds

### **Answer : D**

### Question : The unit of rate of reaction is

- (a) mole/dm<sup>3</sup>
- (b) mole/pound
- (c) mole/dm<sup>3</sup> sec
- (d) mole/ $cm^3$

### Answer : C

### Question : In the rate equation, when the conc. of reactants is unity then rate is equal to

- (a) specific rate constant
- (b) average rate constant
- (c) instantaneous rate constant
- (d) None of above

### **Answer: A**

### Question : The rate of reaction between two specific time intervals is called

- (a) instantaneous rate
- (b) average rate
- (c) specific rate
- (d) ordinary rate

#### **Answer : B**

### Question : Instantaneous rate of a chemical reaction is

(a) rate of reaction in the beginning

(b) rate of reaction at the end

(c) rate of reaction at a given instant

(d) rate of reaction between two specific time intervals

### **Answer : C**

### Question : At the beginning the decrease in the conc. of reactants is

(a) slow

- (b) moderate
- (c) rapid
- (d) None of above

### Answer : C

### Question : The average rate and instantaneous rate of a reaction are equal

- (a) at the start
- (b) at the end
- (c) in the middle
- (d) when two rate have time interval equal to zero

### Answer : D

### Question : The rate of reaction depends upon the

(a) volume

- (b) force
- (c) pressure
- (d) conc. of reactants

### **Answer : D**

Question : For the following reaction:  $NO_2(g) + CO(g) \rightarrow NO(g) + CO_2(g)$ , the rate law is: Rate = k  $[NO_2]_2$ . If 0.1 mole of gaseous carbon monoxide is added at constant temperature to the reaction mixture which of the following statements is true?

- (a) Both k and the reaction rate remain the same
- (b) Both k and the reaction rate increase
- (c) Both k and the reaction rate decrease
- (d) Only k increases, the reaction rate remain the same

### **Answer : A**

### Question : Which one of the following statements for the order of a reaction is incorrect?

(a) Order can be determined only experimentally.

(b) Order is not influenced by stoichiometric coefficient of the reactants.

(c) Order of reaction is sum of power to the concentration terms of reactants to express the rate of reaction.

(d) Order of reaction is always whole number.

### Answer : D

# Question : The rate of the reaction 2NO + $Cl_2 \rightarrow 2NOCl$ is given by the rate equation rate = k [NO]<sub>2</sub> [ $Cl_2$ ]

### The value of the rate constant can be increased by:

(a) increasing the concentration of NO.

- (b) increasing the temperature.
- (c) increasing the concentration of the  $Cl_2$
- (d) doing all of the above

#### Answer : B

#### **Question : Order of reaction can be**

(a) 0

(b) fraction

(c) whole number

(d) integer, fraction, zero

#### Answer : D

# Question : Units of rate constant of first and zero order reactions in terms of molarity M unit are respectively

(a)  $\sec^{-1}$ ,  $\operatorname{Msec}^{-1}$ 

(b) sec<sup>-1</sup>, M

(c) Msec-1, sec $^{-1}$ 

(d) M,  $sec^{-1}$ .

### Answer : A

### Question : A reaction involving two different reactants can never be

- (a) bimolecular reaction
- (b) second order reaction
- (c) first order reaction
- (d) unimolecular reaction

#### **Answer : D**

### **Question : 3A \rightarrow B + C, it would be a zero order reaction when**

(a) the rate of reaction is proportional to square of concentration of A

(b) the rate of reaction remains same at any concentration of A

(c) the rate remains unchanged at any concentration of B and C

(d) the rate of reaction doubles if concentration of B is increased to double

### Answer : B

# Question : A reaction proceeds by first order, 75% of this reaction was completed in 32 min. The time equired for 50% completion is

(a) 8 min

### (b) 16 min

(c) 20 min

(d) 24 min

### Answer : B

### Question : Order of reaction is decided by

- (a) temperature
- (b) mechanism of reaction as well as relative concentration of reactants
- (c) molecularity
- (d) pressure

### Answer : B

### Question : Velocity constant k of a reaction is affected by

- (a) change in the concentration of the reactant
- (b) change of temperature
- (c) change in the concentration of the product
- (d) None of the above

### Answer : B

### **Question : The rate constant for the reaction**

 $2N_2O_5 \rightarrow 4NO_2 + O_2$  is  $3.10 \times 10^{-5}$  sec–1. If the rate is  $2.4 \times 10^{-5}$  mol litre–1 sec<sup>-1</sup> then the concentration of  $N_2O_5$  (in mol litre–1) is :

- (a) 0.04
- (b) 0.8
- (c) 0.07
- (d) 1.4

### Answer : B

### Question : A zero order reaction is one whose rate is independent of

(a) the concentration of the reactants

(b) the temperature of reaction

(c) the concentration of the product

(d) the material of the vessel in which reaction is carried out

### Answer : A

Question : The rate law for a reaction between the substances A and B is given by Rate = k [A]n [B]m On doubling the concentration of A and halving the concentration of B, the ratio of the new rate to the earlier rate of the reaction will be as

- (a) (m + n)
- (b) (n m)
- (c) 2(n m)
- (d) (m n)

### Answer : C

# Question : In the reaction $2A + B \rightarrow A2B$ , if the concentration of A is doubled and that of B is halved, then the rate of the reaction will:

- (a) increase 2 times
- (b) increase 4 times
- (c) decrease 2 times
- (d) remain the same

### **Answer : A**

# Question : The order of a reaction, with respect to one of the reacting component Y, is zero. It implies that:

- (a) the reaction is going on at a constant rate
- (b) the rate of reaction does not vary with temperature
- (c) the reaction rate is independent of the concentration of Y
- (d) the rate of formation of the activated complex is zero

### Answer : C

### Question : If the rate of a gaseous reaction is independent of pressure, the order of reaction is:

- (a) 0 (b) 1 (c) 2
- (d) 3

### **Answer : A**

Question : If the rate of the reaction is equal to the rate constant, the order of the reaction is

(a) 3

(b) 0

(c) 1

(d) 2

Answer : B

### Question : In a reaction, when the concentration of reactant is increased two times, the increase in rate of reaction was four times. Order of reaction is

Answer : C			
(d) 3			
(c) 2			
(b) 1			
(a) zero			

Question : For the reaction $A + 2B \rightarrow C$ , rate is given by $R = [A] [B]^2$ then the order of the reaction is
(a) 3
(b) 6
(c) 5
(d) 7
Answer : C

### Question : The unit of rate constant for a zero order reaction is

(a) mol  $L^{-1} s^{-1}$ 

(b) L mol<sup>-1</sup> s<sup>-1</sup>

(c)  $L^2$  mol<sup>-2</sup> s<sup>-1</sup>

(d) s<sup>-1</sup>

### **Answer: A**

### Question : Which one of the following reactions is a true first order reaction?

- (a) Alkaline hydrolysis of ethyl acetate
- (b) Acid catalyst hydrolysis of ethyl acetate
- (c) Decomposition of  $N_2O$
- (d) Decomposition of gaseous ammonia on a hot platinum surface

### Answer : C

(a) Hydrogenation of ethene

(b) Natural radioactive decay of unstable nuclei

(c) Decomposition of HI on gold surface

(d) Decomposition of  $N_2O$ 

### **Answer : C**

Question : The rate law for the reaction  $xA + yB \rightarrow mP + nQ$  is Rate = k [A]c[B]d. What is the total order of the reaction?

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- (a) (x + y)
- (b) (m + n)
- (c) (c + d)
- (d) x/y

### Answer : C

Question : Half life of a first order reaction is 4s and the initial concentration of the reactant is 0.12 M. The concentration of the reactant left after 16 s is

(a) 0.0075 M

(b) 0.06 M

(c) 0.03 M

(d) 0.015 M

### **Answer : A**

Question : The reaction  $A \rightarrow B$  follows first order kinetics. The timtaken for 0.8 mole of A to produce 0.6 mole of B is 1 hour. What is the time taken for conversion of 0.9 mole of A to produce 0.675 mole of B?

(a) 2 hours

(b) 1 hour

(c) 0.5 hour

(d) 0.25 hour

### Answer : B

# Question : The rate of a first order reaction is $1.5 \times 10^{-2}$ mol L<sup>-1</sup> min<sup>-1</sup> at 0.5 M concentration of the reactant. The half life of the reaction is

(a) 0.383 min

(b) 23.1 min

(c) 8.73 min

(d) 7.53 min

### Answer : B

Question : The rate constant for a first order reaction whose half-life, is 480 seconds is :

(a) 2.88 × 10–3 sec–1

(b) 2.72 × 10–3 sec–1

(c)  $1.44 \times 10^{-3} \text{ sec}^{-1}$ 

(d) 1.44 sec<sup>-1</sup>

Answer : C

# Question : The rate constant of a reaction is $3.00 \times 10^3$ L mol<sup>-1</sup> sec<sup>-1</sup>. The order of this reaction will be:

- (b) 1
- (c) 2
- (d) 3

### Answer : C

Question : The rate constant of a first order reaction is 6.9x10<sup>-3</sup>s<sup>-1</sup> . How much time will it take to reduce the initial concentration to its 1/8th value?

(a) 100 s

(b) 200 s

(c) 300 s

(d) 400 s

### Answer : C

Question : For the reaction  $H_2(g) + Br_2(g) \rightarrow 2HBr(g)$ , the experimental data suggest, rate =  $k[H_2]$  [Br<sub>2</sub>]1/2. The molecularity and order of the reaction are respectively

(a) 2, 3/2

(b) 3/2 ,3/2

(c) 1, 1

(d) 1, 1/2

**Answer : A** 

Question : Nitrogen monoxide, NO, reacts with hydrogen, H2, according to the following equation:

 $2NO(g) + 2H_2(g) \rightarrow N_2(g) + 2H_2O(g)$ 

If the mechanism for this reaction were,

 $2NO(g) + H_2(g) \rightarrow N_2(g) + H_2O_2(g)$ ; slow

 $H2O_2(g) + H_2(g) \rightarrow 2H_2O(g)$ ; fast

Which of the following rate laws would we expect to obtain experimentally?

(a) Rate =  $k[H_2O_2][H_2]$ 

(b) Rate =  $k[NO]_2[H_2]$ 

### (c) Rate = k[NO]<sub>2</sub>[H<sub>2</sub>]<sub>2</sub>

(d) Rate =  $k[NO][H_2]$ 

Answer : C

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# Chemical Kinetics Class 12 Chemistry MCQ

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Chemical Kinetics in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

### **Chemical Kinetics MCQ Questions Class 12 Chemistry with** Answers

I. MCQ - Choose Appropriate Alternative

1. The rate of chemical reaction \_\_\_\_\_\_ with increase in concentration of the reactants. (Increases,

### **Decreases, Does not alter)**

2. Ionic reactions of inorganic compounds are \_\_\_\_\_. (very slow, moderately slow, very fast)

3. The rate of reactions can be determined. (Very Slow, Moderately Slow, Very fast)

4. The sum of exponents of the concentrations of reactants is called \_\_\_\_\_\_. (Order of reaction, **Molecularity, Equilibrium Constant)** 

5. The rate of reaction generally \_\_\_\_\_\_ in the presence of a suitable catalyst. (Increases, Decreases, remains constant)

6. The rate of a reaction \_\_\_\_\_\_ upon the temperature. (depends, slightly depends, does not depends)



<u>VBQs</u>

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7. The minimum energy required to brin Ionization energy, Energy of Activation		ed (Bond energy,
8. Oxidation of SO <sub>2</sub> in the presence of V <b>(Homogenous catalyst, Het</b>		•
9. Hydrolyses of ester in the presence o <b>Heterogeneous catalyst, Negative ca</b> t	•	Homogenous catalyst,
10. Concentration of the reactants reaction. ( <b>Increases, Decreases, Does r</b>		ng a chemical
11. Concentration of the products reaction. (Increases, Decreases, Does I		ng a chemical
12. The rate constant with ter <b>vary)</b>	nperature for a single reaction. <b>(Va</b>	ries, Slightly Varies, Does not
13. The rate of reaction at a particular ti of reaction, Instantaneous rate of rea		ate of reaction, Absolute rate
14. The specific rate constant K has <b>negligible value)</b>	value for all concentrations o	of the reactant. <b>(Fixed, Variable,</b>
15. By increasing the surface area the ra	ate of reaction can be (In	creased, Decreased, Doubled)
16. MnO <sub>2</sub> when heated with KClO <sub>3</sub> <b>catalyst)</b>	(Gives up its own oxygen, F	Produces ozone O <sub>3</sub> , Acts as
17. Reactions with high energy of activa <b>speed, slow speed)</b>	ation proceed with (High	speed, Moderately slow
18. The minimum amount of energy rec <b>(Energy of ionization, Energ</b>		
19. An inhibitor is a catalyst which	rate of reaction. (Increases, D	ecreases, Does not alter)
20 is the change of the conce Velocity Constant, Molecularity)	entration of reactant divided by the	time. (Rate of reaction,
II. Fill in the Blank		
1. The branch of chemistry, which deals	with the study of rates and mechar	nisms of chemical reactions, is

known as \_\_\_\_

2. Such reactions, which proceeds with very high velocities and are completed very quickly are called reactions.

3. Such reactions, which take place very slowly, are called \_\_\_\_\_\_ reactions.

4. Reactions between silver nitrate and sodium chloride to form white precipitates of silver chloride are an example of \_\_\_\_\_ reaction.

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5. Reactions of Organic compounds are slow and are called \_\_\_\_\_\_ reactions.

6. There are some reactions, which proceed slowly with a \_\_\_\_\_\_ speed.

7. The rate of \_\_\_\_\_ reaction can only be determined.

8. The amount of chemical change taking place in concentration of the per unit time is called \_\_\_\_\_\_ of reaction.

9. Rate of reaction is expressed in \_\_\_\_\_.

10. The rate of reaction between two specific interval of time is called \_\_\_\_\_.

11. The addition energy required to bring about a chemical reaction is called \_\_\_\_\_\_.

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12. According to \_\_\_\_\_\_ theory for a chemical reaction to take place, the reacting molecules must come closed together.

13. The addition of \_\_\_\_\_\_ helps the reaction by lowering the energy of activation.

14. The rate of reaction \_\_\_\_\_\_ with the increase in concentration of the reacting molecules.

15. When the concentration of both the reacting molecules is double, the probability of collisions between them will be \_\_\_\_\_\_ times.

16. By \_\_\_\_\_\_ the surface area of the reactants, the rate of reaction is increased.

17. Rate of reaction generally \_\_\_\_\_\_ with the rise of temperature.

18. A \_\_\_\_\_\_ is a substance, which either accelerates or retards the rate of reaction without taking part in the reaction.

19. In the preparation of Oxygen from Potassium Chlorate, \_\_\_\_\_\_ is used as catalyst.

20. In the oxidation of  $SO_2$  to  $SO_3$  by the contact process for the manufacture of  $H_2SO_4$  \_\_\_\_\_\_ is used as catalyst.

21. An unstable intermediate compound formed during a chemical reaction is called \_\_\_\_\_\_.

22. When a catalyst and the reactants are in the same phases, it is known as \_\_\_\_\_ catalyst.

23. When a catalyst and the reactants are in different phases, it is called \_\_\_\_\_\_.

24. When a catalyst increases the rate of reaction, it is called \_\_\_\_\_\_ catalyst.

25. When a catalyst retards the rate of reaction, it is called \_\_\_\_\_\_ catalyst.

26. A negative catalyst \_\_\_\_\_\_ the energy of activation, hence the rate of reaction is decreased.

27. The ratio between the rate of reaction and concentration of reactants is known as \_\_\_\_\_\_.

28. Velocity constant is independent of concentration but depends on \_\_\_\_\_\_.

29. Ionic reactions are \_\_\_\_\_ than molecular reactions.

30. The value of specific rates constant for a reaction \_\_\_\_\_\_ with time.

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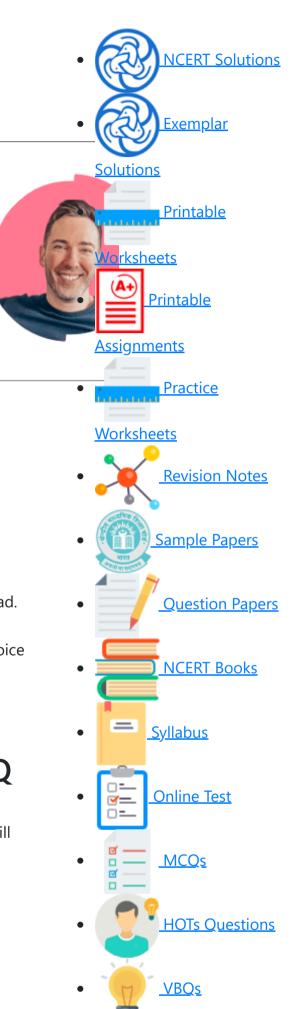
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### Chemistry in Everyday Life MCQ Questions Class 12 Chemistry with Answers

**Question : Which one of the following is employed as a tranquilizer drug?** (a) Promethazine



(b) Valium

(c) Naproxen(d) Mifepristone

Answer : B

### Question : Terfenadine is commonly used as a/an

(a) tranquilizer

(b) antihistamine

(c) antimicrobial

(d) antibiotic

### Answer : B

Question : Which one of the following is not a tranquilizer?

(a) Equanil

(b) Veronal

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- (c) Salvarsan
- (d) Serotonin

### **Answer : C**

#### Question : Tranquillizers are substances used for the treatment of

- (a) cancer
- (b) AIDS
- (c) mental diseases
- (d) physical disorders

### **Answer : C**

### Question : Which one of the following is employed as a tranquilizer drug?

- (a) Promethazine
- (b) Valium
- (c) Naproxen
- (d) Mifepristone

### **Answer: B**

### Question : Which of the following drugs is a tranquilizer and sedative

- (a) Sulphadiazine
- (b) Papaverine
- (c) Equanil
- (d) Mescaline

### **Answer : C**

### Question : Drug which helps to reduce anxiety and brings about calmness is

- (a) tranquillizer
- (b) diuretic
- (c) analgesic
- (d) antihistamine

### **Answer: A**

### Question : The drug used as an antidepressant is

- (a) Luminol
- (b) Tofranil
- (c) Mescaline
- (d) Sulphadiazine

### Answer: B

### Question : Barbituric acid and its derivatives are well known

- (a) antipyretics
- (b) analgesics
- (c) antiseptics
- (d) traquillizers

### **Answer : D**

### Question : Which of the following is a hypnotic drug?

- (a) luminal
- (b) salol
- (c) catechol
- (d) chemisol

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### Answer : A

### Question : Sulpha drugs are used for

(a) precipitating bacteria (b) removing bacteria (c) decreasing the size of bacteria (d) stopping the growth of bacteria **Answer : D** 

### **Question : Streptomycin is effective in the treatment of**

(a) tuberculosis

(b) malaria

(c) typhoid

(d) cholera

### **Answer: A**

#### Question : An antibiotic with a broad spectrum

(a) kills the antibodies

- (b) acts on a specific antigen
- (c) acts on different antigens

(d) acts on both the antigens and antibodies

#### **Answer : C**

### Question : Which of the following is not an antiseptic drug?

- (a) lodoform
- (b) Dettol
- (c) Gammexane
- (d) Genation violet

#### **Answer : C**

#### **Question : Penicillin was first discovered by**

(a) A. Fleming

- (b) Tence and Salke
- (c) S.A. Waksna
- (d) Lewis Pasteur

#### **Answer: A**

#### Question : Veronal, a barbiturate drug is used as

- (a) anaesthetic (b) sedative
- (c) antiseptic
- (d) None of these

### **Answer: B**

#### Question : A drug effective in the treatment of pneumonia, bronchitis, etc, is

(a) streptomycin (b) chloramphenicol (c) penicillin (d) sulphaguanidine **Answer : C** 

#### Question : Commonly used antiseptic 'Dettol' is a mixture of

(a) o-chlorophenozylenol + terpeneol (b) o-cresol + terpeneol (c) phenol + terpeneol (d) chloroxylenol + terpeneol **Answer : D** 

#### **Question : Chloroamphenicol is an :**

- (a) antifertility drug
- (b) antihistaminic
- (c) antiseptic and disinfectant
- (d) antibiotic-broad spectrum

### **Answer : D**

### Question : The drug which is effective in curing malaria is

(a) quinine

(b) aspirin

(c) analgin

(d) equanil

### **Answer: A**

### Question : The use of chemicals for treatment of diseases is called as

(a) isothermotherapy

(b) angiotherapy

(c) physiotherapy

### **Answer : A**

### Question : Which of the following statements is not true aboult enzyme inhibitors?

(a) Inhibit the catalytic activity of the enzyme.

(b) Prevent the binding of substrate.

(c) Generally a strong covalent bond is formed between an inhibitor and an enzyme

(d) Inhibitors can be competitive or non-competitive.

### Answer : C

### Question : Which of the following is not a target molecule for drug function in body?

- (a) Carbohydrates
- (b) Lipids
- (c) Vitamins
- (d) Proteins

### Answer : C

### Question : Which of the following comounds are administered as antacids?

- (i) Sodium carbonate
- (ii) Sodium hydrogencarbonate
- (iii) Aluminium carbonate
- (iv) Magnesium hydroxide
- (a) (i) and (ii)
- (b) (ii) and (iv)
- (c) (i), (ii) and (iii)
- (d) All of these

### Answer : B

### Question : Which of the following method of classification of drugs is useful for doctors?

- (a) On the basis of drug action.
- (b) On the basis of chemical structure.
- (c) On the basis of molecular targets.
- (d) On the basis of pharmacological effect.

### Answer : D

### Question : The function of enzymes in the living system is to

- (a) transport oxygen
- (b) provide energy
- (c) provide immunity
- (d) catalyse biochemical reactions
- Answer : D

### Question : Which one of the following is employed as a tranquilizer?

- (a) Naproxen
- (b) Tetracycline
- (c) Chlorpheninamine
- (d) Equanil
- Answer : D

### Question : Which one of the following is employed as a tranquilizer drug?

- (a) Promethazine
- (b) Valium
- (c) Naproxen

(d) Mifepristone

### Answer : B

### Question : Terfenadine is commonly used as a/an

(a) tranquilizer

(b) antihistamine

(c) antimicrobia

(d) antibiotic

### Answer : B

### Question : Which one of the following is not a tranquilizer?

(a) Equanil

(b) Veronal

- (c) Salvarsan
- (d) Serotonin

### **Answer : C**

#### Question : Tranquillizers are substances used for the treatment of

- (a) cance
- (b) AIDS
- (c) mental disease
- (d) physical disorders

### **Answer : C**

### Question : Which one of the following is employed as a tranquilizer drug?

- (a) Valium Promethazin
- (b) Equanil
- (c) Naproxen
- (d) Mifepristone

### **Answer: B**

### Question : Which of the following drugs is a tranquilizer and sedative

- (a) Sulphadiazine
- (b) Papaverine
- (c) Equanil
- (d) Mescaline

### **Answer : C**

### Question : Drug which helps to reduce anxiety and brings about calmness is

- (a) tranquillizer
- (b) diuretic
- (c) analgesic
- (d) antihistamine

### **Answer: A**

### Question : The drug used as an antidepressant is

- (a) Lumino
- (b) Tofranil
- (c) Mescaline
- (d) Sulphadiazine

### **Answer: B**

### Question : Barbituric acid and its derivatives are well known

- (a) antipyretic
- (b) analgesics
- (c) antiseptic
- (d) traquillizers

### **Answer : D**

### Question : Which of the following is a hypnotic drug

- (a) lumina
- (b) salol
- (c) catecho
- (d) chemisol

### **Answer: A**

Question : Which of the following is used for inducing sleep?

(a) Paracetamol

(b) Chloroquine

(c) Bithional

(d) Barbituric acid derivatives

### **Answer : D**

### **Question : Aspirin is**

(a) antibiotic

(b) antipyretic

(c) sedative(d) psychedelic

### Answer : B

### **Question : An antipyretic is**

(a) quinin
(b) paracetamol
(c) lumina
(d) piperazine
Answer : B

#### Question : Barbituric acid and its derivatives are well known

- (a) antipyretics
- (b) analgesics
- (c) antiseptics
- (d) traquillizers

### Answer : D

### Question : The drug used for prevention of heart attacks is

- (a) aspiri
- (b) valium
- (c) chloramphenicol
- (d) cephalsoprin

### Answer : A

### Question : Sulpha drugs are used for

- (a) precipitating bacteria
- (b) removing bacteria
- (c) decreasing the size of bacteria
- (d) stopping the growth of bacteria

### **Answer : D**

### Question : Aspirin falls under which class of drugs ?

- (a) Analgesic
- (b) Antibiotic
- (c) Antifertilit
- (d) antacid
- Answer : A

### Question : Which of the following term means pain killer

- (a) Antibiotic
- (b) Analgesic
- (c) Antipyretic
- (d) Penicillin

### Answer : B

### Question : Which one of the following can possibly be used as analgesic without causing addiction

### and mood modification ?

- (a) Diazepam
- (b) Morphine
- (c) N-Acetyl-para-aminophenol

(d) Tetrahydrocannabinol Answer : C

Question : Aspirin is known as (a) acetyl salicylic acid (b) phenyl salicylate (c) acetyl salicylate (d) methyl salicylic acid

Answer : A

Question : Which one among the following is not an analgesic?(a) Ibuprofen(b) Naproxen

(c) Aspirin (d) Valium

### **Answer : D**

#### Question : Which of the following statements about aspirin is not true?

- (a) It is effective in relieving pain.
- (b) It is a neurologically active drug.
- (c) It has antiblood clotting action.
- (d) It belongs to narcotic analgesics.

#### **Answer : D**

#### **Question : Salol can be used as**

- (a) antiseptic
- (b) antipyretic
- (c) analgesic
- (d) None of these

#### **Answer: A**

#### Question : Various phenol derivatives, tincture of iodine (2 – 3%) I<sub>2</sub> in (water / alcohol) and some dyes

- like methylene blue are
- (a) antiseptics
- (b) disinfectants
- (c) analgesics
- (d) antipyretics

### **Answer : A**

#### **Question : Sulpha drugs are used for**

- (a) precipitating bacteria
- (b) removing bacteria
- (c) decreasing the size of bacteria
- (d) stopping the growth of bacteria

#### **Answer : D**

#### **Question : Streptomycin is effective in the treatment of**

- (a) tuberculosis
- (b) malaria
- (c) typhoid
- (d) cholera
- **Answer: A**

#### Question : An antibiotic with a broad spectrum

- (a) kills the antibodies
- (b) acts on a specific antigen
- (c) acts on different antigens
- (d) acts on both the antigens and antibodies

### Answer : C

#### Question : Which of the following is not an antiseptic drug?

- (a) lodoform
- (b) Dettol
- (c) Gammexane

(d) Genation violet

### Answer : C

#### Question : Penicillin was first discovered by

(a) A. Fleming

(b) Tence and Salke

(c) S.A. Waksna

(d) Lewis Pasteur

**Answer : A** 

### Question : Veronal, a barbiturate drug is used as

(a) anaesthetic

(b) sedative

(c) antiseptic (d) None of these **Answer: B** 

#### Question : A drug effective in the treatment of pneumonia, bronchitis, etc, is

(a) streptomycin (b) chloramphenicol (c) penicillin (d) sulphaguanidine **Answer : C** 

#### Question : Commonly used antiseptic 'Dettol' is a mixture of

(a) o-chlorophenozylenol + terpeneol (b) o-cresol + terpeneol (c) phenol + terpeneol (d) chloroxylenol + terpeneol **Answer : D** 

### **Question : Chloroamphenicol is an :**

(a) antifertility drug

(b) antihistaminic

(c) antiseptic and disinfectant

(d) antibiotic-broad spectrum

**Answer : D** 

#### Question : The drug which is effective in curing malaria is

- (a) quinine
- (b) aspirin
- (c) analgin
- (d) equanil

### **Answer: A**

#### Question : An antibiotic contains nitro group attached to aromatic nucleus. It is

- (a) penicillin
- (b) streptomycin
- (c) tetracycline

(d) chloramphenicol

### **Answer : D**

#### Question : Which of the following is not a function of aspirin?

- (a) Relief from arthritic pain
- (b) Relief from postoperative pain.
- (c) Prevents platelet coagulation.
- (d) Prevention of heart attacks.

### **Answer : D**

#### Question : Arsenic drugs are mainly used in the treatment of

- (a) Jaundice
- (b) Typhoid
- (c) Syphilis
- (d) Cholera

### Answer : C

### **Question : Bithional is an example of**

(a) disinfectant

(b) antiseptic

(c) antibiotic

(d) analgesic

### **Answer: B**

### **Question : Penicillin is an :**

(a) antibiotic

(b) anaesthetic

(c) antiseptic

(d) antipyretic

### Answer : A

#### Question : Which of the following is a broad spectrum drug?

(a) Plasmoquine

(b) Chloroquine

(c) Chloramphenicol

(d) D.D.T.

#### Answer : C

### Question : Bithional is added to soap as an additive to function as a/an

(a) softener

(b) hardener

(c) dryer

(d) antiseptic

**Answer : D** 

### Question : Antiseptics and disinfectants either kill or prevent growth of microorganisms. Identify

### which of the following statements is not true:

(a) Chlorine and iodine are used as strong disinfectants.

(b) Dilute solutions of boric acid and hydrogen Peroxide are strong antiseptics.

(c) Disinfectants harm the living tissues.

(d) A 0.2% solution of phenol is an antiseptic while 1% solution acts as a disinfectant.

**Answer: B** 

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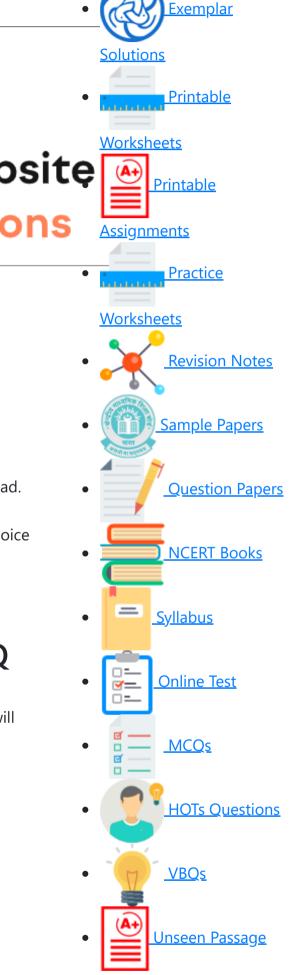
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### Coordination Compounds MCQ Questions Class 12 Chemistry with Answers

Question : The number of unpaired electrons in Ni(CO)<sub>4</sub> is



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(a) 0 (b) 1 (c) 3 (d) 5

Answer : A

**Question : The organometallic compound is :** 

(a) Ti(OCOCH<sub>3</sub>)<sub>4</sub>
(b) Ti(C<sub>2</sub>H<sub>5</sub>)<sub>4</sub>
(c) Ti(OC<sub>6</sub>H<sub>5</sub>)<sub>4</sub>
(d) Ti(OC<sub>2</sub>H<sub>5</sub>)<sub>4</sub>

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### Question : Mercuric chloride is soluble in KI solution due to :

- (a) the formation of complex ion
- (b) common iodide ion
- (c) none of the above
- (d) both (a) and (b)

### Answer : A

### **Question** :

The EAN of Zn in  $Zn(OH)_4^{2-}$  complex is:

(a) 16(b) 26(c) 36

(d) 46

### Answer : C

### Question : The reagent commonly used to determine hardness of water titrimetrically is

- (a) oxalic acid
- (b) disodium salt of EDTA
- (c) sodium citrate
- (d) sodium thiosulphate

### Answer : B

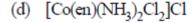
### Question : Which of the following is not considered as an organometallic compound?

- (a) cis-platin
- (b) Ferrocene
- (c) Zeise's salt
- (d) Grignard reagent

### **Answer : A**

### Question : Which of the following does not have optical isomer?

- (a)  $[Co(NH_3)_3Cl_3]$
- (b)  $[Co(en)_3]Cl_3$
- (c) [Co(en)2Cl2]Cl



### Answer : C

Question : An aqueous solution of CoCl<sub>2</sub> on addition of excess of concentrated HCl turns blue due to formation of :

Answer : C

**Question : The diamagnetic species is** 

(a) 
$$[Ni(CN)_4]^{2-}$$
 (b)  $[NiCl_4]^{2-}$   
(c)  $[CoCl_4]^{2-}$  (d)  $[CoF_6]^{2-}$ 

Question : The correct order for the wavelength of absorption in the visible region is :

(a) 
$$[Ni(NO_2)_6]^{4-} < [Ni(NH_3)_6]^{2+} < [Ni(H_2O)_6]^{2+} < [Ni(H_2O)_6]^{2+} < [Ni(H_2O)_6]^{2+} < [Ni(NH_3)_6]^{2+} < [Ni(NH_3)_6]^{2+} < [Ni(NH_3)_6]^{2+} < [Ni(NO_2)_6]^{4-} < [Ni(NO_2)_6]^{4-} < [Ni(NO_2)_6]^{2+} < [Ni(H_2O)_6]^{2+} < [Ni(H_2O)_6]^{4-} < [Ni(NO_2)_6]^{4-} <$$

### **Answer : A**

### Question : According to the postulates of Werner for coordination compounds

(a) primary valency is ionizable

(b) secondary valency is ionizable

(c) primary and secondary valencies are non-ionizable

(d) only primary valency is non-ionizable.

#### **Answer: A**

#### **Question : Which of the following postulates of Werner's theory is incorrect?**

(a) Primary valencies are satisfied by negative ions.

(b) Secondary valencies are satisfied by neutral molecules or negative ions.

(c) Secondary valence is equal to the coordination number and it depends upon the nature of ligand attached to metal.

(d) The ions/ groups bound by the secondary linkages to the metal have charecteristic spatial arrangements.

### Answer : C

### **Question : CrCl<sub>3</sub> has primary valence of**

(a) 3

(b) 4

(c) 2

(d) 1

**Answer : A** 

Question : One mole of the complex compound  $Co(NH_3)_5Cl_3$ , gives 3 moles of ions on dissolution in water. One mole of the same complex reacts with two moles of AgNO<sub>3</sub> solution to yield two moles of AgCl (s). The structure of the complex is

(a) [Co(NH<sub>3</sub>)<sub>3</sub>Cl<sub>3</sub>]. 2 NH<sub>3</sub>
(b) [Co(NH<sub>3</sub>)<sub>4</sub>Cl<sub>2</sub>] Cl . NH<sub>3</sub>
(c) [Co(NH<sub>3</sub>)<sub>4</sub>Cl] Cl<sub>2</sub>. NH<sub>3</sub>
(d) [Co(NH<sub>3</sub>)<sub>5</sub>Cl] Cl<sub>2</sub>

Answer : D

### Question : When AgNO<sub>3</sub> is added to a solution of $Co(NH_3)_5Cl_3$ , the precipitate of AgCl shows two

### ionisable chloride ions. This means :

- (a) Two chlorine atoms satisfy primary valency and one secondary valency
- (b) One chlorine atom satisfies primary as well as secondary valency
- (c) Three chlorine atoms satisfy primary valency
- (d) Three chlorine atoms satisfy secondary valency

### **Answer : A**

# Question : Which one is the most likely structure of CrCl<sub>3</sub>. 6H<sub>2</sub>O if 1/3 of total chlorine of the compound is precipitated by adding AgNO<sub>3</sub>

(a) CrCl<sub>3</sub>. 6H<sub>2</sub>O
(b) [ Cr (H<sub>2</sub>O)<sub>3</sub> Cl<sub>3</sub>]. (H<sub>2</sub>O)<sub>3</sub>
(c) [ CrCl<sub>2</sub> (H<sub>2</sub>O)<sub>4</sub> ] Cl . 2H<sub>2</sub>O
(d) [ CrCl (H<sub>2</sub>O)<sub>5</sub> ] Cl<sub>2</sub> . H<sub>2</sub>O

### Answer : C

### Question : K<sub>4</sub>[Fe(CN)<sub>6</sub>] is a :

(a) double salt(b) complex compound(c) acid(d) base

### Answer : B

Question : The number of ions formed on dissolving one molecule of FeSO <sub>4</sub> (NH <sub>4</sub> )2S	O <sub>4</sub> .6H <sub>2</sub> O in
water is:	
(a) 4	
(b) 5	
(c) 3	
(d) 6	

### Answer : B

### Question : The solution of K<sub>4</sub>[Fe(CN)<sub>6</sub>] in water will

- (a) give a test K<sup>+</sup>
- (b) give a test  $Fe^{2+}$
- (c) give a test of CN-
- (d) give a test of  $[Fe(CN)_6]^{4-}$

### Answer : A

#### Question : In the coordination compound, K<sub>4</sub>[Ni(CN)<sub>4</sub>], the oxidation state of nickel is

(a) 0

(b) +1

(c) +2

(d) –1

### Answer : A

### Question : The coordination number of a central metal atom in a complex is determined by

(a) the number of ligands around a metal ion bonded by sigma and pi-bonds both(b) the number of ligands around a metal ion bonded by pi-bonds(c) the number of ligands around a metal ion bonded by sigma bonds(d) the number of only anionic ligands bonded to the metal ion.

### Question : The oxidation state of Cr in [Cr(NH<sub>3</sub>)4Cl<sub>2</sub>]<sup>+</sup> is

(a) 0 (b) + 1 (c) + 2 (d) + 3

### **Answer : D**

Question : In Ni(CO)4 – , oxidation number of Ni is :

(a) 4 (b) – 4

(c) 0

(d) + 2

### Answer : C

### Question : [EDTA]<sup>4-</sup> is a :

(a) monodentate ligand(b) bidentate ligand(c) quadridentate ligand(d) hexadentate ligand

### **Answer : D**

### Question : The compound having the lowest oxidation state of iron is:

(a) K<sub>4</sub>Fe(CN)<sub>6</sub>
(b) K<sub>2</sub>FeO<sub>4</sub>
(c) Fe<sub>2</sub>O<sub>3</sub>
(d) Fe(CO)<sub>5</sub>

### Answer : D

### Question : The coordination number and the oxidation state of the element 'E' in the complex [E

(en)2 (C2O4)]NO2 (where (en) is ethylene diamine) are, respectively,

- (a) 6 and 2
- (b) 4 and 2
- (c) 4 and 3
- (d) 6 and 3

# Question : Some salts although containing two different metallic elements give test for only one of them in solution. Such salts are

(a) complex

(b) double salts

(c) normal salts

(d) None of these

Answer : A

**Question : Coordination number of Ni in [Ni(C<sub>2</sub>O<sub>4</sub>)<sub>3</sub>]<sup>4–</sup> is** (a) 3 (b) 6

### **Answer: B**

### Question : According to Lewis, the ligands are

- (a) acidic in nature
- (b) basic in nature
- (c) some are acidic and others are basic
- (d) neither acidic nor basic

### **Answer: B**

### Question : Ligand in a complex salt are

(a) anions linked by coordinate bonds to a central metal atom or ion (b) cations linked by coordinate bonds to a central metal or ion (c) molecules linked by coordinate bonds to a central metal or ion (d) ions or molecules linked by coordinate bonds to a central atom or ion

### Answer : C

### Question : The ligand N(CH<sub>2</sub>CH<sub>2</sub>NH<sub>2</sub>)<sub>3</sub> is

- (a) tridentate
- (b) pentadentate
- (c) tetradentate
- (d) bidentate

### Answer : C

### Question : An example of ambidentate ligand is

- (a) Ammine
- (b) Aquo
- (c) Chloro
- (d) Thiocyanato

### **Answer : D**

### Question : Which of the following does not form a chelate ?

- (a) EDTA
- (b) Oxalate
- (c) Pyridine
- (d) Ethylenediamine

#### Answer : C



### **Question : A** *bidenate* **ligand always**

(a) has bonds formed to two metals ions

(b) has a charge of +2 or -2

(c) forms complex ions with a charge of +2 or -2

(d) has two donor atoms forming simultaneously two sigma ( $\Sigma$ ) bonds.

### **Answer: B**

### Question : An ambident ligand is one which

(a) is linked to the metal atom through two donor atoms

(b) has two donor atoms, but only one of them has the capacity to form a coordinate bond [or a sigma ( $\Sigma$ ) bond]

(c) has two donor atoms, but either of two can form a coordinate bond

(d) forms chelate rings.

### Answer : C

### Question : NH<sub>2</sub>-NH<sub>2</sub> serves as

(a) Monodentate ligand

(b) Chelating ligand

(c) Bridging ligand

(d) Both (a) and (c)

### Answer : C

### Question : Which one of the following is NOT a ligand ?

(a) PH<sub>3</sub> (b) NO+

(c) Na+

(d) F–

Answer : C

### Question : The IUPAC name of K<sub>3</sub>[Ir(C<sub>2</sub>O<sub>4</sub>)<sub>3</sub>] is

(a) potassium trioxalatoiridium (III)

(b) potassium trioxalatoiridate (III)

(c) potassium tris (oxalato) iridium (III)

(d) potassium tris (oxalato) iridate (III)

### Answer : B

### **Question : Which one does not belong to ligand?**

(a) PH<sub>3</sub>
(b) NO+
(c) BF<sub>3</sub>

(d) Cl–

### Answer : C

### Question : Which ligand is expected to be bidentate?

(a) 2 C<sub>2</sub>O<sup>2</sup><sub>4-</sub>

(b)  $CH_3C = N$ (c) Br -(d)  $CH_3NH_2$ 

Answer : A

### Question : Which one of the following ligands forms a chelate

(a) Acetate

(b) Oxalate

(c) Ammonia

(d) Cyanide

### Answer : B

### Question : Choose the correct statement.

(a) Coordination number has nothing to do with the number of groups or molecules attached to the central atom

(b) Coordination number is the number of coordinating sites of all the ligands connected to the central

atom or the number of coordinate bonds formed by the metal atom with ligands

(c) Werner's coordination theory postulates only one type of valency

(d) All the above are correct

### Answer : B

### Question : O<sub>2</sub> is a

(a) Monodentate ligand

(b) Bidenate ligand

(c) Tridentate ligand

(d) Hexadenate ligand

### **Answer: B**

### Question : The stabilisation of cooordination compounds due to chelation is called the chelate effect. Which of the following is the most stable complex species ?

(a) [Fe(CO)<sub>5</sub>]
(b) 3[Fe(CN)<sub>6</sub>] (c) 3[Fe(C<sub>2</sub>O<sub>4</sub>)<sub>3</sub>] (d) 3[Fe(H<sub>2</sub>O)<sub>6</sub>] -

### Answer : C

# Question : A chelating agent has two or more than two donor atoms to bind to a single metal ion. Which of the following is not a chelating agent ? (a) thiosulphato (b) oxalato (c) glycinato

(d) ethane - 1, 2-diamine

### Answer : A

### Question : Which of the following species is not expected to be a ligand?

(a) NO (b) NH<sub>4</sub> (c) NH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>NH<sub>2</sub> (d) Both (a) and (b)

### Answer : B

Question : Which of the following complexes are homoleptic ? (i)  $3[Co(NH_3)_6]^{3+}$  (ii)  $[Co(NH_3)4Cl]^+$ (iii)  $2[Ni(CN)4]^{2-}$  (iv)  $[Ni(NH_3)4Cl_2]$ (a) (i) and (ii) (b) (ii) and (iii) (c) (iii) and (iv) (d) (i) and (iii)

Answer : C

### Question : Which of the following complexes are heteroleptic ?

(i) 3[Cr(NH<sub>3</sub>)<sub>6</sub>]<sup>3+</sup> (ii) [Fe(NH<sub>3</sub>)<sub>4</sub>Cl<sub>2</sub>]<sup>+</sup>
(iii) 4[Mn(CN)<sub>6</sub>]<sup>2-</sup> (iv) [Co(NH<sub>3</sub>)4Cl<sub>2</sub>]
(a) (i), (iv)
(b) (ii) and (iv)
(c) (i) and (ii)
(d) (i) and (iv)

### **Answer : B**

#### Question : Central atoms/ions in coordination compounds are.

(a) Lewis acid

(b) Lewis bases

(c) Neutral molecules

(d) All of these

### **Answer : A**

#### Question : What is the denticity of the ligand ethylenediaminetetra actetate ion?

(a) 4

(b) 2

(c) 6

(d) 1

### Answer : C

### Question : K<sub>3</sub>[Al(C<sub>2</sub>O<sub>4</sub>)<sub>3</sub>] is called

(a) Potassium aluminooxalate

(b) Potassium trioxalatoaluminate (III)

(c) Potassium aluminium (III) oxalate

(d) Potassium trioxalatoaluminate (VI)

### Answer : C

### Question : The hypothetical complex chlorodiaquatriamminecobalt (III) chloride can be represented

as

(a) [CoCl(NH<sub>3</sub>)<sub>3</sub>(H<sub>2</sub>O)<sub>2</sub>]Cl<sub>2</sub>
(b) [Co(NH<sub>3</sub>)<sub>3</sub>(H<sub>2</sub>O)Cl<sub>3</sub>]
(c) [Co(NH<sub>2</sub>)<sub>3</sub>(H2O)2 Cl]
(d) [Co(NH<sub>3</sub>)<sub>3</sub>(H<sub>2</sub>O)<sub>3</sub>]Cl<sub>3</sub>

### **Answer : A**

#### Question : The IUPAC name of the coordination compound K<sub>3</sub>[Fe(CN)<sub>6</sub>] is

(a) Tripotassium hexacyanoiron (II)

(b) Potassium hexacyanoiron (II)

(c) Potassium hexacyanoferrate (III)

(d) Potassium hexacyanoferrate (II)

### Answer : C

### Question : The IUPAC name for the complex [Co(ONO)(NH<sub>3</sub>)<sub>5</sub>]Cl<sub>2</sub> is

(a) pentaamminenitrito-N-cobalt(II) chloride
(b) pentaamminenitrito-N-cobalt(III) chloride
(c) nitrito-N-pentaamminecobalt(III) chloride
(d) nitrito-N-pentaamminecobalt(II) chloride

#### Question : The IUPAC name of K<sub>2</sub>[PtCl<sub>6</sub>] is

(a) hexachloroplatinate potassium

(b) potassium hexachloroplatinate (IV)

(c) potassium hexachloroplatinate

(d) potassium hexachloroplatinum (IV)

#### **Answer: B**

### Question : The IUPAC name of [Ni(NH<sub>3</sub>)<sub>4</sub>] [NiCl<sub>4</sub>] is

- (a) Tetrachloronickel (II) tetraamminenickel (II)
- (b) Tetraamminenickel (II) tetrachloronickel (II)
- (c) Tetraamminenickel (II) tetrachloronickelate (II)
- (d) Tetrachloronickel (II) tetrachloronickelate (0)

#### Answer : C

#### Question : As per IUPAC nomenclature, the name of the complex $[Co(H_2O)_4(NH_3)_2]Cl_3$ is :

- (a) Tetraaquadiaminecobalt (III) chloride
- (b) Tetraaquadiamminecobalt (III) chloride
- (c) Diaminetetraaquacoblat (II) chloride
- (d) Diamminetetraaquacobalt (III) chloride

#### Answer : D

#### Question : The IUPAC name of the complex [Co(NH<sub>3</sub>)<sub>4</sub>(H<sub>2</sub>O)Cl]Cl<sub>2</sub> is

- (a) aquatetramminechloridocobalt (III) chloride
- (b) chloridoaquatetramminechloridocobalt (III) chloride
- (c) chloridoaquatetramminechloridocobalt (III) chloride
- (d) tetrammineaquachloridocobalt (III) chloride

#### Answer : D

#### Question : As per IUPAC nomenclature, the name of the complex $[Co(H_2O)_4(NH_3)_2]Cl_3$ is :

- (a) Tetraaquadiaminecobalt (III) chloride
- (b) Tetraaquadiamminecobalt (III) chloride
- (c) Diaminetetraaquacobalt (II) chloride
- (d) Diamminetetraaquacobalt (III) chloride

#### **Answer : D**

### Question : Chemical formula for iron (III) hexacyanoferrate (II) is

(a)  $Fe[Fe(CN)_6]$ (b)  $Fe_3[Fe(CN)_6]$ (c)  $Fe_3[Fe(CN)_6]_4$ 

(d)  $Fe_4[Fe(CN)_6]_3$ 

Answer : A

### Question : The ligands in anti-cancer drug cis-platin are (a) NH<sub>3</sub>, Cl (b) NH<sub>3</sub>, H<sub>2</sub>O (c) Cl, H<sub>2</sub>O (d) NO. Cl

(d) NO, Cl

**Answer : A** 

#### **Question : Which statement is true for ferrocene?**

(a) All Fe-C are of equal length

- (b) It has sandwich type structure
- (c) It was the first discovered organometallic compound
- (d) All of these.

#### Answer : D

Question : During estimation of nickel, we prepare nickel dimethylglyoxime, a scarlet red solid. This

- compound is \_\_\_\_\_.
- (a) ionic
- (b) covalent
- (c) metallic
- (d) non-ionic complex.

#### Answer : D

Question : Which of the following metal ions will form complexes with the same magnetic moment and geometry irrespective of the nature of ligands?

(a) Ni<sup>2+</sup>
(b) Fe<sup>2+</sup>
(c) Cu<sup>2+</sup>
(d) Co<sup>2+</sup>

### Answer : C

#### **Question** :

[Fe(NO<sub>2</sub>)<sub>3</sub>Cl<sub>3</sub>] and [Fe(ONO)<sub>3</sub>Cl<sub>3</sub>] shows

(a) linkage isomerism

(b) geometrical isomerism

(c) optical isomerism

(d) none of the above

#### Answer: B

Question : Among the following the compound that is both paramagnetic and coloured is

(a)  $K_2Cr_2O_7$  (b)  $(NH_4)_2[TiCl_6]$ 

(c) VOSO4 (d) K<sub>3</sub>Cu(CN)<sub>4</sub>

**Answer : A** 

Question : Which of the following complex has zero magnetic moment (spin only)?

(a)  $[Ni(NH_3)_6]Cl_2$  (b)  $Na_3[FeF_6]$ (c)  $[Cr(H_2O)_6]SO_4$  (d)  $K_4[Fe(CN)_6]$ 

Answer : C

### Question : Silver chloride dissolves in (a) Water (b) Conc. HCl

(c) NH<sub>4</sub>OH (d) CCl<sub>4</sub>

### Answer : C

### Question : The IUPAC name of the complex Hg [Co(CNS )<sub>4</sub>] is

(a) mercury tetrathiocyanatocobaltate (II)

- (b) mercury cobalttetrasulphocyano (II)
- (c) mercury tetrasulphocyanidecobalt (II)
- (d) tetrasulphocyantocobalt mercurate (II)

### Answer : A

Question : In the isoelectronic series of metal carbonyl, the CO bond strength is expected to increase in the order:

- (a)  $[Mn(CO)_6]^+ < [Cr(CO)_6] < [V(CO)_6]^-$
- $(d) \quad [V(CO)_6]^- < [Cr(CO)_6] < [Mn(CO)_6]^+$
- (c)  $[V(CO)_6]^- < [Mn(CO)_6]^+ < [Cr(CO)_6]^+$
- $(d) \quad [Cr(CO)_6] < [Mn(CO)_6]^+ < [V(CO)_6]^-$

### Answer : B

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### Electrochemistry Class 12 Chemistry MCQ

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Electrochemistry in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

### **Electrochemistry MCQ Questions Class 12 Chemistry with** Answers

**Question : Effect of dilution on conductivity of solution:** 



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a) Increases

b) Decreases c) Unchanged d) None of the above

### Answer: A

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Question : Through a solution of CuSO<sub>4</sub> a current of 3 amperes was passed for 2 hours. At cathode 3 g of  $Cu^{2+}$ ions were discharged. The current efficiency is [At. wt. of Cu = 63.5]

a) 33.3%

b) 42.2%

c) 48.7%

d) 54.4%

#### Answer : B

**Question : Which shows electrical conductance?** 

- a) Sodium
- b) Diamond
- c) Potassium
- d) Graphite

#### Answer : D

Question : Which of the following reactions is used to make a fuel cell?

(a) 
$$Cd(s) + 2Ni(OH)_{3}(s) \longrightarrow CdO(s)$$
  
+  $2Ni(OH)_{2}(s) + H_{2}O(l)$   
(b)  $Pb(s) + PbO_{2}(s) + 2H_{2}SO_{4}(aq) \longrightarrow$   
 $2PbSO_{4}(s) + 2H_{2}O(l)$ 

(c) 
$$2H_2(g) + O_2(g) \longrightarrow 2H_2O(l)$$
  
(d)  $2Fe(s) + O_2(g) + 4H^+(aq) \longrightarrow$ 

$$2Fe^{2+}(aq) + 2H_2O(l)$$

### Answer : C

### Question : Time required to deposit one millimole of aluminium metal by the passage of 9.65 amperes through aqueous solution of aluminium ion is

a) 30 s
b) 10 s
c) 30,000 s
d) 10,000 s

#### **Answer : A**

Question : The products formed when an aqueous solution of NaBr is electrolysed in a cell having inert electrodes are:

a) Na and Br<sub>2</sub>
b) Na and O<sub>2</sub>
c) H<sub>2</sub> ,Br<sub>2</sub> and NaOH
d) H<sub>2</sub> and O<sub>2</sub>

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Question : Equivalent conductance of an electrolyte containing NaF at infinite dilution is 90.1 Ohm<sup>-1</sup> cm<sup>2</sup>. If NaF is replaced by KF what is the value of equivalent conductance?

a) 90.1 Ohm<sup>-1</sup>cm<sup>2</sup>
b) 111.2 Ohm<sup>-1</sup>cm<sup>2</sup>
c) 0
d) 222.4 Ohm<sup>-1</sup>cm<sup>2</sup>

### Answer : A

Question : For a cell reaction involving two electron change, the standard EMF of the cell is 0.295 V at 2°C. The equilibrium constant of the reaction at 25°C will be

a) 29.5 × 10<sup>-2</sup> b) 10 c) 1 × 10<sup>10</sup> d) 2.95 × 10<sup>-10</sup>

### Answer : C

Question : A 0.5 M NaOH solution offers a resistance of 31.6 ohm in a conductivity cell at room temperature. What shall be the approximate molar conductance of this NaOH solution if cell constant of the cell is 0.367 cm<sup>-1</sup>.

a) 234 S cm<sup>2</sup> mole<sup>-1</sup>
b) 23.2 S cm<sup>2</sup> mole<sup>-1</sup>
c) 4645 S cm<sup>2</sup> mole<sup>-1</sup>
d) 5464 S cm<sup>2</sup> mole<sup>-1</sup>

**Answer : D** 

#### Question : Which of the following statements is incorrect regarding electrochemistry?

a) It is the study of production of electricity from energy released during spontaneous chemical reactions.

b) NaOH, Cl<sub>2</sub>, alkali and alkaline earth metals are prepared by electrochemical methods.

c) The demerit associated with electrochemical methods is that they are more polluting. Thus they are ecodestructive.

d) Electrochemical reactions are more energy efficient and less polluting.

### Answer : C

### Question : What flows in the internal circuit of a galvanic cell?

a) lons

b) Electrons

- c) Electricity
- d) Atoms

### **Answer : A**

#### Question : Which of the following statements about galvanic cell is incorrect

a) anode is positive

b) oxidation occurs at the electrode with lower reduction potential

c) cathode is positive

d) reduction occurs at cathode

Answer : A

Question : Reaction that takes place at graphite anode in dry cell is

a)  $Zn^{2+} + 2e + \rightarrow Zn(s)$ 

b)  $Zn(s) \rightarrow Zn^{2+} + 2e^{-1}$ 

c) Mn2+ + 2<sup>e</sup>-  $\rightarrow$ Mn(s)

d) Mn(s)→Mn- + e- +1.5V

#### Question : In which of the following conditions salt bridge is not required in a galvanic cell?

a) When galvanic cell is used in geyser.

b) When distance between oxidation half cell and reduction half cell is negligible.

c) Electrolytic solutions used in both the half cells are of same concentration.

d) When both the electrodes are dipped in the same electrolytic solution.

#### Answer : D

### Question : Which device converts chemical energy of a spontaneous redox reaction into electrical energy?

- a) Galvanic cell
- b) Electrolytic cell
- c) Daniell cell
- d) Both a) and c)

#### Answer : D

#### Question : To deposit one equivalent weight of silver at cathode, the charge required will be

- a) 9.65 × 10<sup>4</sup> C
- b) 9.65 × 10<sup>3</sup> C
- c) 9.65 × 10<sup>5</sup> C
- d) 9.65 × 10<sup>7</sup> C

### **Answer : A**

### Question : The volume of oxygen gas liberated at NTP by passing a current of 9650 coulombs through acidified water is

- a) 1.12 litre
- b) 2.24 litre
- c) 11.2 litre
- d) 22.4 litre

Question : Three faradays electricity was passed through an aqueous solution of iron (II) bromide. The weight of iron metal (at. wt = 65) deposited at the cathode (in gm) is

a) 56

b) 84

c) 112

d) 168

#### Question : Faraday's laws of electrolysis will fail when

- a) temperature is increased
- b) inert electrodes are used
- c) a mixture of electrolytes is used
- d) None of these cases

### **Answer : D**

### Question : The electric charge for electrode decomposition of one gram equivalent of a substance is

- a) one ampere per second
- b) 96500 coulombs per second
- c) one ampere for one hour
- d) charge on one mole of electrons

#### **Answer : D**

### Question : In electrolysis of dilute H<sub>2</sub>SO<sub>4</sub> using platinum electrodes

- a)  $H_2$  is evolved at cathode
- b) NH<sub>2</sub> is produced at anode
- c) Cl<sub>2</sub> is obtained at cathode
- d)  $O_2$  is produced

### **Answer: A**

### Question : The difference between the electrode potentials of two electrodes when no current is drawn hrough the cell iscalled \_\_\_\_\_.

- a) Cell potentials
- b) Cell emf
- c) Potential difference
- d) Cell voltage

#### Answer : B

### Question : Which of the following pair(s) is/are incorrectly matched?

(i) R (resistance) – ohm ( $\mathbf{\Omega}$ )

(ii) (resistivity) – ohm metre ( $\Omega$ m)

(iii) G (conductance) – seimens or ohm (S)

(iv) (conductivity) – seimens metre<sup>-1</sup> (Sm<sup>-1</sup>)

a) (i), (ii) and (iii)

b) (ii) and (iii)

**c)** (i), (ii) and (iv)

d) (iii) only

### Answer : C

### **Question : The reference electrode is made by using**

a) ZnCl<sub>2</sub>

b) CuSO<sub>4</sub>

c) HgCl<sub>2</sub>

d) Hg<sub>2</sub>Cl<sub>2</sub>

### **Answer : D**

### Question : The standard hydrogen electrode potential is zero, because

- a) hydrogen oxidized easily
- b) electrode potential is considered as zero
- c) hydrogen atom has only one electron
- d) hydrogen is a very light element

### Answer : B

### Question : Without losing its concentration ZnCl<sub>2</sub> solution cannot be kept in contact with

- a) Au
- b) Al
- c) Pb
- d) Ag

### Answer : B

### Question : On the basis of the following $\mathbf{E}^\circ$ values, the strongest oxidizing agent is :

```
[Fe(CN)_6]^{4-} + [Fe(CN)_6]^{3-} + e-; E^\circ = -0.35 V
```

Fe2+  $\rightarrow$  Fe3+ + e-; E° = -0.77 V

### a) [Fe(CN)<sub>6</sub>]<sup>4–</sup>

b) Fe<sup>2+</sup>

c) Fe<sup>3+</sup>

d) [Fe(CN)<sub>6</sub>]<sup>3-</sup>

Answer :C

Question : Standard electrode potential of three metals X, Y and Z are – 1.2 V, + 0.5 V and – 3.0 V, espectively. The reducing powerbof these metals will be :

a) Y > Z > X

b) X > Y > Z

c) Z > X > Y d) X > Y > Z

### Answer : C

Question : Standard electrode potential for  $Sn^{4+} / Sn^{2+}$  couple is + 0.15 V and that for the  $Cr^{3+} / Cr$  couple is – 0.74 V. These two couples in their standard state are connected to make a cell. The cell potential will be

- a) + 1.19 V
- b) + 0.89 V
- c) + 0.18 V
- d) + 1.83 V

#### **Answer : B**

#### Question : Standard reduction potentials of the half reactions are given below :

- $F_2(q) + 2e \rightarrow 2F (aq); E^\circ = + 2.85 V$
- $Cl_2(q) + 2e \rightarrow 2Cl (aq); E^\circ = + 1.36 V$
- $Br_2(l) + 2e \rightarrow 2Br(aq); E^\circ = + 1.06 V$
- $I_2(s) + 2e \rightarrow 2I (aq); E^\circ = + 0.53 V$

The strongest oxidising and reducing agents respectively are

- a)  $F_2$  and  $I_-$
- b) Br<sub>2</sub> and Cl-
- c) Cl<sub>2</sub> and Br–
- d)  $Cl_2$  and  $l_2$

### Answer : A

### Question : A button cell used in watches functions as following

Zn(s) + Ag2O(s) + H2O(l)

 $2Ag(s) + Zn_2 + (aq) + 2OH - (aq)$ 

### If half cell potentials are :

```
Zn_2+(aq) + 2e \rightarrow Zn(s); Eo = -0.76 V
```

```
Ag_2O(s) + H_2O(l) + 2e \rightarrow 2Ag(s) + 2OH - (aq); Eo = 0.34 V
```

The cell potential will be :

a) 0.42 V

b) 0.84 V

c) 1.34 V

d) 1.10 V

Answer : D

# Question : The oxidation potentials of A and B are +2.37 and +1.66 V respectively. In chemical reactions

- a) A will be replaced by B
- b) A will replace B
- c) A will not replace B
- d) A and B will not replace each other

### Answer : B

#### Question : A smuggler could not carry gold by depositing iron on the gold surface since

- a) gold is denser
- b) iron rusts
- c) gold has higher reduction potential than iron
- d) gold has lower reduction potential than iron

### Answer : B

### Question : Which cell will measure standard electrode potential of copper electrode ?

a) Pt (s) |H<sub>2</sub> (g, 0.1 bar) |H+ (aq., 1 M) ||Cu<sup>2+</sup> (aq., 1 M) | Cu b) Pt (s) |H<sub>2</sub> (g, 1 bar) |H+ (aq., 1 M) ||Cu<sup>2+</sup> (aq., 2 M) | Cu c) Pt (s) |H<sub>2</sub> (g, 1 bar) |H+ (aq., 1 M) ||Cu<sup>2+</sup> (aq., 1 M) | Cu d) Pt (s) |H<sub>2</sub> (g, 1 bar) |H+ (aq., 0.1 M) ||Cu<sup>2+</sup> (aq., 1 M) | Cu

### Answer : C

### Question : Which of the following statement is not correct about an inert electrode in a cell ?

- a) It does not participate in the cell reaction.
- b) It provides surface either for oxidation or for reduction reaction.
- c) It provides surface for conduction of electrons
- d) It provides surface for redox reaction.

### **Answer : D**

### Question : In the electrochemical reaction $2Fe^{3+} + Zn \rightarrow Zn^{2+} + 2Fe^{2+}$ , on increasing the

concentration of Fe<sup>2+</sup>

a) increases cell emf

b) increases the current flow

c) decreases the cell emf

d) alters the pH of the solution

Answer : C

Question : The standard e.m.f. of a galvanic cell involving cell reaction with n = 2 is found to be
0.295 V at 25°C. The equilibrium constant of the reaction would be (Given F = 96500 C mol-1; R =
8.314JK–1mol <sup>–1</sup> )

- a) 2.0x1011
- b) 4.0x1012
- c) 1.0x102
- d) 1.0x10<sup>10</sup>

### Answer : D

Question : On electrolysis of dilute sulphuric acid using platinum electrodes, the product obtained at the anode will be

- a) hydrogen
- b) oxygen
- c) hydrogen sulphide
- d) Sulphur dioxide

#### **Answer : B**

Question : If 0.5 amp current is passed through acidified silver nitrate solution for 100 minutes. The mass of silver deposited oncathode, is (eq.wt.of silver nitrate = 108)

- a) 2.3523 g
- b) 3.3575 g
- c) 5.3578 g
- d) 6.3575 g

### Answer : B

# Question : Molar ionic conductivities of a two-bivalent electrolytes $x^{2+}$ and $y^{2-}$ are 57 and 73 respectively. The molar conductivity of the solution formed by them will be

- a) 130 S cm<sup>2</sup> mol<sup>-1</sup>
- b) 65 S cm<sup>2</sup> mol<sup>-1</sup>
- c) 260 S cm<sup>2</sup> mol<sup>-1</sup>
- d) 187 S cm<sup>2</sup> mol<sup>-1</sup>

### Answer : A

Question : The standard emf of a cell, involving one electron change is found to be 0.591 V at 25°C. The equilibrium constant of the reaction is (F = 96500 C mol<sup>-1</sup>)

a) 1.0 × 10<sup>1</sup>

b) 1.0 × 10<sup>5</sup>

c) 1.0 × 10<sup>10</sup>

d) 1.0 ×10<sup>30</sup>

### Answer : C

### **Question : For the galvanic cell**

### Zn | Zn2+ (0.1M) || Cu<sup>2+</sup> (1.0M)|Cu the cell potential increase if:

- a) [Zn2+] is increased
- b) [Cu2+] is increased
- c) [Cu2+] is decreased
- d) surface area of anode is increased

### Answer: B

### Question : Which of the following is the use of electrolysis?

- a) Electrorefining
- b) Electroplating
- c) Both a) & b)
- d) None of these

### Answer : D

Question : An electrolytic cell contains a solution of  $Ag_2SO_4$  and has platinum electrodes. A current is passed until 1.6 gm of  $O_2$  has been liberated at anode. The amount of silver deposited at cathode would be

- a) 107.88 gm
- b) 1.6 gm
- c) 0.8 gm
- d) 21.60 gm

### Answer : D

### Question : When 9650 coulombs of electricity is passed through a solution of copper sulphate, the amount of copper deposited is (given at. wt. of Cu = 63.6)

a) 0318g

b) 3.18 g

c) 31.8g

. . . .

### Answer : B

Question : Find the charge in coulombs required to convert 0.2 mole VO3  $^{-2}$  into VO4  $^{-3}$   $^{-1}$ 

a) 1.93 × 10<sup>4</sup>

b) 9.65 × 10<sup>4</sup>

c) 1.93 × 10<sup>5</sup>

d) 9.65 × 10<sup>5</sup>

### Answer : A

# Question : A silver cup is plated with silver by passing 965 coulombs of electricity. The amount of Ag deposited is :

- a) 107.89 g
- b) 9.89 g
- c) 1.0002 g
- d) 1.08 g

### Answer : D

### Question : The number of coulombs required to reduce 12.3 g of nitrobenzene to aniline is :

- a) 115800 C
- b) 5790 C
- c) 28950 C
- d) 57900 C

### Answer : D

### Question : The amount of electricity that can deposit 108 g of Ag from AgNO<sub>3</sub> solution is:

a) 1 F

- b) 2 A
- c) 1 C
- d) 1 A

### Answer : A

Question : If 0.01 M solution of an electrolyte has a resistance of 40 ohms in a cell having a cell
constant of 0.4 cm <sup>-1</sup> , then its molar conductance in ohm <sup>-1</sup> cm <sup>2</sup> mol <sup>-1</sup> is

- a) 10<sup>2</sup>
- b) 10<sup>4</sup>
- c) 10
- d) 10<sup>3</sup>

#### Answer : D

Question : Specific conductance of a 0.1 N KCl solution at 23°C is 0.012 ohm<sup>-1</sup> cm<sup>-1</sup>. Resistance of cell containing the solutionat same temperature was found to be 55 ohm. The cell constant is

a) 0.0616 cm<sup>-1</sup>

b) 0.66 cm<sup>-1</sup>

c) 6.60 cm<sup>-1</sup>

d) 660 cm<sup>-1</sup>

Answer : B

### **Question : The unit of equivalent conductivity is**

### a) ohm cm

- b) ohm-1 cm<sup>2</sup> (g equivalent)-1
- c) ohm cm<sup>2</sup> (g equivalent)

d) S cm<sup>-2</sup>

### Answer : B

Question : The resistance of 0.01 N solution of an electrolyte was found to be 220 ohm at 298 K using a onductivity cell with a cell constant of 0.88cm–1. The value of equivalent conductance of solution is –

- a) 400 mho  $\mathrm{cm}^2\,\mathrm{g}\,\mathrm{eq}^{-1}$
- b) 295 mho  $\text{cm}^2$  g  $\text{eq}^{-1}$
- c) 419 mho cm<sup>2</sup> g eq–1
- d) 425 mho cm2 g  $eq^{-1}$

### Answer : A

# Question : Specific conductance of 0.1 M HNO<sub>3</sub> is $6.3 \times 10^{-2}$ ohm<sup>-1</sup> cm<sup>-1</sup>. The molar conductance of the solution is

- a) 100 ohm<sup>-1</sup> cm<sup>2</sup>
- b) 515 ohm<sup>-1</sup> cm<sup>2</sup>
- c) 630 ohm<sup>-1</sup> cm<sup>2</sup>
- d) 6300 ohm<sup>-1</sup> cm<sup>2</sup>

### Answer : C

# Question : The specific conductance of a 0.1 N KCl solution at 23°C is 0.012 ohm–1cm<sup>-1</sup>. The resistance of cell containing the solution at the same temperature was found to be 55 ohm. The cell constant will be

- a) 0.142 cm<sup>-1</sup>
- b) 0.66 cm<sup>-1</sup>
- c) 0.918 cm<sup>-1</sup>
- d) 1.12 cm<sup>-1</sup>

### Answer : B

Question : The unit of specific conductivity is

a) ohm cm<sup>-1</sup>

b) ohm cm<sup>-2</sup>

c) ohm<sup>-1</sup> cm

d)  $ohm^{-1} cm^{-1}$ 

Answer : D

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# Question : Which of the following solutions of KCI will have the highest value of specific conductance?

- a) 1.0 N
- b) 0.1 N
- c) 1.0 ×10<sup>-2</sup>N
- d) 1.0 ×10<sup>-3</sup>N

### Answer : A

### Question : The cell constant of a conductivity cell \_\_\_\_\_.

- a) changes with change of electrolyte.
- b) changes with change of concentration of electrolyte.
- c) changes with temperature of electrolyte.
- d) remains constant for a cell.

### Answer : D

### Question : The standard reduction potential at 298K for the following half cells are given: Which is the strongest reducing agent?

- a) Zn(s)
- b) Cr(s)
- c) Both
- d) None of these
- Answer : A

### Question : Faraday's laws of electrolysis are related to the

- a) Atomic number of the cation
- b) Atomic number of the anion
- c) Equivalent weight of the electrolyte
- d) Speed of the cation

### Answer : C

### Question : Which of the following statement is wrong about galvaniccell ?

- a) cathode is positive charged
- b) anode is negatively charged
- c) reduction takes place at the anode
- d) reduction takes place at the cathode

### Answer : C

### Question : In the electrolytic cell, flow of electrons if from

a) Cathode to anode in solution

b) Cathode to anode through external supply

c) Cathode to anode through internal supply

d) Anode to cathode through internal supply

### Answer : C

Question : metal having negative reduction potential when dippedin the solution of its own ions, has a tendency :

a) to pass into the solution

b) to be deposited from the solution

c) to become electrically positived) to remain neutra

#### **Answer : A**

Question : A hydrogen electrode is immersed in a solution with pH = 0 (HCl). By how much will the potential (reduction) change if an equivalent amount of NaOH is added to the solution. (Take H<sub>2</sub> p = 1 atm), T = 298 K.

a) increase by 0.41 V b) increase by 59 mV c) decrease by 0.41 V d) decrease by 59 mV

### Answer : A

#### Question : During the electrolysis of fused NaCl, the reaction thatoccurs at the anode is :

- a) Chloride ions are oxidized
- b) Chloride ions are reduced
- c) Sodium ions are oxidized
- d) Sodium ions are reduced

### Answer : A

### Question : In electroplating the article to be electroplated is made :

- a) cathode
- b) anode
- c) either cathode or anode
- d) simply suspended in the electrolytic bath.

#### **Answer : A**

### Question : The electric charge for electrode deposition of one gram equivalent of a substance is

- a) One ampere per second
- b) 96.500 coulombs per second
- c) One on one mole of electrons
- d) Charge on one mole of electrons

### Answer : D

### Question : When a lead storage battery is discharged

a) SO<sub>2</sub>

- b) Lead is formed
- c) Lead sulphate is consumed
- d) Sulphuric acid is consumed

### Answer : D

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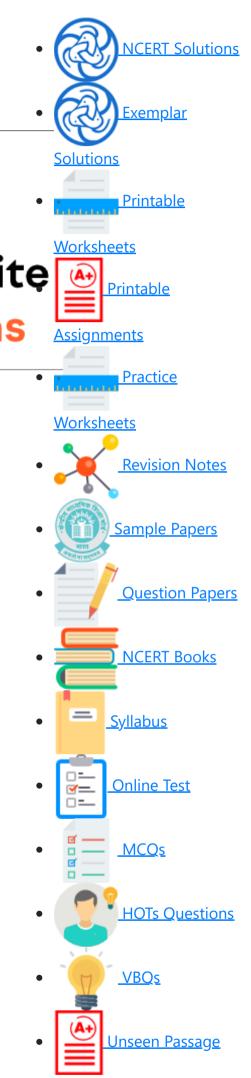
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(a) Ag, Cu and Pb (c) Ag, Cu and Sn (b) Ag, Mg and Pb (d) Al, Cu and Pb

#### Answer: A

Question : Match List I with List II and select the correct answer using the codes given below the list

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13/09/202	2, 19:56		CBSE Class 12 Chemistry General Principles and Processes of Isolation of Elements MCQs, Multiple Choice Ques	stions for Chemistry
	List I		List II	• Entrepreneurship
1.	Ti	А.	Bauxite	<ul> <li>Marketing</li> </ul>
2.	Si	В.	Cerussite	<ul> <li>Painting and Sculpture</li> </ul>
3.	Al	С.	Van-Arkel method	
4.	Pb	D.	Zone refining	<ul> <li><u>Physical Education</u></li> </ul>
				• <u>English</u>
				<ul> <li><u>Graphics Design</u></li> </ul>
(a) 1	–B, 2–A, 3–C, 4–D			• <u>Geography</u>
. ,	I–B, 2–C, 3–A, 4–B			• <u>Hindi</u>
(c) 1	–С, 2–А, 3–В, 4–D			• Home Science
(d)	I-C, 2-D, 3-A, 4-B			• <u>History</u>
Δns	wer : D			<ul> <li>Informatics Practices</li> </ul>

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- **Answer: B**

## Question : The main reactions occurring in blast furnace during extraction of iron from haematite are (i) $Fe_2O_2 + 3CO_2$

(i)	$Fe_2O_3 + 3CO \longrightarrow 2Fe + 3CO_2$
	$FeO + SiO_2 \longrightarrow FeSiO_2$

- (ii)  $\operatorname{Fe}_2O_3 + 3C \longrightarrow 2\operatorname{Fe} + 3\operatorname{CO}$ (iv)  $\operatorname{CaO} + \operatorname{SiO}_2 \longrightarrow \operatorname{CaSiO}_3$
- (a) (i) and (iii) (b) (ii) and (iv) (c) (i) and (iv) (d) (i), (ii) and (iii)

#### **Answer : C**

#### Question : Which process of metallurgy of copper is represented by below equation?

 $2CuFeS_2 + O_2 \longrightarrow Cu_2S + 2FeS + SO_2$ 

(a) Concentration (b) Roasting (c) Reduction (d) Purification

#### **Answer: B**

#### Question : The undesired material which is present with the element in its ore is known as \_\_\_\_\_.

(a) slag (b) gang (c) flux (d) sludge

Question : Ellingham diagrams are plots drawn between \_\_\_\_ and \_\_\_\_ for formation of oxides.

(a)Temperature and change in enthalpy

(b)Change in Gibbs free energy and Pressure

(c)Change in Gibbs free energy and temperature

(d)Change in enthalpy and pressure

**Answer: B** 

Question : In the electrolysis of aluminium, the anodes in the cell have to be replaced frequently because .

(a) they melt in the high temperature inside the cell.

(b) they are burned away by O reacting with the carbon anode to form  $CO_2$  gas.

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(c) they get coated in a thick crust of solid  $AI_2O_3$ .

(d)the AI formed on them has to be removed.

#### Answer : B

## Question : During the formation of metal oxides from metal and oxygen, what happens to the change in Gibbs free energy with the little increase of temperature?

- (a)Increases
- (b)Decreases
- (c)Remains same
- (d)Insufficient data

#### Answer : A

### Question : Which of the following is not an ore of iron?

(a)Siderite (b)Haematite (c)Limonite

(d)Dolomite

#### Answer : D

#### Question : Distillation is very useful for low boiling metals like\_\_\_\_.

(a)zinc and mercury
(b)sodium and magnesium
(c)iron and copper
(d)gold and silver
Answer : A

#### Question : Which one of the following is an ore of silver ?

- (a) Argentite
- (b) Stibnite
- (c) Haematite
- (d) Bauxite

#### Answer : A

#### **Question : Cinnabar is an ore of**

(a) Hg (b) Cu (c) Pb (d) Zn

#### **Answer : A**

#### Question : An example of an oxide ore is

(a) Bauxite

(b) Malachite

(c) Zinc blende

(d) Feldspar

Answer : A

### Question : The natural materials from which an element can be extracted economically are called

(a) ores

(b) minerals

(c) gangue

(d) None of these

#### Question : The most abundant metal on the surface of the earth is

(a) Fe

(b) Al

(c) Ca

(d) Na

#### **Answer: B**

#### **Question : Which of the following is an ore of tin ?**

(a) Carborundum

(b) Epsomite

(c) Cassiterite

(d) Spodumene

#### Answer : C

#### **Question : Which of the following is chalcopyrite?**

(a) CuFeS<sub>2</sub>
(b) FeS<sub>2</sub>
(c) KMgCl<sub>3</sub>.6H<sub>2</sub>O
(d) Al<sub>2</sub>O<sub>3</sub>.2H<sub>2</sub>O

#### Answer : A

#### **Question : Haematite is the ore of**

(a) Pb (b) Cu (c) Fe (d) Au

#### Answer : C

#### **Question : Composition of azurite mineral is**

(a) CuCO<sub>3</sub>CuO
(b) Cu(HCO<sub>3</sub>)<sub>2</sub>. Cu(OH)<sub>2</sub>
(c) 2CuCO<sub>3</sub>.Cu(OH)<sub>2</sub>
(d) CuCO<sub>3</sub>. 2Cu(OH)<sub>2</sub>

#### Answer : C

#### Question : Which one of the following is a mineral of iron ?

(a) Malachite

(b) Cassiterite

(c) Pyrolusite

(d) Magnetite

Answer : D

#### Question : All ores are minerals, while all minerals are not ores because

(a) the metal can't be extracted economically from all the minerals(b) minerals are complex compounds(c) the minerals are obtained from mines(d) all of these are correct

#### Question : Which one of the following is not a sulphide ore?

- (a) Magnetite
- (b) Iron pyrites
- (c) Copper glance
- (d) Sphalerite

#### **Answer : A**

#### Question : The impurities associated with mineral used in metallurgy are called collectively?

- (a) Slag
- (b) Flux
- (c) Gangue
- (d) Ore

#### Answer : C

#### Question : The most abundant element in the earth's crust (by weight) is

(a) Si

(b) Al

(c) O

(d) Fe

#### Answer : C

#### **Question : Malachite is an ore of**

- (a) iron
- (b) copper
- (c) mercury
- (d) zinc

#### Answer : B

#### Question : Cassiterite is an ore of

(a) Mn (b) Ni (c) Sb (d) Sn

#### Answer : D

#### Question : Galena is an ore of

(a) Pb (b) Hg (c) Zn

(d) None of these

Answer : A

## **Question : The metal always found in the free states is** (a) Au

(b) Ag

(c) Cu

(d) Na

#### Question : Matrix is defined as -

- (a) the unwanted foreign material present in the ore
- (b) the flux added to remove the unwanted impurities from ore
- (c) the slag formed as a result of the reaction of flux with gangue
- (d) the material used in the reduction of metal oxide to metal

#### **Answer : A**

#### Question : Which of the following pair is incorrectly matched ?

- (a) Magnetite  $Fe_3O_4$
- (b) Copper glance Cu<sub>2</sub>S
- (c) Calamine  $ZnCO_3$
- (d) Zincite ZnS

#### Answer : C

#### Question : Which one of the following ores is best concentrated by froth-flotation method ?

- (a) Galena
- (b) Cassiterite
- (c) Magnetite
- (d) Malachite

#### Answer : C

#### Question : Froth floatation process is used for the metallurgy of

- (a) chloride ores
- (b) amalgams
- (c) oxide ores
- (d) sulphide ores

#### Answer : D

#### (a) levigation

(b) electromagnetic separation

(c) floatation

(d) liquefaction

#### Answer : B

Question : While extracting an element from its ore, the ore is grounded and leached with dil. potassium cyanide solution to form the soluble product potassium argento cyanide. The element is

(a) Lead

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- (b) Chromium
- (c) Manganese
- (d) Silver

#### **Answer : D**

#### Question : The method of concentrating the ore which makes use of the difference in density between ore and impurities is called

#### (a) levigation

- (b) leaching
- (c) magnetic separation
- (d) liquifaction

#### **Answer: A**

#### **Question : Leaching is a process of**

- (a) reduction
- (b) concentration
- (c) refining
- (d) oxidation

#### **Answer: B**

#### Question : Which one of the following ores is concentrated by chemical leaching method?

- (a) Galena
- (b) Copper pyrite
- (c) Cinnabar
- (d) Argentite

#### **Answer : D**

#### Question : Electromagnetic separation is used in the concentration of

(a) copper pyrites

(b) bauxite

(c) cassiterite

(d) cinnabar

Answer : C

#### Question : For which ore of the metal, froth floatation method is used for concentration?

(a) Horn silver

(b) Bauxite

(c) Cinnabar

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(d) Heamatite

#### **Answer : C**

#### Question : Which of the following metal is leached by cyanide process ?

(a) A	٩g
-------	----

- (b) Na
- (c) Al
- (d) Cu

#### **Answer: A**

#### Question : Which one of the following ores is not concentrated by froth floatation process?

- (a) Copper pyrites
- (b) Pyrargyrite
- (c) Pyrolusite
- (d) Zinc blende

#### Answer : C

#### Question : In froth flotation process many chemicals (frother , collector, activator, and depressant) are used. Which of the following is a frother ?

(a) CuSO<sub>4</sub>

- (b) NaCN+ alkali
- (c) Pine oil
- (d) Potassium xanthate

#### **Answer : C**

#### **Question : Froth flotation process is based on**

- (a) wetting properties of ore particle
- (b) specific gravity of ore particles
- (c) magnetic properties of ore particles
- (d) electrical properties of ore particles

#### **Answer : A**

#### Question : In the froth flotation process of concentration of ores, the ore particles float because they:

(a) are light

(b) are insoluble

(c) have the surface which is not wetted easily

(d) have a constant electrical charge

#### **Question : Main function of roasting is**

- (a) to remove volatile substances
- (b) oxidation
- (c) reduction
- (d) slag formation

#### Answer : A

#### Question : Roasting is generally done in case of the

- (a) oxide ores
- (b) silicate ores
- (c) sulphide ores
- (d) carbonate ores

#### Answer : C

#### Question : Heating of pyrites in air for oxidation of sulphur is called

- (a) roasting
- (b) calcination
- (c) smelting
- (d) slagging

#### Answer : A

#### Question : The role of calcination in metallurgical operations is

- (a) to remove moisture
- (b) to decompose carbonates
- (c) to drive off organic matter
- (d) to decompose carbonates and drive off moisture and organic matter

#### **Answer : D**

#### Question : General method for the extraction of metal from oxide ore is

(a) carbon reduction

(b) reduction by aluminium

(c) reduction by hydrogen

(d) electrolytic reduction

Answer : A

#### Question : Function of the flux added during smelting is

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- (a) to make ore porous
- (b) to remove gangue
- (c) to make reduction easier
- (d) to precipitate slag

#### Answer : B

#### **Question : Process followed before reduction of carbonate ore is**

- (a) calcination
- (b) roasting
- (c) liquation
- (d) polling

#### Answer : A

#### Question : Calcination is the process in which :

(a) ore is heated above its melting point to expel  $H_2O$  or  $CO_2$  or  $SO_2$ 

- (b) ore is heated below its melting point to expel volatile impurities
- (c) ore is heated above its melting point to remove S, As and Sb as SO<sub>2</sub>, As2O<sub>3</sub> and Sb2O<sub>3</sub> respectively
- (d) ore is heated below its melting point to expel  $H_2O$  or  $CO_2$

#### Answer : D

## Question : When a metal is to be extracted from its ore and the gangue associated with the ore is silica, then

- (a) an acidic flux is needed
- (b) a basic flux is needed
- (c) both acidic and basic fluxes are needed
- (d) Neither of them is needed

#### Answer: B

Question : Which of the following fluxes is used to remove acidic impurities in metallurgical process?

(a) Silica

(b) Lime stone

(c) Sodium chloride

(d) Sodium carbonate

Answer : B

Question : Which of the following reactions is an example for calcination process ?

(a)  $2Ag + 2HCI + (O) \rightarrow 2AgCI + H_2O$ 

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(b) 2Zn + O<sub>2</sub> →2ZnO

(c)  $2ZnS + 3O_2 \rightarrow 2ZnO + 2SO_2$ 

(d) MgCO<sub>3</sub>  $\rightarrow$  MgO+CO<sub>2</sub>

#### **Answer : D**

#### Question : After partial roasting the sulphide of copper is reduced by

- (a) cyanide process
- (b) electrolysis
- (c) reduction with carbon
- (d) self reduction

#### **Answer : D**

#### Question : Hydro-metallurgical process of extraction of metals is based on

- (a) complex formation
- (b) hydrolysis
- (c) dehydration
- (d) dehydrogenation

#### **Answer: A**

#### $\label{eq:Question: 2CuFeS_2 + O_2 \rightarrow Cu_2S + 2FeS + SO_2 \ Which \ process \ of \ metallurgy \ of \ copper \ is \ represented$ by above equation?

- (a) Concentration
- (b) Roasting
- (c) Reduction
- (d) Purification

#### **Answer: B**

#### Question : Which of the following is not used as a collector ?

(a) Pine oil

(b) Xanthates

(c) Cresols

(d) Fatty acids

**Answer: B** 

Question : According to Ellingham diagram, the oxidation reaction of carbon to carbon monoxide may be used to reduce which one of the following oxides at the lowest temperature ?

(a) Al<sub>2</sub>O<sub>3</sub>

(b) Cu<sub>2</sub>O

- (c) MgO
- (d) ZnO

#### Answer : B

#### Question : The process does not involve a catalyst is :

- (a) Haber process
- (b) Contact process
- (c) Thermite process
- (d) Ostwald process

#### Answer : C

#### **Question : Chief ore of Al is :**

(a) cryolite

(b) bauxite

(c) feldspar

(d) kaolin

#### **Answer : B**

#### **Question : Froth floatation is a process of:**

(a) Oxidation

(b) Reduction

(c) Refining

(d) Concentration

#### **Answer : D**

#### **Question : Flux is used to remove :**

(a) basic impurities

(b) acidic impurities

(c) all types of impurities

(d) acidic and basic both impurities

#### Answer : D

#### **Question : Mac Arthur process is used for the extraction of:**

(a) Au

(b) Pt

(c) Cu

#### Answer : A

Question : In the extraction of copper from its sulphide ore, the metal is formed by reduction of Cu<sub>2</sub>O with :

(a) FeS (b) CO

(c) Cu<sub>2</sub>S

(d) SO<sub>2</sub>

Answer : C

#### Question : Which of the following is a carbonate ore?

- (a) Pyrolusite
- (b) Malachite
- (c) Diaspore
- (d) Cassiterite

#### Answer : B

## Question : Carbon and CO gas are used to reduce which of the following pairs of metal oxides for extraction of metals?

(a) FeO, SnO
(b) SnO, ZnO
(c) BaO, Na<sub>2</sub>O<sub>2</sub>
(d) FeO, ZnO

#### **Answer : D**

## Question : In metallurgical process of aluminium, cryolite is mixed with alumina in its molten state, because it

- (a) decreases the amount of alumina
- (b) oxidises the alumina
- (c) increases the melting point of alumina
- (d) decreases the melting point of alumina

#### **Answer : D**

#### Question : Match list I with list II and select the correct answer using the codes given below the lists:

#### List I

#### List II

- I. Cyanide process A. Ultrapure Ge
- II. Floatation process B. Pine oil
- III. Electrolytic reduction C. Extraction of Al
- IV. Zone refining D. Extraction of Au
- (a) I-C, II-A, III-D, IV-B
  (b) I-D,II-B,III-C,IV-A
  (c) I-C,II-B,III-D, IV-A
  (d) I-D,II-A,III-C,IV-B

#### Answer : D

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CBSE Class 12 Chemistry Haloalkanes and Haloarenes MCQs with answers available in Pdf for free download. The MCQ Questions for Class 12 Chemistry with answers have been prepared as per the latest syllabus, NCERT books and examination pattern suggested in Standard 12 by CBSE, NCERT and KVS. Multiple Choice Questions are an important part of Term 1 and Term 2 exams for Grade 12 Chemistry and if practiced properly can help you to get higher marks. Refer to more Chapter-wise MCQs for NCERT Class <u>12 Chemistry</u> and also download more latest study material for all subjects

## Haloalkanes and Haloarenes Class 12 Chemistry MCQ

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Haloalkanes and Haloarenes in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

## Haloalkanes and Haloarenes MCQ Questions Class 12 **Chemistry with Answers**

Question : Among the following which one can have a meso form?

(a)  $CH_3CH(OH)CH(CI)C_2H_5$ (b) CH<sub>3</sub>CH(OH)CH(OH)CH<sub>3</sub> (c)  $C_2H_5CH(OH)CH(OH)CH_3$ (d) HOCH<sub>2</sub>cH(Cl)CH<sub>3</sub>

**Answer: B** 

Question : Which of the following compounds has the highest boiling point?

(a) CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CI (b) CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CI (c) CH<sub>3</sub>CH(CH<sub>3</sub>)CH<sub>2</sub>CI

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#### (d) (CH<sub>3</sub>)<sub>3</sub>CCl

#### Answer : B

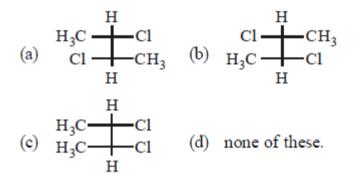
Question : In the following sequence of the reations, what is D?

$$\underbrace{\bigcirc}^{CH_3} \xrightarrow{[0]} A \xrightarrow{SOC1_2} B$$
$$\underbrace{\xrightarrow{NaN_3} C \xrightarrow{Heat} D}$$

- (a) Primary amine
- (b) An amide
- (c) Phenyl isocyanate
- (d) A chain lengthened hydrocarbon

#### Answer : C

#### **Question : Which of the following is optically inactive?**



#### Answer : C

#### **Question : The pesticide DDT slowly changes to**

- (a) CCl<sub>3-</sub>CHO and chlorobenzene
- (b) p, p'-Dichlorodiphenylethene
- (c) p, p'-Dichlorodiphenyldichloroethane
- (d) p, p'-Dichlorodiphenyldichloroethene

#### Answer : D

#### Question : Rectified spirit is a mixture of

(a) 95% ethyl alcohol + 5% water

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(b) 94% ethyl alcohol + 4.53 water
(c) 94.4% ethyl alcohol + 5.43% water
(d) 95.87% ethyl alcohol + 4.13% water

Answer : D

Question : Which of the following is an example of SN<sub>2</sub> reaction?

(a) 
$$CH_3Br + OH^- \longrightarrow CH_3OH + Br^-$$

(b) 
$$CH_3 - CH - CH_3 + OH^- \longrightarrow CH_3 - CH - CH_3$$
  
 $|_{Br} OH$ 

(c) 
$$CH_3CH_2OH \xrightarrow{-H_2O} CH_2 = CH_2$$

(d) 
$$(CH_3)_3C - Br + OH^- \longrightarrow (CH_3)_3COH + Br^-$$

#### **Answer: A**

#### Question : Which of the following pairs is/are correctly matched?

	Reaction	Product
I.	RX+AgCN	RNC
Π.	RX+KCN	RCN
	RX+KNO <sub>2</sub> RX+AgNO <sub>2</sub>	R - N = 0 R-O-N=0

- (a) Only I
- (b) I and II
- (c) III and IV
- (d) I, II, III and IV

#### **Answer: B**

Question : The solution of a chemical compound reacts with AgNO<sub>3</sub> solution to form a white preciptate of Y which dissolves in NH<sub>4</sub>OH to give a complex Z. When Z is treated with dilute HNO<sub>3</sub>, Y reappears. The chemical compound X can be

(a) NaCl

(b) CH<sub>3</sub>Cl

(c) NaBr

(d) Nal

#### **Answer : A**

#### Question : The synthesis of alkyl fluorides is best accomplished by

- (a) Finkelstein reaction
- (b) Swarts reaction
- (c) Free radical fluorination

#### **Answer: B**

#### Question : Which of the following is a primary halide?

(a) Isopropyl iodide

(b) Secondary butyl iodide

(c) Tertiary butyl bromide

(d) Neohexyl chloride

#### **Answer : D**

#### Question : When two halogen atoms are attached to same carbon atom then it is :

(a) vic-dihalide

- (b) *gem*-dihalide
- (c)  $\alpha$ ,  $\omega$  -halide
- (d)  $\alpha$ ,  $\beta$  -halide

#### Answer : B

#### **Question : Gem-dibromide is**

(a)  $CH_3CH(Br)CH_2(Br)$ 

(b) CH<sub>3</sub>CBr<sub>2</sub>CH<sub>3</sub>

(c)  $CH_2(Br)CH_2CH_2$ 

(d)  $CH_2BrCH_2Br$ 

#### Answer : B

Question : How many structural isomers are possible for a compound with molecular formula  $C_3H_7CI$ ?

(a) 2

(b) 5

- (c) 7
- (d) 9

#### Answer : A

#### Question : The compound which contains all the four 1°, 2°, 3° and 4° carbon atoms is

- (a) 2, 3-dimethyl pentane
- (b) 3-chloro-2, 3-dimethylpentane
- (c) 2, 3, 4-trimethylpentane
- (d) 3, 3-dimethylpentane

#### **Answer: B**

#### Question : IUPAC name of (CH<sub>3</sub>)<sub>3</sub>CCI

(a) 3-Chlorobutane

(b) 2-Chloro-2-methylpropane

(c) *t*-butyl chloride

(d) *n*-butyl chloride

Answer : B

#### Question : IUPAC name of CH<sub>3</sub>CH<sub>2</sub>C(Br) = CH—Cl is

- (a) 2-bromo-1-chloro butene
- (b) 1-chloro-2-bromo butene
- (c) 3-chloro-2-bromo butene
- (d) None of the above

#### Answer : A

#### Question : The IUPAC name of CH<sub>2</sub> = CH - CH<sub>2</sub>Cl is

- (a) Allyl chloride
- (b) 1-chloro-3-propene
- (c) Vinyl chloride
- (d) 3-chloro-1-propene

#### Answer : D

#### **Question : Which of the following halide is 2°?**

- (a) Isopropyl chloride
- (b) Isobutyl chloride
- (c) *n*-propyl chloride
- (d) *n*-butyl chloride

#### Answer : A

#### **Question : Benzene hexachloride is**

- (a) 1, 2, 3, 4, 5, 6 hexachlorocyclohexane
- (b) 1, 1, 1, 6, 6, 6 hexachlorocyclohexane
- (c) 1, 6 phenyl 1, 6 chlorohexane
- (d) 1, 1 phenyl 6, 6 -chlorohexane

#### **Answer : A**

#### Question : Phosgene is a common name for

(a) phosphoryl chloride

(b) thionyl chloride

(c) carbon dioxide and phosphine

(d) carbonyl chloride

Answer : D

Question : C – X bond is strongest in

(a) CH<sub>3</sub>Cl

(b) CH<sub>3</sub>Br

(c)  $CH_3F$ 

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## Answer : C

#### Question : Which of the following will have the maximum dipole moment?

(a) CH<sub>3</sub>F

(b) CH<sub>3</sub>Cl

(c)  $CH_3Br$ 

(d) CH<sub>3</sub>I

#### Answer : B

#### Question : The decreasing order of boiling points of alkyl halides is

- (a) RF > RCI > RBr > RI
- (b) RBr > RCI > RI > RF
- (c) RI > RBr > RCI > RF
- (d) RCI > RF > RI > RBr

#### Answer : C

#### **Question : Halogenation of alkanes is**

- (a) a reductive process
- (b) an oxidative process
- (c) an isothermal process
- d) an endothermal process

#### Answer : B

#### Question : Ethylene dichloride can be prepared by adding HCl to

- (a) Ethane
- (b) Ethylene
- (c) Acetylene
- (d) Ethylene glycol

#### Answer : D

Question : In which of the following conversions, phosphorus pentachloride is used as the reagent?

(a)  $H_2C = CH_2 \rightarrow CH_3CH_2CI$ (b)  $CH_3CH_2OH \rightarrow CH_3CH_2CI$ 

(c) H<sub>3</sub>C-O - CH<sub>3</sub>  $\rightarrow$  CH<sub>3</sub>Cl

(d) CH - CH  $\rightarrow$  CH<sub>2</sub> = CHCl

Answer : B

## Question : The best method for the conversion of an alcohol into an alkyl chloride is by treating the alcohol with

#### (a) $PCI_5$

(b) dry HCl in the presence of anhydrous ZnCl<sub>2</sub>

(c)  $SOCl_2$  in presence of pyridine

(d) None of these

#### Answer : C

#### Question : Which of the following is liquid at room temperature (b.p. is shown against it) ?

- (a) CH<sub>3</sub>I 42°C
- (b) CH<sub>3</sub>Br 3°C
- (c) C<sub>2</sub>H<sub>5</sub> C I 12°C
- (d) CH<sub>3</sub>F –78°C

#### Answer : A

## Question : The catalyst used in the preparation of an alkyl chloride by the action of dry HCl on an alcohol is

(a) anhydrous AlCl<sub>3</sub>

(b) FeCl<sub>3</sub>

(c) anhydrous ZnCl<sub>2</sub>

(d) Cu

#### Answer : C

#### **Question : Chlorobenzene is prepared commercially by**

(a) Raschig process

(b) Wurtz Fittig reaction

(c) Friedel-Craft's reaction

(d) Grignard reaction

#### Answer : A

#### Question : In the preparation of chlorobenzene from aniline, the most suitable reagent is

(a) Chlorine in the presence of ultraviolet light

(b) Chlorine in the presence of  $AICI_3$ 

(c) Nitrous acid followed by heating with  $Cu_2Cl_2$ 

(d) HCl and  $Cu_2Cl_2$ 

Answer : C

#### **Question : Which of the following possesses highest melting point?**

- (a) Chlorobenzene
- (b) m-dichlorobenzene
- (c) o-dichlorobenzene
- (d) p-dichlorobenzene

#### Answer : D

#### Question : Conant Finkelstein reaction for the preparation of alkyl iodide is based upon the fact that

- (a) Sodium iodide is soluble in methanol, while sodium chloride is insoluble in methanol
- (b) Sodium iodide is soluble in methanol, while NaCl and NaBr are insoluble in methanol
- (c) Sodium iodide is insoluble in methanol, while NaCl and NaBr are soluble
- (d) The three halogens differ considerably in their electronegativity

#### Answer: B

#### Question : Which of the following reactions is an example of nucleophilic substitution reaction?

- (a) 2 RX + 2 Na  $\rightarrow$  R R + 2 NaX
- (b)  $RX + H2 \rightarrow RH + HX$
- (c)  $RX + Mg \rightarrow RMgX$
- (d)  $RX + KOH \rightarrow ROH + KX$

#### Answer : D

#### **Question : Which one is most reactive towards SN1 reaction ?**

- (a)  $C_6H_5CH(C_6H_5)Br$
- (b)  $C_6H_5CH(CH_3)Br$
- (c)  $C_6H_5C(CH_3)(C_6H_5)Br$
- (d)  $C_6H_5CH_2Br$

#### Answer : C

#### Question : A Grignard reagent may be made by reacting magnesium with

(a) Methyl amine

(b) Diethyl ether

(c) Ethyl iodide

(d) Ethyl alcohol

Answer : C

## Question : Which one of the following halogen compounds is difficult to be hydrolysed by SN1 mechanism?

(a) Tertiary butyl chloride

(b) Isopropyl chloride

(c) Benzyl chloride

#### (d) Chlorobenzene

#### **Answer : D**

#### Question : The order of reactivity of the given haloalkanes towards nucleophile is :

- (a) RI > RBr > KCl
- (b) RCI > RBr > RI
- (c) RBr > RCl > Rl
- (d) RBr > RI > RCI

#### **Answer : A**

#### **Question : Most reactive halide towards SN1 reaction is**

- (a) *n*-Butyl chloride
- (b) sec-Butyl chloride
- (c) tert-Butyl chloride
- (d) Allyl chloride

#### Answer : C

#### Question : In $S_{\mbox{\scriptsize N}}1$ reaction, the recemization takes place. It is due to

- (a) inversion of configuration
- (b) retention of configuration
- (c) conversion of configuration

#### (d) Both (a) and (b)

#### Question : The order of reactivities of the following alkyl halides for a $S_N2$ reaction is

- (a) RF > RCI > RBr > RI
- (b) RF > RBr > RCI > RI
- (c) RCI > RBr > RF > RI

#### (d) RI > RBr > RCI > RF

Question : Optically active isomers but not mirror images are called

(a) enantiomers

(b) mesomers

(c) tautomers

(d) diastereomers

**Question : SN<sub>2</sub> mechanism proceeds through intervention of** 

(a) carbonium ion

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- (b) transition state
- (c) free radical
- (d) carbanion

#### Answer : B

#### Question : Which among MeX, RCH<sub>2</sub>X, R<sub>2</sub>CHX and R<sub>3</sub>CX is most reactive towards SN<sub>2</sub> reaction?

(a) MeX

(b) RCH<sub>2</sub>X

(c)  $R_2 C HX$ 

(d)  $R_3CX$ 

#### **Answer : A**

#### Question : Isopropyl chloride undergoes hydrolysis by

- (a) S<sub>N</sub>1 mechanism
- (b) S<sub>N</sub>2 mechanism
- (c)  $S_N 1$  and  $S_N 2$  mechanisms
- (d) Neither  $S_N 1$  nor  $S_N 2$  mechanism

#### Answer : C

#### Question : Tertiary alkyl halides are practically inert to substitution by SN2 mechanism because of

- (a) steric hindrance
- (b) inductive effect
- (c) instability
- (d) insolubility

#### **Answer : B**

#### Question : Which of the following is the correct order of decreasing SN2 reactivity?

(a)  $R_2 C HX > R_3 CX > RCH_2 X$ 

(b) RCHX >  $R_3CX > R_2CHX$ 

(c)  $RCH_2X > R_2CHX > R_3CX$ 

(d)  $R_3CX > R_2CHX > RCH_2X$ .

(X is a halogen)

Answer : C

#### Question : Which of the following alkyl halides is used as a methylating agent?

(a)  $C_2H_5Br$ 

(b)  $C_6H_5CI$ 

(c) CH<sub>3</sub>I

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(d)  $C_2H_5CI$ 

#### **Answer : C**

#### Question : Which of the following is an optically active compound ?

- (a) 1-Butanol
- (b) 1-Propanol
- (c) 2-Chlorobutane
- (d) 4-Hydroxyheptane

#### Answer : C

#### Question : An important chemical method to resolve a racemic mixture makes use of the formation of

- (a) a meso compound
- (b) enantiomers
- (c) diasteromers
- (d) racemates

#### **Answer : C**

#### Question : The process of separation of a racemic modification into d and I -enantiomers is called

- (a) Resolution
- (b) Dehydration
- (c) Revolution
- (d) Dehydrohalogenation

#### **Answer: A**

#### Question : Mg reacts with RBr best in

- (a)  $C_2H_5OC_2H_5$
- (b)  $C_6H_5OCH_3$
- (c)  $C_6H_5N(CH_3)_2$
- (d) Equally in all the three

#### **Answer : A**

#### Question : Which of the following will have a mesoisomer also?

(a) 2, 3- Dichloropentane

(b) 2, 3-Dichlorobutane

(c) 2-Chlorobutane

(d) 2-Hydroxypropanoic acid

#### **Answer : B**

#### Question : Which of the following compounds is optically active ?

(a) CH<sub>3</sub>CHCICOOH

(b) CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>3</sub>

(c)  $(CH_3)_2CHOH$ 

(d) (CH<sub>3</sub>)<sub>3</sub>CCI

#### Answer : B

#### **Question : Racemic compound has**

- (a) equimolar mixture of enantiomers
- (b) 1 : 1 mixture of enantiomer and diastereomer
- (c) 1 : 1 mixture of diastereomers
- (d) 1:2 mixture of enantiomers

#### Answer : A

## Question : 2-Bromopentane is heated with potassium ethoxide in ethanol. The major product obtained is

- (a) 2-ethoxypentane
- (b) pentene-1
- (c) trans-2-pentene
- (d) cis-pentene-2

#### Answer : C

#### Question : An organic molecule necessarily shows optical activity if it

- (a) contains asymmetric carbon atoms
- (b) is non-polar
- (c) is non-superimposable on its mirror image
- (d) is superimposable on its mirror image

#### Answer : C

**Question :** Chloroform and conc. HNO<sub>3</sub> react to produce

(a) CHCl<sub>2</sub>NO<sub>2</sub>

(b) CHCl<sub>2</sub>HNO<sub>3</sub>

(c) CCI<sub>3</sub>NO<sub>2</sub>

(d) CCI<sub>3</sub>NO<sub>3</sub>

Answer : C

Question : Benzene diazonium chloride reacts with hypophosphorous acid to produce:

- (a) phenol
- (b) benzene
- (c) p-hydroxyazobenzene
- (d) benzonitrile

#### **Answer : B**

#### Question : B.H.C. is used as an

(a) Insecticide

(b) Disinfectant

(c) Mosquito repellent

(d) Antiseptic

#### **Answer: A**

#### Question : Which one of the following produces acyl halide by treatment with PCI<sub>5</sub>?

(a) Alcohols

(b) Esters

(c) Acids

(d) Carbonyl compounds

#### Answer : C

#### **Question : 5 Gammexane is**

(a) Chloral

(b) BHC

(c) DDT

(d) HCB

#### **Answer: B**

#### Question : Among the following, insecticide is

(a) BHC

(b) Phosphene

(c) Chloral

(d) Aspirin

#### **Answer : A**

#### Question : Which of the following is a chiral compound?

(a) hexane

(b) n-butane

(c) methane

(d) 2,3,4-trimethylhexane

#### **Answer : D**

#### Question : Which of the following is most stable?

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- (a) 1-butene
- (b) 1-pentene
- (c) 2-butene
- (d) 2-pentene

#### Answer : C

Question : Among the following, the most reactive towards alcoholic KOH is :

(a) CH<sub>2</sub>=CHBr

(b)  $CH_3COCH_2CH_2Br$ 

(c) CH<sub>3</sub>CH<sub>2</sub>Br

#### (d) CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>Br

#### **Answer : B**

#### Question : Among the following, the dissociation constant is highest for

(a)	C <sub>6</sub> H <sub>5</sub> OH	(b)	$\rm C_6H_5CH_2OH$
(c)	$\mathrm{CH}_3\mathrm{C} \equiv \mathrm{CH}$	(d)	$\operatorname{CH}_3\operatorname{NH}_3^+\operatorname{Cl}^-$

#### Answer : D

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# Polymers Class 12 Chemistry MCQ

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Polymers in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

# **Polymers MCQ Questions Class 12 Chemistry with Answers**

Question : Which of the following is not linear polymer?

(a) Bakelite (b) Polyester (c) Cellulose (d) High density polyethene

Answer : A

### Question : A polymer is formed when simple chemical units

(a) combine to form long chains (b) combine to form helical chains (c) break up (d) become round

Answer: A

**Question : Polymer formation from monomers starts by** 



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- (a) condensation reaction between monomers
- (b) coordinate reaction between monomers
- (c) conversion of monomer to monomer ions by protons
- (d) hydrolysis of monomers.

### Answer : A

### Question : On the basis of mode of formation, polymers can be classified?

- (a) as addition polymers only
- (b) as condensation polymers only
- (c) as copolymers
- (d) both as addition and condensation polymers

### Answer : D

### Question : In addition polymer monomer used is

- (a) unsaturated compounds
- (b) saturated compounds
- (c) bifunctional saturated compounds
- (d) trifunctional saturated compounds

### Answer : A

### Question : Nylon 66 belongs to the class of

- (a) Addition polymer
- (b) Condensation polymer
- (c) Addition homopolymer
- (d) Condensation heteropolymer

### **Answer : D**

# Question : A polymer made from a polymerization reaction that produces small molecules (such as water) as well as the polymer is classified as a/an ..... polymer.

(a) addition(b) natural(c) condensation(d) elimination

### Answer : C

### **Question : In elastomer, intermolecular forces are**

(a) strong (b) weak(c) nil(d) None of these

**Answer: B** 

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### Question : A thermoplastic among the following is

(a) bakelite(b) polystyrene(c) terylene(d) urea-formaldehyde resin

### Answer : B

### Question : Which is an example of thermosetting polymer?

(a) Polythene (b) PVC(c) Neoprene (d) Bakelite

### Answer : C

#### **Question : Which is not true about polymers?**

- (a) Polymers do not carry any charge
- (b) Polymers have high viscosity
- (c) Polymers scatter light
- (d) Polymers have low molecular weight

#### Question : Which of the following belongs to the class of natural polymers?

- (a) Proteins (b) Cellulose
- (c) Rubber (d) All of these

#### Question : Which of the following natural products is not a polymer?

- (a) DNA (b) Cellulose
- (c) ATP (d) Urease

#### Question : Among the following a natural polymer is

- (a) cellulose (b) PVC
- (c) teflon (d) polyethylene

### **Question : Which of the following is not a biopolymer ?**

- (a) Proteins (b) Rubber
- (c) Cellulose (d) RNA

### **Question : Rayon is :**

(a) synthetic plastic (b) natural rubber

(c) natural silk (d) artificial silk

#### **Question : Protein is a polymer of:**

(a) glucose (b) terephthalic acid

**Question : Natural silk is a** 

(a) polyester (b) polyamide

(c) polyacid (d) polysaccharide

**Question : Polymers are:** 

(a) micromolecules (b) macromolecules

(c) sub-micromolecules (d) None of the above

### **Question : Which of the following is/are a semisynthetic polymers?**

- (a) Cellulose acetate (b) Polyvinyl chloride
- (c) Cellulose nitrate (d) Both (a) and (c)

#### Question : Which of the following is not linear polymer ?

- (a) Bakelite (b) Polyester
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- (c) terylene (d) urea-formaldehyde resin

#### **Question : Which is an example of thermosetting polymer?**

(a) Polythene (b) PVC

(c) Neoprene (d) Bakelite

#### Question : Which of the following is thermoplastic ?

(a) Bakelite (b) Polyethylene

(c) Terylene (d) All of these

#### **Question : Thermosets are:**

(a) cross-linked polymers

- (b) don't melt or soften on heating
- (c) cross-linking is usually developed at the time of moulding where they harden reversibly

#### (d) all of the above

#### **Question : Which is/are true for elastomers?**

(a) These are synthetic polymers possessing elasticity

(b) These possess very weak intramolecular forces or attractions between polymer chains

(c) Vulcanised rubber is an example of elastomer

(d) All of the above

Question : Among the following polymers the strongest molecular forces are present in

(a) elastomers (b) fibres

(c) thermoplastics (d) thermosetting polymers

#### Question : Three dimensional molecular structure with cross links are formed in the case of a

- (a) thermoplastic (b) thermosetting plastic
- (c) Both (a) and (b) (d) None of the above

#### Question : Which of the following polymer is an example of fibre ?

- (a) Silk (b) Dacron
- (c) Nylon-66 (d) All of these

#### Question : Which of the following statements is not correct for fibres?

- (a) Fibres possess high tensile strength and high modulus.
- (b) Fibres impart crystalline nature.
- (c) Characteristic features of fibres are due to strong intermolecular forces like hydrogen bonding.

### (d) All are correct.

#### Question : Which of the following is/are examples of fibres?

- (a) Polyesters (b) Polyamide
- (c) Polythene (d) Both (a) and (b)

### Question : Which of the following can be repeatedly soften on heating?

- (i) Polystyrene (ii) Melamine
- (iii) Polyesters (iv) Polyethylene
- (v) Neoprene
- (a) (i) and (iii) (b) (i) and (iv)
- (c) (iii), (iv) and (v) (d) (ii) and (iv)

#### Question : Which of the following does not undergo addition polymerization?

- (a) Vinylchloride
- (b) Butadiene
- (c) Styrene

(d) All of the above undergoes addition polymerizations

Question : Which of the following is a cross linked polymer?

(a) PVC (b) Bakelite

(c) Polyethylene (d) Rubber

Question : Fibres that have good resistance to stains, chemicals, insects and fungi is

(a) Acrylic (b) Terylene

(c) Nylon (d) All of these

### Question : Which of the following statements is not true about low density polythene?

(a) Tough

(b) Hard

### (c) Poor conductor of electricity

(d) Highly branched structure

### **Question : Low density polythene is prepared by**

### (a) Free radical polymerisation

- (b) Cationic polymerisation
- (c) Anionic polymerisation
- (d) Ziegler-Natta polymerisation

### **Question : The monomer of teflon is**

(a) CHF == CH<sub>2</sub> (b) CF<sub>2</sub> == CF<sub>2</sub>

(c)  $CHC_1 == CHCI$  (d)  $CHF == CHC_1$ 

### Question : The monomer(s) used in the preparation of Orlon, a substitute for wool is/are

(a) caprolactam

### (b) tetrafluoroethene

(c) styrene and 1, 3-butadiene

(d) acrylonitrile

### **Question : Orlon is a polymer of**

(a) styrene (b) tetrafluoroethylene

(c) vinyl chloride (d) acrylonitrile

### Question : Which of the following polymer is used for manufacturing of buckets, dustbins, pipes etc

?

### (b) High density polythene

(c) Teflon

(d) Polyacrylonitrile

### Question : Which of the following catalyst is used in preparation of high density polythene ?

(a) Peroxide catalyst

(b) Ziegler - Natta catalyst

(c) Wilkinson's catalyst

(d) Pd - catalyst

#### **Question : Which of the following statements is false?**

(a) Artificial silk is derived from cellulose.

- (b) Nylon-66 is an example of elastomer.
- (c) The repeat unit in natural rubber is isoprene.
- (d) Both starch and cellulose are polymers of glucose.

#### **Question : Melamine plastic crockery is a copolymer of:**

(a) HCHO and melamine (b) HCHO and ethylene

(c) melamine and ethylene (d) None of these

#### **Question : Caprolactam polymerises to give**

(a) terylene (b) teflon

(c) glyptal (d) nylon-6

#### Question : Nylons, polysters and cotton, all posses strength due to:

(a) intermolecule H-bonding

#### (b) van der Waals' attraction

(c) dipole-dipole interaction

(d) None of the above

#### Question : Nylon 66 is a polyamide obtained by the reaction of

(a)  $COOH(CH_2)4 COOH + NH_2C_6H_4NH_2$ 

#### (b) $COOH(CH_2)_4 COOH + NH_2 (CH_2)_6 NH_2$

(c) COOH (CH<sub>2</sub>)<sub>6</sub> COOH + NH<sub>2</sub> (CH<sub>2</sub>)<sub>4</sub> NH<sub>2</sub>

(d)  $COOHC_6H_4 COOH_(p) + NH_2 (CH_2)_6 NH_2$ 

(a) van der waal's

### (b) hydrogen bonding

(c) dipole-dipole interactions

(d) None of these

### Question : The plastic household crockery is prepared by using

(a) melamine and tetrafluoroethane

(b) malonic acid and hexamethyleneamine

(c) melamine and vinyl acetate

(d) melamine and formaldehyde

#### Question : Which of the following is currently used as a tyre cord ?

- (a) Terylene (b) Polyethylene
- (c) Polypropylene (d) Nylon 6

#### Question : Of the following which one is classified as polyester polymer?

- (a) Terylene (b) Bakelite
- (c) Melamine (d) Nylon-66

#### Question : Which one of the following is not a condensation polymer

- (a) Melamine (b) Glyptal
- (c) Dacron (d) Neoprene

#### **Question : Bakelite is obtained from phenol by reacting with**

- (a)  $(CH_2OH)_2$  (b)  $CH_3CHO$
- (c)  $CH_3 COCH_3$  (d) HCHO

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#### Answer : D

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# Solutions Class 12 Chemistry MCQ

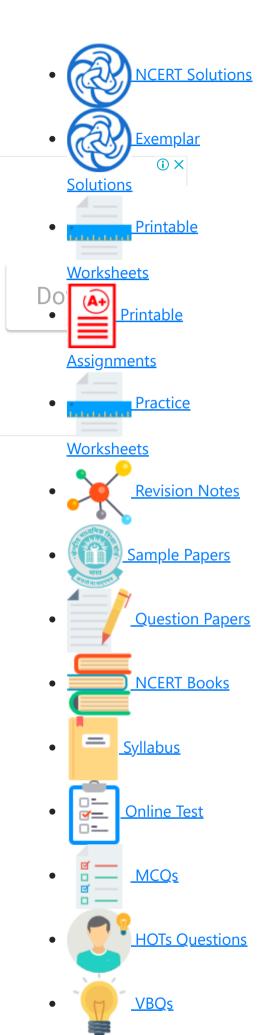
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# Solutions MCQ Questions Class 12 Chemistry with Answers

### Question : An aqueous solution of hydrochloric acid

a) Obeys Raoult's law

b) Shows negative deviation from Raoult's law



c) Shows positive deviation from Raoult's lawd) Obeys Henry's law at all compositions

### **Answer : B**



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Question : Exactly 1 g of urea dissolved in 75 g of water gives a solution that boils at 100.114°C at 760 torr. The molecular weight of urea is 60.1. The boiling point elevation constant for water is

a) 1.02 b) 0.51 c) 3.06 d) 1.51

Answer : B

# Question: 12g of urea is dissolved in 1 litre of water and 68.4 g of sucrose is dissolved in 1 litre of water. The lowering of vapour pressure of first case is

- a) equal to second
- b) greater than second
- c) less than second
- d) double that of second

#### **Answer : A**

Question : At a particular temperature, the vapour pressures of two liquids A and B are respectively 120 and 180 mm of mercury. If 2 moles of A and 3 moles of B are mixed to form an ideal solution, the vapour pressure of the solution at the same temperature will be (in mm of mercury)

a) 156 b) 145

c) 150 d) 108

#### Answer : A

### Question : The freezing point of equimolal aqueous solution will be highest for

a)  $C_6H_5NH_3 + CI^-$  b)  $Ca(NO_3)_2$ c)  $La(NO_3)_2$  d)  $C_6H_{12}O_6$ 

#### Answer : D

### Question : Which of the following 0.10 m aqueous solutions will have the lowest freezing point ?

a) Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> b) C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> c) KCl d) C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>

#### **Answer : A**

Question : A solution containing 10g per dm<sup>3</sup> of urea (molecular mass = 60 gmol<sup>-1</sup>) is isotonic with a 5% solution of a non volatile solute. The molecular mass of this non volatile solute is

a) 300 g mol<sup>-1</sup> b) 350 g mol<sup>-1</sup> c) 200 g mol<sup>-1</sup> b) 250 g mol<sup>-1</sup>

### **Answer : A**

Question : The vapour pressure of a solvent decreases by 10 mm of Hg when a non-volatile solute was added to the solvent. The mole fraction of the solute in the solution is 0.2. What should be the mole fraction of the solvent if the decrease in the vapour pressure is to be 20 mm of Hg ?

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a) 0.8 b) 0.6 c) 0.4 d) 0.2

### Answer : B

Question : A solution containing 1.8 g of a compound (empirical formula  $CH_2O$ ) in 40 g of water is observed to freeze at -0.465° C. The molecular formula of the compound is (K<sub>f</sub> of water = 1.86 kg K mol<sup>-1</sup>)

a) C<sub>2</sub>H<sub>4</sub>O<sub>2</sub> b) C<sub>3</sub>H<sub>6</sub>O<sub>3</sub> c) C<sub>4</sub>H<sub>8</sub>O<sub>4</sub> d) C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>

Answer : D

Question : 10 Which observation(s) reflect(s) colligative properties?

(i) A 0.5 m NaBr solution has a higher vapour pressure than a 0.5 m BaCl<sub>2</sub> solution at the same

#### temperature

(ii) Pure water freezes at the higher temperature than pure methanol

(iii) a 0.1 m NaOH solution freezes at a lower temperature than pure water

#### Choose the correct answer from the codes given below

```
a) (i), (ii) and (iii) b) (i) and (ii)
c) (ii) and (iii) d) (i) and (iii)
```

#### Answer : D

Question : The vapour pressure of benzene at 30°C is 121.8 mm. By adding 15 g of non-volatile solute in 250 g of benzene, its vapour pressure is decreased to 120.2 mm. The molecular weight of solute is :

a) 156.6 g b) 267.4 g c) 356.3 g d) 467.4 g

### Answer : C

Question : Pure benzene freezes at 5.45°C. A 0.374 m solution of tetrachloroethane in benzene freezes at 3.55°C. The Kf for benzene is:

a) 0.508 b) 5.08 c) 50.8 d) 508

#### **Answer: B**

Question : 0.450 g of urea (mol.wt.60) in 22.5 g of water show 0.170°C of elevation in boiling point. The molal elevation constant of water is:

a) 0.051°C b) 0.51°C c) 5.1°C d) 0.83°C

#### Answer: B

#### **Question : The colligative property is not represented by :**

a) elevation in boiling point

- b) osmotic pressure
- c) optical activity
- d) relative lowering of vapour pressure

#### Answer : C

Question : 20 g of a substance were dissolved in 500 mL of water and the osmotic pressure of the solution was found to be 600 mm of mercury at 15°C. The molecular weight of substance is :

a) 998 b) 1028

c) 1098 d) 1198

Answer : D

Question : Which one of the statements given below concerning properties of solutions, describes a colligative effect?

- a) Boiling point of pure water decreases by the addition of ethanol
- b) Vapour pressure of pure water decreases by the addition of nitric acid
- c) Vapour pressure of pure benzene decreases by the addition of naphthalene
- d) Boiling point of pure benzene increases by the addition of toluene

#### Answer : C

Question : The average osmotic pressure of human blood is 7.8 bar at 37°C. What is the concentration of an aqueous NaCl solution that could be used in the blood stream?

a) 0.16 mol/L b) 0.31 mol / L c) 0.60 mol / L d) 0.45 mol / L

#### Answer: B

Question : A 5% solution (by mass) of cane sugar in water has freezing point of 271 K and freezing point of pure water is 273.15 K. The freezing point of a 5% solution (by mass) of glucose in water is

a) 271 K b) 273.15K c) 269.07 K d) 277.23 K

#### Answer : C

Question : The vapour pressure of pure benzene at a certain temperature is 0.850 bar. A non-volatile, nonelectrolyte solid weighing 0.5 g is added to 39.0 g of benzene (molar mass 78 g/mol). The vapour pressure of the solution then is 0.845 bar. What is the molecular mass of the solid substance?

a) 58 b) 180 c) 170 d) 145.

#### Answer : C

Question : 0.01 M solution of KCl and BaCl2 are prepared in water. The freezing point of KCl is found to be –2°C. What is the freezing point of BaCl2 to be completely ionised ?

a) – 3°C b) + 3°C c) – 2°C d) – 4°C

#### Answer : A

Question : "The importance of many pure substance in life depends on their composition." Which of the following statement justify the above fact?

a) 1 ppm of fluoride ions in water prevents tooth decay.

b) 1.5 ppm of fluoride ions causes tooth decay.

c) Concentration above 1.5 ppm can be poisonous.

### d) All of the above.

Question : Which of the following fluoride is used as rat poison?

a) CaF<sub>2</sub> b) KF

c) NaF d) MgF<sub>2</sub>

Question : Most of the processes in our body occur in

a) solid solution **b) liquid solution** 

c) gaseous solution d) colloidal solution

### Question : The term homogenous mixtures signifies that

- a) its composition is uniform throughout the mixture.
- b) its properties are uniform throughout the mixture.

### c) both composition and properties are uniform throughout the mixture.

d) neither composition nor properties are uniform throughout the mixture.

### **Question : Which of the following mixture is(are) called solution?**

- (i) water + ammonia (ii) water + acetone
- (iii) acetone + alcohol (iv) hexane + water
- **a) (i), (ii) and (iii)** b) (i), (iii) and (iv)
- c) (i) and (iv) d) (ii) and (iii)

#### Question : Which of the following is a quantitative description of the solution?

- a) Dilute b) Concentrated
- c) Saturated d) Molar

#### Question : When a solute is present in trace quantities the following expression is used

- a) Gram per million b) Milligram percent
- c) Microgram percent d) Parts per million

#### Question : Molarity of liquid HCl will be, if density of solution is 1.17 gm/cc

a) 36.5 **b) 32.05** 

c) 18.25 d) 42.10

### Question : 1 M, 2.5 litre NaOH solution is mixed with another 0.5 M, 3 litre NaOH solution. Then find out the molarity of resultant solution

#### a) 0.80 M b) 1.0 M

#### c) 0.73 M d) 0.50 M

### Question : An X molal solution of a compound in benzene has mole fraction of solute equal to 0.2. The value of X is

a) 14 **b) 3.2** 

c) 1.4 d) 2

#### Question : The molarity of the solution containing 7.1 g of Na2SO4 in 100 ml of aqueous solution is

a) 2 M b) 0.5 M

c) 1 M d) 0.05 M

Question : The vapour pressure of pure benzene at 25°C is 640 mm Hg and that of solution of solute A is 630 mm Hg. The molality of solution is

**a) 0.2 m** b) 0.4 m

c) 0.5 m d) 0.1 m

#### Question : 4.0 g of NaOH is dissolved in 100 ml solution. The normality of the solution is

a) 0.1 N b) 0.5 N

c) 4.0 N d) 1.0 N

#### **Question : The molarity of pure water is**

a) 50 M b) 18 M

c) 55.6 M d) 100 M

### Question : An aqueous solution of glucose is 10% in strength. The volume in which 1 g mole of it is dissolved, will be

a) 9 litre **b) 1.8 litre** 

c) 8 litre d) 0.9 litre

#### Question : 10 g of NaCl is dissolved in 106g of the solution. Its concentration is

a) 100 ppm b) 0.1 ppm

c) 1 ppm d) 10 ppm

### Question : On adding a solute to a solvent having vapour pressure 0.80 atm, vapour pressure reduces to 0.60 atm. Mole fraction of solute is

a) 0.25 b) 0.75

c) 0.50 d) 0.33

Question : 2.5 litres of NaCl solution contain 5 moles of the solute. What is the molarity?

a) 5 molar b) 2 molar

c) 2.5 molar d) 12.5 molar

Question : The mole fraction of the solute in one molal aqueous solution i

a) 0.009 b) 0.018

c) 0.027 d) 0.036

# Question : 5 ml of N HCl, 20 ml of N/2 H2SO4 and 30 ml of N/3 HNO3 are mixed together and volume made to one litre. The normality of the resulting solution is

a)N5

b)N10

c)N20

d) N/40

Question : 25ml of a solution of barium hydroxide on titration with a 0.1 molar solution of hydrochloric acid gave a titre value of 35ml. The molarity of barium hydroxide solution was

a) 0.07 b) 0.14

c) 0.28 d) 0.35

### Question : Mole fraction of the solute in a 1.00 molal aqueous solution is

a) 0.1770 b) 0.0177

c) 0.0344 d) 1.7700

#### Question : What is the normality of a 1 M solution of H3PO4 ?

a) 0.5 N b) 1.0 N

c) 2.0 N d) 3.0 N

### Question : The volume of 4 N HCl and 10 N HCl required to make 1 litre of 6 N HCl are

a) 0.75 litre of 10 N HCl and 0.25 litre of 4 N HCl

b) 0.50 litre of 4 N HCl and 0.50 litre of 10 N HCl

### c) 0.67 litre of 4 N HCl and 0.33 litre of 10 N HCl

d) 0.80 litre of 4 N HCl and 0.20 litre of 10 N HCl

### Question : Molarity of H2SO4 is 18 M. Its density is 1.8 g/ml. Hence molality is

a) 36 b) 200

### **c) 500** d) 18

Question : 200 ml of water is added to 500 ml of 0.2 M solution. What is the molarity of this diluted solution ?

a) 0.5010 M b) 0.2897 M

c) 0.7093 M d) 0.1428 M

Question : How many grams of concentrated nitric acid solution should be used to prepare 250 mL of 2.0M HNO3 ? The concentrated acid is 70% HNO3

a) 90.0 g conc. HNO3 b) 70.0 g conc. HNO<sub>3</sub>

c) 54.0 g conc. HNO3 d) 45.0 g conc. HNO<sub>3</sub>

# Question : For preparing 0.1 N solution of a compound from its impure sample of which the percentage purity is known, the weight of the substance required will be

a) Less than the theoretical weight

#### b) More than the theoretical weight

c) Same as the theoretical weight

d) None of these

# Question : If N/10 50 ml H2SO4, N/3 30 ml HNO3, N/2 10 ml HCl is mixed and solution is made to 1L. Then normality of resultant solution is

a)*N/*20

b)*N/*40

#### c) 50/N

d) *N* 

### Question : A solution made by dissolving 40 g NaOH in 1000 g of water is

a) 1 molar b) 1 normal

c) 1 molal d) None of these

#### Question : Which of the following concentration terms is/are independent of temperature?

a) Molality only

#### b) Molality and mole fraction

- c) Molarity and mole fraction
- d) Molality and normality

# Question : A solution is prepared by dissolving 10 g NaOH in 1250 mL of a solvent of density 0.8 mL/g. The molality of the solution in mol kg–1 is

#### a) 0.25 b) 0.2

c) 0.008 d) 0.0064

Question : Which of the following units is useful in relating concentration of solution with its vapour pressure?

a) mole fraction b) parts per million

c) mass percentage d) molality

Question : Which of the following concentration unit is independent of temperature ?

a) Normality b) Molarity

c) Formality d) Molality

#### Question : Which of the following factor do not affect solubility of solid solute in liquid ?

- a) Temperature **b) Pressure**
- c) Nature of solute d) All of these

Question : When a solid solute is added to the solvent, some solute dissolves and its concentration increases in solution. This process is known as \_\_\_\_\_. Some solute particles in solution collide with the solid solute particles and get separated out of solution. This process is known as \_\_\_\_\_.

- a) Crystallization, dissolution.
- b) Dissolution, saturation.
- c) Saturation, crystallization.
- d) Dissolution, crystallization.

#### Question : At the state of dynamic equilibrium, for solute + solvent solution.

- a) Rate of dissolution = Rate of unsaturation.
- b) Rate of dissolution = Rate of unsaturation.
- c) Rate of dissolution = Rate of saturation
- d) Rate of crystallization = Rate of saturation.

#### **Question : Which of the following statements is incorrect?**

a) A solution in which no more solute can be dissolved at the same temperature and pressure is called a saturated solution.

- b) An unsaturated solution is one in which more solute can be dissolved at the same temperature.
- c) The solution which is in dynamic equilibrium with undissolved solute is the saturated solution.

d) The minimum amount of solute dissolved in a given amount of solvent is its solubility.

# Question : On dissolving sugar in water at room temperature solution feels cool to touch. Under which of the following cases dissolution of sugar will be most rapid ?

- a) Sugar crystals in cold water.
- b) Sugar crystals in hot water.
- c) Powdered sugar in cold water.

#### d) Powdered sugar in hot water.

Question : The solubility of a solid in a liquid is significantly affected by temperature changes.

Solute + Solvent  $\rightarrow$  Solution.

The system being in a dynamic equilibrium must follow

Le-chatelier's principle. Considering the Le-chatelier's

principle which of the following is correct?

a)  $\triangle H_{sol} > 0$ ; solubility  $\uparrow$ ; temperature  $\downarrow$ 

**b)** △H<sub>sol</sub> < 0; solubility ↓ ; temperature ↑

- c)  $\triangle H_{sol} > 0$ ; solubility 1; temperature 1
- d)  $\triangle H_{sol} < 0$ ; solubility  $\downarrow$ ; temperature  $\uparrow$

# Question : The statement "If 0.003 moles of a gas are dissolved in 900 g of water under a pressure of 1 atmosphere, 0.006 moles will be dissolved under a pressure of 2 atmospheres", illustrates

a) Dalton's law of partial pressure

- b) Graham's law
- c) Raoult's law

### d) Henry's law

# Question : According to Henry's law, the amount of gas that will dissolve in blood plasma or any other liquid is determined by which of these factor?

#### a) Solubility of the gas in the liquid.

- b) The total pressure of the gas mixture .
- c) pH of the liquid.
- d) The osmotic pressure of the gas mixture.

# Question : Henry's law constant of oxygen is $1.4 \times 10-3$ mol. lit-1. atm-1 at 298 K. How much of oxygen is dissolved in 100 ml at 298 K when the partial pressure of oxygen is 0.5 atm?

a) 1.4 g b) 3.2 g

c) 22.4 mg d) 2.24 mg

#### Question : At equillibrium the rate of dissolution of a solid solute in a volatile liquid solvent is \_\_\_\_\_.

a) less than the rate of crystallisation.

b) greater than the rate of crystallisation.

#### c) equal to the rate of crystallisation.

d) zero

Question : A beaker contains a solution of substance 'A'. Precipitation of substance 'A' takes place when small amount of 'A' is added to the solution. The solution is \_\_\_\_\_.

a) saturated **b) supersaturated** 

# Question : Maximum amount of a solid solute that can be dissolved in a specified amount of a given liquid solvent does not depend upon \_\_\_\_\_.

a) Temperature b) Nature of solute

c) Pressure d) Nature of solvent

Question : Low concentration of oxygen in the blood and tissues of people living at high altitude is due to \_\_\_\_\_.

a) low temperature

### b) low atmospheric pressure

- c) high atmospheric pressure
- d) both low temperature and high atmospheric pressure

### Question : Value of Henry's constant *K*H \_\_\_\_\_.

#### a) increases with increase in temperature.

- b) decreases with increase in temperature.
- c) remains constant.
- d) first increases then decreases.

### Question : The value of Henry's constant *K*H is \_\_\_\_\_.

a) greater for gases with higher solubility.

### b) greater for gases with lower solubility.

- c) constant for all gases.
- d) not related to the solubility of gases.

# Question : Which of the followingfactor(s) affect the solubility of a gaseous solute in the fixed volume of liquid solvent ?

(i) Nature of solute (ii) Temperature (iii) Pressure

### a) (i) and (iii) at constant T

- b) (i) and (ii) at constant P
- c) (ii) and (iii) only

d) (iii) only

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# Solutions MCQ Questions Class 12 Chemistry with Answers

**Solutions** 

1. Molarity is the number of moles of a solute dissolved per \_\_\_\_\_. (dm<sup>3</sup> of a solution, dm<sup>3</sup> of solvent, Kg of solvent)

2. Molality is defined as the number of moles of solute dissolved per \_\_\_\_\_. (dm<sup>3</sup> of solution, kg of solvent, kg of solute)

3. The solubility of a solute \_\_\_\_\_\_ with the increase of temperature. (increases, decreases, does not alter)

4. The loss of electron during a chemical reaction is known as \_\_\_\_\_\_. (Oxidation, Reduction, Neutralization)

5. The gain of electron during a chemical reaction is known as \_\_\_\_\_\_. (Oxidation, Reduction, Neutralization)

6. The ions, which are attracted towards the anode, are known as \_\_\_\_\_\_. (Anins, Cations, Positron).

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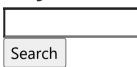
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7. The pH of a neutral solution is	(1.7, 7, 14)
8. A current of one ampere flowing for <b>Faraday)</b>	r one minute is equal to (One coulomb, 60 coulomb, one
9. A substance, which does not allow e <b>Conductor, Electrolyte)</b>	electricity to pass through, is known as (Insulator,
10. Such substances, which allow elect called ( <b>Electrolytes, Insula</b>	tricity to pass through them and are chemically decomposed, are <b>tors, Metallic conductors)</b>
11 is an example of strong a	acid. (Acetic Acid, Carbonic Acid, Hydrochloric Acid)
12 is an example of weak ac	cid. (Hydrochloric Acid, Acetic Acid, Sulphuric Acid)
13. When NH4Cl is hydrolyzed, the sol	lution will be (Acidic, Basic, Neutral)
14. When Na2CO3 is hydrolyzed, the s	solution will be (Acidic, Basic, Neutral)
15. When blue hydrated copper sulpha remains blue)	ate is heated (It changes into white, it turns black, it
16. Sulphur has the highest oxidation	number in (SO <sub>2</sub> , H <sub>2</sub> SO <sub>4</sub> , H <sub>2</sub> SO <sub>3</sub> )
17. The reaction between an acid and <b>Hydrolysis, Neutralization)</b>	a base to form a salt and water is called (Hydration,
18 is opposite of Neutraliza	ation. (Hydration, Hydrolysis, Ionization)
19. The substance having pH value 7 is	s (Basic, Acidic, Neutral)
20. An aqueous solution whose pH is z	zero is (Alkaline, Neutral, Strongly Acidic)
21. Solubility product of slightly solubl	le salt is denoted by (Kc, Kp, Ks <sub>p</sub> )
22. The increase of oxidation number i	is known as (Oxidation, Reduction, Hydrolysis)
23. The decrease of Oxidation number	r is known as (Oxidation, Reduction, Electrolysis)
24. One molar solution of glucose con	ntains gms of glucose per dm3 of solution. (180, 100, 342)
25. The number of moles of solute pre <b>Normality)</b>	esent per dm3 of solution is called (Molality, Molarity,
26. 'M' is the symbol used for represer	nting (Molality, Molarity, Normality)
27. 1 mole of H2SO4 is equal to	(98gms, 49gms, 180gms)
28. Buffer solution tends to	pH.
29. The logarithm of reciprocal of hydi	roxide ion is represented as (pH, pOH, pOH)
30. In water molecules surro	ound solute particles. (Hydration, Hydrolysis, Neutralization)

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# arch ywords



### II. Fill in the Blank

1. A mixture of two or more substances, which are homogeneously mixed, is called a \_\_\_\_\_.

2. is defined as the amount of solute dissolved in a given amount of solvent.

3. A solution is composed of two components \_\_\_\_\_ and \_\_\_\_\_.

4. A solution containing one mole of solute per dm<sup>3</sup> of solution is called one \_\_\_\_\_\_ solution.

5. Molarity is denoted by \_\_\_\_\_.

6. 1M solution of NaOH contains \_\_\_\_\_ gms of it dissolved per dm<sup>3</sup> of solution.

7. A solution containing one mole of solute dissolved by per kg of solvent is called \_\_\_\_\_\_ solution. 8. Molality is denoted by \_\_\_\_\_.

9. 1M solution of H<sub>2</sub>SO<sub>4</sub> contains \_\_\_\_\_ gms of it per kg of solvent.

10. The process in which ions are surrounded by water molecules is called \_\_\_\_\_.

11. The water molecules attached with the hydrated substance are called \_\_\_\_\_\_.

12. Hydrated copper sulphate evolves \_\_\_\_\_\_ water molecules on heating.

13. The interaction between salt and water to produce acids and bases is called \_\_\_\_\_\_.

14. The products of ionic concentration in a saturated solution at a certain temperature are called the

15. Solubility product constant expressed as \_\_\_\_\_

- 16. The suppression of ionization by adding a common ion is called \_\_\_\_\_\_
- 17. The process of dissociation of an electrolyte into ions is known as \_\_\_\_\_

18. The chemical decomposition of a compound in a solution or in fused state brought about by a flow of

electric current is known as \_\_\_\_\_

- 19. Electrolysis is performed in an electrolytic cell, which is known as \_\_\_\_\_
- 20. The positive electrode of a voltmeter is called \_\_\_\_\_\_ and negative as \_\_\_\_\_\_.
- 21. A solution, which tends to resist changes in pH is called a \_\_\_\_\_\_ solution.

22. A mixture of acetic acid and sodium acetate acts as a \_\_\_\_\_

- 23. According to Sorenson \_\_\_\_\_\_ is defined as negative logarithm of the hydrogen ion concentration.
- 24. pH is mathematically expressed as \_\_\_\_\_.
- 25. The pH of a neutral solution is \_\_\_\_\_
- 26. \_\_\_\_\_ substances have pH values lower than 7.

27. \_\_\_\_\_ solutions have pH values more than 7.

- 28. Oxidation is \_\_\_\_\_ of electron.
- 29. Reduction is the \_\_\_\_\_ of electron.
- 30. Such chemical reactions in which the oxidation number of atoms or ions is changed are called

\_\_\_\_\_ reactions.

31. Oxidation number of a free element is \_\_\_\_\_.

- 32. Oxidation number of Oxygen in a compound is \_\_\_\_\_
- 33. The sum of oxidation number of any formula of a compound is \_\_\_\_\_\_.
- 34. The oxidation number of any ion is equal to the \_\_\_\_\_ on the ion.
- 35. \_\_\_\_\_\_ is the reaction in which an acid reacts with a base to form salt and water.
- 36. \_\_\_\_\_\_ are organic compounds which change colour in accordance with the pH of the medium.
- 37. An indicator that changes from colourless to pink in the presence of an alkaline solution is called

38. An indicator that changes from red to yellow in the presence of an alkaline solution is called \_\_\_\_\_\_

39. Dissociation constant is denoted by \_\_\_\_\_.

- 40. According to Bronsted-Lowry Concept, \_\_\_\_\_ is the donor of proton and \_\_\_\_\_ is the acceptor of proton.
- 41. According to Arrhenius, acid is substance that produces \_\_\_\_\_\_ ions when dissolved in water.
- 42. According to Arrhenius, base is a substance that produces \_\_\_\_\_\_ ions when dissolved in water.
- 43. When ionic product is less than ksp, the solution will \_\_\_\_\_\_.
- 44. When ionic product is greater than ksp, the solution will \_\_\_\_\_\_
- 45. The electrode at which oxidation takes place is called \_\_\_\_\_\_.
- 46. The electrode at which reduction takes place is called \_\_\_\_\_
- 47.  $H_3O$ + ion is called \_\_\_\_\_ ion.
- 48. The logarithm of reciprocal of hydroxyl ion (OH)- is called \_\_\_\_\_\_.
- 49. Aqueous solution of NH<sub>4</sub>Cl is \_\_\_\_\_\_ while that of NaHCO<sub>3</sub> is \_\_\_\_\_\_
- 50. The ionic product of [H+] and [OH-] of pure water is \_\_\_\_\_.
- 51. An increase in the oxidation number of an element or ion during a chemical change is called \_\_\_\_\_\_.
- 52. A decrease in the oxidation number of an element or ion during a chemical change is called \_\_\_\_\_\_.
- 53. The degree of dissociation \_\_\_\_\_\_ with the increase in temperature.
- 54. The degree of dissociation \_\_\_\_\_\_ with the dilution of electrolytic solution.
- 55. A \_\_\_\_\_\_ consists of an electrode immersed in solution of its ion.
- 56. The potential difference between the electrode and the solution of its salt at equilibrium position is

called \_\_\_\_\_ potential.

57. If the pH of a solution is 14, the solution is \_\_\_\_\_.

58. If the pH of a solution is 4, the solution is \_\_\_\_\_.

59. The oxidation number of Mn in KMnO<sub>4</sub> is \_\_\_\_\_\_.

60. The oxidation number of Fe in FeCl<sub>3</sub> is \_\_\_\_\_.

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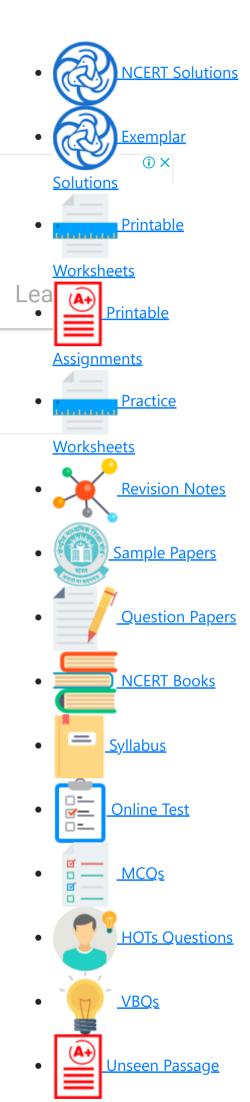
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# Surface Chemistry Class 12 Chemistry MCQ

Class 12 Chemistry students should refer to the following multiple-choice questions with answers for Surface Chemistry in standard 12. These MCQ questions with answers for Grade 12 Chemistry will come in exams and help you to score good marks

# Surface Chemistry MCQ Questions Class 12 Chemistry with Answers

Question : The physical adsorption of gases on the solid surface is due to:



(a) Covalent bond (b) Hydrogen bond (c) Ionic bond (d) Van der waal's forces

# Answer : D

Question : The electrical charge on a colloidal particle is observed by:

(a) Ultramicroscope (b) Scattering (c) Brownian movement (d) Electrophoresis

# **Answer : D**

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# **Question : Lyophilic colloids are stable due to:**

- (a) Small size of the particle
- (b) Large size of the particle
- (c) Layer of dispersion medium on the particle
- (d) Charge on the particle

# Answer : C

# **Question : Purple of cassius is colloidal solution of :**

(a) Silver (b) Lead(c) Gold (d) Mercury

# Answer : C

# Question : Milk is colloid in which :

- (a) Liquid is dispersed in liquid
- (b) Gas is dispersed in liquid
- (c) Sugar is dispersed in water
- (d) Solid is dispersed in liquid

# **Answer : A**

# **Question : The colloid is :**

(a) urea(b) blood(c) cane sugar (d) NaCl

# Answer: B

# Question : The movement of colloidal particles, under applied electric current is known as :

(a) electrodialysis(b) dialysis(c) electrophoresis(d) none of the above

# Answer : C

# Question : The size of colloidal particle is

(a) 10–3 to 10–9 m (b) 10–3 to 10–12 m (c) 10–6 to 10–9 m (d) 10–12 to 10–19 m

# Answer : C

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# **Question : A catalyst**

(a) changes the equilibrium constant(b) lowers the activation energy(c) increases the forward and backward reactions at different speeds.(d) follows same mechanism for the reaction.

# Answer : B

**Question : Which of the following is a lyophobic colloidal solution ?** 

- (a) Aqueous starch solution
- (b) Aqueous protein solution
- (c) Gold solution
- (d) Polymer solvent in some organic solvents

# Answer : C

Question : The density of gold is 19 g/cm<sup>3</sup>. If  $1.9 \times 10^{-4}$  g of gold is dispersed in one litre of water to give a sol having spherical gold particles of radius 10 nm, then the number of gold particles per mm<sup>3</sup> of the sol will be

(a)  $1.9 \times 10^{12}$  (b)  $6.3 \times 10^{14}$ (c)  $6.3 \times 10^{10}$  (d)  $2.4 \times 10^{6}$ 

# Answer : D

# Question : Which of the following electrolyte will have maximum flocculation value for Fe(OH)<sub>3</sub> sol?

(a) NaCl (b) Na<sub>2</sub>S (c) (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> (d) K<sub>2</sub>SO<sub>4</sub>

# **Answer : A**

# **Question : Preparation of Lyophobic sols by chemical method involves**

(a) double decomposition(b) oxidation & reduction

- (c) hydrolysis
- (d) all of these

# **Answer : D**

Question : A colloidal solution is subjected to an electrical field. The particles move towards anode. The coagulation of same sol is studied using NaCl, BaCl<sub>2</sub> and AlCl<sub>3</sub> solutions. Their coagulating power should be

- (a)  $NaCl > BaCl_2 > AlCl_3$
- (b)  $BaCl_2 > AlCl_3 > NaCl$
- (c)  $AlCl_3 > BaCl_2 > NaCl$
- (d)  $BaCl_2 > NaCl > AlCl_3$

# Answer : C

Question : Under the influence of an electric field, the particles in a sol migrate towards cathode.

The coagulation of the same sol is studied using NaCl, Na<sub>2</sub>SO<sub>4</sub> and Na<sub>3</sub>PO<sub>4</sub> solutions. Their

## coagulating values will be in the order

(a)  $NaCl > Na_2SO_4 > Na_3PO_4$ (b)  $Na_2SO_4 > Na_3PO_4 > NaCl$ (c)  $Na_3PO_4 > Na_2SO_4 > NaCl$ (d)  $Na_2SO_4 > NaCl > Na_3PO_4$ 

## Answer : A

# Question : Gold numbers of protective colloids A, B, C and D are 0.50, 0.01, 0.10 and 0.005, respectively. The correct order of their protective powers is

```
(a) D < A < C < B (b) C < B < D < A
(c) A < C < B < D (d) B < D < A < C
```

## Answer : C

## Question : The disease kala azar is caused by

(a) colloidal antimony(b) milk of magnesia(c) argyrols(d) colloidal gold

# **Answer : A**

Question : Which one of the following impurities present in colloidal solution cannot be removed by electrodialysis?

(a) Sodium chloride(b) Potassium sulphate(c) Urea(d) Calcium chloride

# Answer : C

## **Question : Choose the correct option**

# Assertion: A catalyst is more effective in finely divided form. Reason: Finely divided form has more surface area.

(a) If both Assertion and Reason are correct and the Reason is a correct explanation of the Assertion.

(b) If both Assertion and Reason are correct but Reason is not a correct explanation of the Assertion.

(c) If the Assertion is correct but Reason is incorrect.

(d) If both the Assertion and Reason are incorrect.

# **Answer : A**

# Question : Adsorbed acetic acid on activated charcoal is:

(a) adsorber (b) absorber

(c) adsorbent (d) adsorbate

## **Question : Adsorption is always**

(a) endothermic

# (b) exothermic

(c) exothermic in case of physical and endothermic in case of chemical

Question : Which is not correct regarding the physical adsorption of a gas on surface of solid ?

(a) On increasing temperature, adsorption increases continuously

(b) Enthalpy and entropy changes are negative

(c) Adsorption is more for some specific substance

(d) Reversible

## Question : How many layers are adsorbed in chemical adsorption ?

(a) One (b) Two

(c) Many (d) Zero

## Question : Adsorption due to strong chemical forces is called

- (a) Chemisorption (b) Physisorption
- (c) Reversible adsorption (d) Both (b) and (c)

## Question : In physical adsorption, gas molecules are bound on the solid surface by

(a) chemical forces (b) electrostatic forces

## (c) gravitational forces (d) van der Waal's forces

## **Question : Which of the following statements is not correct ?**

(a) Physical adsorption is due to van der Waal's forces

# (b) Chemical adsorption first decreases with increase in temperature.

- (c) Physical adsorption is reversible
- (d) Adsorption energy for a chemical adsorption is generally greater than that of physical adsorption.

## Question : Adsorption of gases on solid surface is exothermic reaction because

(a) free energy increases (b) enthalpy is positive

(c) entropy increases (d) enthalpy is negative

# Question : The gas which is least adsorbed on charcoal (under identical conditions) is

- (a) HCl (b) O<sub>2</sub>
- (c) CO<sub>2</sub> (d) NH<sub>3</sub>

## Question : Adsorption is accompanied by

(a) decrease in enthalpy and increase in entropy

(b) increase in enthalpy and increase in entropy

# (c) decrease in enthalpy and decrease in entropy

(d) increase in enthalpy and decrease in entropy

# **Question : Choose the incorrect statement in respect of physisorption?**

(a) It is not specific in nature

(b) It arises because of van der Waal's force

(c) It is reversible in nature

(d) Enthalpy of adsorption is in the range 80-240 kJ mol<sup>-1</sup>

# Question : The term 'sorption' stands for \_\_\_\_\_

- (a) absorption
- (b) adsorption

## (c) both absorption and adsorption

(d) desorption

## **Question : Extent of physisorption of a gas increases with \_\_\_\_\_.**

(a) increase in temperature.

## (b) decrease in temperature.

- (c) decrease in surface area of adsorbent.
- (d) decrease in strength of van der Waal's forces.

# Question : Extent of adsorption of adsorbate from solution phase increases with \_\_\_\_\_.

## (a) increase in amount of adsorbate in solution.

- (b) decrease in surface area of adsorbent.
- (c) increase in temperature of solution.
- (d) decrease in amount of adsorbate in solution.

## Question : Which of the following is not a favourable condition for physical adsorption ?

- (a) High pressure
- (b) Negative  $\triangle H$
- (c) Higher critical temperature of adsorbate

## (d) High temperature

## Question : Physical adsorption of a gaseous species may change to chemical adsorption with

(a) decrease in temperature

# (b) increase in temperature

- (c) increase in surface area of adsorbent
- (d) decrease in surface area of adsorbent

Question : In physisorption adsorbent does not show specificity for any particular gas because

## (a) involved van der Waal's forces are universal.

(b) gases involved behave like ideal gases.

(c) enthalpy of adsorption is low.

(d) it is a reversible process.

## Question : Which of the following is an example of absorption ?

(a) Water on silica gel .

## (b) Water on calcium chloride.

- (c) Hydrogen on finely divided nickel.
- (d) Oxygen on metal surface.

Question : For adsorption of a gas on a solid, the plot of  $\log x/m$  vs log P is linear with slope equal to (*n* being whole number)

- (a) *k* (b) log *k*
- (c) *n* (d) 1/*n*

Question : The adsorption of a gas on a solid surface varies with pressure of the gas in which of the following manner

# (a) Fast $\rightarrow$ slow $\rightarrow$ independent of the pressure

- (b) Slow  $\rightarrow$  fast  $\rightarrow$  independent of the pressure
- (c) Independent of the pressure  $\rightarrow$  fast  $\rightarrow$  slow
- (d) Independent of the pressure  $\rightarrow$  slow  $\rightarrow$  fast

## Question : Hair cream is an example of

- (a) gel (b) sol
- (c) aerosol (d) foam

# **Question :** In Freundlich adsorption isotherm, the value of 1/*n* is :

# (a) between 0 and 1 in all cases

- (b) between 2 and 4 in all cases
- (c) 1 in case of physical adsorption
- (d) 1 in case of chemisorption

# Question : Which is adsorbed in maximum amount by activated charcoal ?

# (a) N<sub>2</sub> (b) CO<sub>2</sub>

# (c) Cl<sub>2</sub> (d) O<sub>2</sub>

Question : If dispersed phase is a liquid and the dispersion medium is a solid, the colloid is known as

(a) a sol **(b) a gel** 

(c) an emulsion (d) a foam

Question : According to Freundlich adsorption isotherm, the amount of gas adsorbed at very high pressure

(a) reaches a constant limiting value

## (b) goes on increasing with pressure

- (c) goes on decreasing with pressure
- (d) increase first and decreases later with pressure

## Question : Which is not correct regarding the adsorption of a gas on surface of solid?

## (a) On increasing temperature, adsorption increases continuously

- (b) Enthalpy and entropy changes are -ve
- (c) Adsorption is more for some specific substance
- (d) This Phenomenon is reversible

## Question : Alloy is an example of

- (a) gel (b) solidified emulsion
- (c) solid solution (d) sol

## Question : Which of the following is related to adsorption?

- (i)  $\triangle H = -ve$  (ii)  $\triangle S = -ve$
- (iii)  $-T \triangle S = -ve$  (iv)  $\triangle G = -ve$
- (a) (i), (ii) and (iv) (b) (ii) and (iii)
- (c) (iii) only (d) (i), (iii) and (iv)

## Question : The role of a catalyst in a reversible reaction is to

- (a) increase the rate of forward reaction
- (b) decrease the rate of backward reaction
- (c) alter the equilibrium constant of the reaction

## (d) allow the equilibrium to be achieved quickly

## **Question : Catalytic poisons act by :**

(a) making the products chemically inactive.

(b) increasing the rate of the backward reaction.

(c) chemical combination with any one of the reactants.

(d) preferential adsorption on the catalyst surface.

**Question : A catalyst :** 

(a) lowers the activation energy

(b) changes the rate constant

(c) changes the product

(d) itself destroyed in the reaction

https://www.studiestoday.com/mcq-chemistry-cbse-class-12-chemistry-surface-chemistry-mcqs-set-292285.html

# Question : Active charcoal is a good catalyst because it

(a) is made up of carbon atoms.

(b) is very reactive.

# (c) has more adsorption power.

(d) has inert nature toward reagents.

# Question : Which of the following kind of catalysis can be explained by the adsorption theory ?

- (a) Homogeneous catalysis
- (b) Acid base catalysis

# (c) Heterogeneous catalysis

(d) Enzyme catalysis

Question : According to the adsorption theory of catalysis, the speed of the reaction increases because-

# (a) Adsorption lowers the activation energy of the reaction

(b) The concentration of reactant molecules at the active centres of the catalyst becomes high due to strong adsorption

- (c) In the process of adsorption, the activation energy of the molecules becomes large
- (d) Adsorption produces heat which increases the speed of the reaction

# Question : Catalyst increases the rate of reaction by

(a) decreasing threshold energy

# (b) decreasing activation energy

- (c) increasing activation energy
- (d) decreasing equilibrium constant

# Question : A catalyst can affect reversible reaction by

(a) changing equilibrium constant

(b) slowing forward reaction

(d) None of these

# Question : Which one of the following is an example of homogeneous catalysis ?

(a) Haber's process of synthesis of ammonia

(b) Catalytic conversion of  $SO_2$  to  $SO_3$  in contact process

(c) Catalytic hydrogenation of oils

(d) Acid hydrolysis of methyl acetate.

# Question : Identify the correct statement regarding enzymes

(a) Enzymes are specific biological catalysts that cannot be poisoned.

(b) Enzymes are normally heterogeneous catalysts that are very specific in their action.

(c) Enzymes are specific biological catalysts that can normally function at very high temperatures (T  $\approx$ 1000K).

# (d) Enzymes are specific biological catalysts that possess well-defined active sites.

## **Question : A biological catalyst is**

(a) an enzyme (b) a carbohydrate

(c) an amino acid (d) a nitrogenous base

# Question : The action of enzymes in living system is to :

- (a) supply energy to tissues
- (b) enhance immunity
- (c) circulate oxygen

# (d) enhance the rate of biochemical reactions.

# Question : Hydrolysis of urea is an example of

(a) homogenous catalysis (b) heterogenous catalysis

(c) biochemical catalysis (d) zeolite catalysis

# Question : The efficiency of an enzyme in catalysing a reaction is due to its capacity

- (a) to form a strong enzyme-substrate complex
- (b) to decrease the bond energies of substrate molecule
- (c) to change the shape of the substrate molecule

## (d) to lower the activation energy of the reaction

# Question : What is the role of molybdenum in Haber's process for manufacture of ammonia?

(a) As catalytic poison (b) As a catalytic promoter

(c) As a catalyst (d) As a reactant

# Question : Which of the following step(s) is/are not involved in the mechanism of adsorption theory of heterogeneous catalyst?

(i) Diffusion of reactants to the surface of the catalyst.

(ii) Sorption of reactant molecules on the surface of the catalyst.

(iii) Occurrence of chemical reaction on the catalyst's surface through formation of an intermediate.

(iv) Desorption of reaction products from the catalyst's surface.

(v) Diffusion of reaction products away from the catalyst's surface.

- (a) (i) only (b) (ii) and (iv)
- **(c) (ii) only** (d) (i), (ii) and (v)

# Question : Milk is a colloid in which a

# (a) liquid is dispersed in a liquid

- (b) solid is dispersed in a liquid
- (c) gas is dispersed in a liquid
- (d) sugar is dispersed in a liquid

# Question : Butter is a colloid formed when

- (a) Fat is dispersed in water
- (b) Fat globules are dispersed in water

# (c) Water is dispersed in fat

(d) None of the above

# Question : The size of colloidal particles is between

- (a)  $10^{-7} 10^{-9}$  cm (b)  $10^{-9} 10^{-11}$  cm
- (c)  $10^{-5} 10^{-7}$  cm (d)  $10^{-2} 10^{-3}$  cm

# Question : An aerosol is a :

# (a) dispersion of a solid or liquid in a gas

- (b) dispersion of a solid in a liquid
- (c) dispersion of a liquid in a liquid
- (d) solid solution

# Question : An example of dispersion of a liquid in a gas is :

(a) milk (b) vegetable oil

(c) foam (d) mist

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CBSE vide Circular No.Acad-51/2021 dated 5th July, 2021, notified that in the session 2021-2022, Board Examinations would be conducted in two terms, i.e.. Term I and Term II. This decision was taken due to the uncertainty arising out of COVID 19 Pandemic. Term I...

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Ministry of Education, Govt of India vide letter No. F.No. 12-5/2020-IS-4 dated 16.12.2021 has intimated that under the banner Azadi ka Amrit Mahotsav the National Yogasanasports Federation has decided to run a project of 750 million Surya Namaskar from 01 January 2022...

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Question : Malachite is an ore of

(a) Silver (b) Mercury

(c) Magnesium

(d) Copper

Answer : D

Question : The chief ore of Hg is

(a) Pyrolusite

(b) Bauxite

(c) Galena

(d) Cinnabar



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# **Answer : D**

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- **Question : Which of the following is greatest paramagnetic?**
- (a) Cu<sup>+</sup> (c) Fe<sup>3+</sup>
- (d) Cu<sup>2+</sup>
- Answer : C

# (a) Iron

**Question : The transition element which shows the highest oxidation state is:** 

- (b) Vanadium
- (c) Manganese
- (d) Chromium

# **Answer : C**

# **Question : Gun metal is**

(a) Cu + Zn (b) Cu + Sn + Zn(c) Cu + Sn (d) Zn + Sn

# **Answer: B**

# **Question : Transition elements form coloured ions due to :**

(a) d-d transition (b) fully filled d-orbitals (c) smaller atomic radii (d) availability of s-electrons

# **Answer: A**

# **Question : CuSO<sub>4</sub> and KCN react to produce**

(a) CuCN<sub>2</sub> (b) CuCN (c)  $K_3[Cu(CN)_4]$ (d)  $K_4[Cu(CN)_6]$ 

# **Answer : C**

- (b) Fe<sup>2+</sup>

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# **Question : German silver is an alloy of:**

(a) Fe, Cr, Ni

(b) Ag, Cu Au

(c) Cu, Zn, Ni

(d) Cu, Zn, Sn

**Answer : C** 

# **Question : The composition of duralumin is**

(a) Al 94%, Mg 6%
(b) Cu 56%, Zn 24%, Ni 20%
(c) Cu 95%, Al 5%
(d) Al 95%, Cu 4%, Mn 0.5%, Mg 0.5%

## **Answer : D**

## Question : The number of water molecules in Mohr's salt is

(a) 2

(b) 4

(c) 6

(d) 8

# Answer : C

# Question : Philosopher's wool on heating with BaO at 1100°C produces

(a)  $Ba + ZnCl_2$ (b)  $BaCdO_2$ (c)  $BaZnO_2$ (d)  $BaO_2 + Zn$ 

# Answer : C

# Question : Which of the following is not an ore of iron?

(a) limonite

(b) casiterite

(c) magnetite

(d) none of these

## **Answer: B**

## Question : Lanthanide for which +II and +III oxidation states are common is

(a) La

(b) Pr

(c) Ce

(d) Eu

# **Answer : D**

# Question : F<sub>2</sub> is formed by reacting K<sub>2</sub>MnF<sub>6</sub> with

(a) SbF<sub>5</sub> (b) MnF<sub>3</sub>

(c) KSbF<sub>6</sub> (d) MnF<sub>4</sub>

# **Answer : A**

Question : The colour imparted by Co(II) compounds to glass is:

(a) Green(b) Deep blue(c) Yellow(d) Red

# Answer : B

# Question : Which of the following radioisotopes is used as anticancerous?

(a) Na-24(b) C-14(c) U-235(d) Co-60

# **Answer : D**

# Question : Which of the following compound is coloured?

(a) TiCl<sub>3</sub>
(b) FeCl<sub>3</sub>
(c) CoCl<sub>2</sub>
(d) All of these

# Answer : D

# Question : To obtain silver from silver amalgam, it is heated in vessel which is made of

(a) Cu

(b) Fe

(c) Ni

(d) Zn

# **Answer : B**

# Question : The colour of copper sulphide is

(a) Blue (b) Black

(c) Red

(d) Green

# Answer : B

# Question : Which one of the following statements concerning lanthanide elements is false?

(a) Lanthanides are separated from one another by ion exchange method

- (b) The ionic radii of trivalent lanthanides steadily increase with increase in atomic number
- (c) All lanthanides are highly dense metals
- (d) Most characteristic oxidation state of lanthanides is +3

# Question : What is wrong about transition metals?

(a) Diamagnetic

(b) Paramagnetic

(c) Form complexes

(d) Shows variable oxidation state

# **Answer : A**

Question : Correct electronic configuration of Cr (Z = 24) is (a)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^7 4s^1$ (b)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^1$  (c)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d 7 4s^2$ (d)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d 6 4s^2$ 

# Answer : B

# Question : Which of the following configuration is correct for iron ?

(a)  $1s^2, 2s^2 2p^6, 3s^2 3p^6 3d4$ (b)  $1s^2, 2s^2 2p^6, 3s^2 3p^6 3d^6 4s^2$ (c)  $1s^2, 2s^2 2p6, 3s^2 3p^6 3d2$ (d)  $1s^2, 2s^2 2p6, 3s^2 3p^6 3d24s^2$ 

## Answer: B

# Question : Which one of the following ions has electronic configuration [Ar] 3d<sup>6</sup>?

(a) Ni<sup>3+</sup>
(b) Mn<sup>3+</sup>
(c) Fe<sup>3+</sup>
(d) Co<sup>3+</sup>

(At. Nos. Mn = 25, Fe = 26, Co = 27, Ni = 28)

## **Answer : D**

# Question : Which of the following element does not belong to first transition series?

(a) Fe

(b) V

(c) Ag

(d) Cu

# Answer : C

# Question : $(n-1)d^{10}ns^2$ is the general electronic configuration of

(a) Fe, Co, Ni

(b) Cu, Ag, Au

(c) Zn, Cd, Hg

(d) Se, Y, La

# Answer : C

# Question : The last electron in d-block elements goes to

(a) (n-1) d

(b) nd

(c) np (d)

(n-1) s

Answer : A

# Question : The elements which exhibit both vertical and horizontal similarites are

- (a) inert gas elements
- (b) representative elements
- (c) rare elements
- (d) transition elements

# **Answer : D**

Question : An atom has electronic configuration  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^3 4s^2$  in which group would it be placed?

- (a) Fifth
- (b) Fifteenth
- (c) Second
- (d) Third

## **Answer: A**

# Question : In 3d-series atomic number (Z) varies from

- (a) *Z* = 21- 30
- (b) *Z* = 22 30
- (c) *Z* = 20 30
- (d) *Z* = 31- 40

# **Answer: A**

## Question : The valence shell of transition elements consists of

- (a) nd orbitals
- (b) (n-1) d orbitals
- (c) ns np nd orbitals
- (d) (n-1) d ns orbitals

# **Answer : D**

# Question : Number of unpaired electrons in $Ni^{2+}(Z=28)$ is

(b) 2

(c) 6

(d) 8

**Answer: B** 

# Question : Which of the following element is not a member of transition elements ?

(a) Zn

# (b) Pt

(c) Ce

(d) Mo

# Answer : C

Question : The number of unpaired electrons in gaseous species of Mn<sup>3+</sup>, Cr<sup>3+</sup> and V<sup>3+</sup> respectively are.

(a) 4, 3 and 2

(b) 3, 3 and 2

(c) 4, 3 and 2

(d) 3, 3 and 3

# Answer : C

	Answer : A		
	(d) Ca		
	(c) V		
	(b) Ti		
	(a) Sc		
Question : The first element in the 3d-transition series is			

Question : Which of the following has more unpaired d-electrons?

(a) Zn<sup>+</sup>

(b) Fe<sup>2+</sup>

(c) Ni<sup>+</sup>

(d) Cu<sup>+</sup>

# Answer : B

Question : The number of unpaired electrons in a nickel atom in ground state are (At. No. of Ni = 28)

(a) 2

(b) 5

(d) 7

# Answer : A

Question : Which one of the following is an example of non-typical transition elements ?

(a) Li, K, Na

(b) Be, Al, Pb

(c) Zn, Cd, Hg

(d) Ba, Ga, Sr.

# Answer : C

# **Question : Which of the following has the maximum number of unpaired electrons?**

1-1	<b>T</b> . O	
(a)	112	+

- (b) Fe2+
- (c) Cr+
- (d) Cu+

# Answer : C

# Question : The outer electronic configuration of Ag is $4d^{10} 5s^1$ , it belongs to

- (a) 5th period, group 4
- (b) 4th period, group 5
- (c) 5th period, group 11
- (d) 6th period, group 9

# Answer : C

# **Question : Manganese belongs to**

- (a) 1st transition series
- (b) 2nd transition series
- (c) 3rd transition series
- (d) 4th transition series

# **Answer : A**

# Question : The no. of unpaired electrons in $Mn^{7+}$ ions (At. no. of Mn = 25) is

- (a) 0
- (b) 1
- (c) 2
- (d) 3

# Answer : A

Question : Which one of the following species is paramagnetic?

(a) N2

(b) Co

(c) Cu+

(d) Zn

Answer : B

# Question : Which of the following species is/are paramagnetic?

Fe<sup>2+</sup>, Zn0, Hg<sup>2+</sup>, Ti<sup>4+</sup>

- (a) Fe<sup>2+</sup> only
- (b) Zn0 and Ti<sup>4+</sup>
- (c)  $Fe^{2+}$  and  $Hg^{2+}$
- (d) Zn0 and  $Hg^{2+}$

# Answer : A

# Question : In first transition series, the melting point of Mn is low because

- (a) due to  $d^{10}$  configuration, metallic bonds are strong
- (b) due to  $d^7$  configuration, metallic bonds are weak
- (c) due to  $d^5$  configuration, metallic bonds are weak
- (d) None of these

# Answer : C

# Question : The transition metals have a less tendency to form ions due to

- (a) high ionisation energy
- (b) low heat of hydration of ions
- (c) high heat of sublimation
- (d) All of these

# Answer : D

# Question : The common oxidation states of Ti are

- (a) + 2 and + 3
- (b) + 3 and + 4
- (c) 3 and 4
- (d) + 2, + 3 and + 4

# Answer : D

# Question : Maximum oxidation state is shown by

(a) Os

(b) Mn

(c) Co

(d) Cr

# Answer : A

Question : Which one of the elements with the following outer orbital configurations may exhibit the largest number of oxidation states?

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(a)  $3d^54s^1$ 

(b) 3*d*<sup>5</sup>4*s*<sup>2</sup>

(c)  $3d^24s^2$  (d)  $3d^34s^2$ 

# Answer : B

# Question : Which of the following pairs has the same size?

(a) Fe<sup>2+</sup> , Ni<sup>2+</sup>

(b) Zr<sup>4+</sup> , Ti<sup>4+</sup>

(c) Zr<sup>4+</sup> , Hf <sup>4+</sup>

(d) Zn2+ , Hf  $^{4+}$ 

# Answer : C

Question : For the four successive transition elements (Cr, Mn, Fe and Co), the stability of +2 oxidation state will be there in which of the following order?

- (a) Mn > Fe > Cr > Co
- (b) Fe > Mn > Co > Cr
- (c) Co > Mn > Fe > Cr
- (d) Cr > Mn > Co > Fe

# Answer : A

# Question : Iron exhibits +2 and + 3 oxidation states. Which of the following statements about iron is incorrect ?

(a) Ferrous oxide is more basic in nature than the ferric oxide.

- (b) Ferrous compounds are relatively more ionic than the corresponding ferric compounds.
- (c) Ferrous compounds are less volatile than the corresponding ferric compounds.
- (d) Ferrous compounds are more easily hydrolysed than the corresponding ferric compounds.

# **Answer : D**

Question : Four successive members of the first row transition elements are listed below with their atomic numbers. Which one of them is expected to have the highest third ionization enthalpy?

(a) Vanadium (Z = 23)

(b) Chromium (Z = 24)

(c) Manganese (Z = 25)

(d) Iron (Z = 26)

Answer : C

Question : Of the following outer electronic configurations of atoms, the highest oxidation state is achieved by which one of them ?

(a)  $(n-1)d^3 ns^2$ 

(b)  $(n-1)d^5 ns^1$ 

(c)  $(n-1)d^8 ns^2$ 

(d)  $(n-1)d^5 ns^2$ 

# Answer : D

# Question : For *d* block elements the first ionization potential is of the order

- (a) Zn > Fe > Cu > Cr
- (b) Sc = Ti < V = Cr
- (c) Zn < Cu < Ni < Co

(d) V > Cr > Mn > Fe

# **Answer : A**

# Question : Which of the following does not represent the correct order of the properties indicated ?

(a)  $Ni^{2+} > Cr^{2+} > Fe^{2+} > Mn^{2+}$  (size) (b) Sc > Ti > Cr > Mn (size) (c)  $Mn^{2+} > Ni^{2+} < Co^{2+} < Fe^{2+}$  (unpaired electron) (d)  $Fe^{2+} > Co^{2+} > Ni^{2+} > Cu^{2+}$  (unpaired electron) Answer : A

# Question : Zinc and mercury do not show variable valency like *d*-block elements because

- (a) they are soft
- (b) their *d*-shells are complete
- (c) they have only two electrons in the outermost subshell
- (d) their *d*-shells are incomplete

# Answer : B

# Question : Which of the following transition element shows the highest oxidation state ?

(a) Mn

(b) Fe

(d) Cr

**Answer : A** 

Question : Which of the following elements does not showvariable oxidation states?

(a) Copper

(b) Iron

(c) Zinc

(d) Titanium

# Question : Which one of the following transition elements does not exhibit variable oxidation state?

(a)	Ni	

- (b) Cu
- (c) Fe
- (d) Sc

# Answer : D

Question : Electronic configuration of a transition element X in +3 oxidation state is [Ar]3d<sup>5</sup>. What is its atomic number ?

(a) 25

(b) 26

(c) 27

(d) 24

## **Answer: B**

# Question : Metallic radii of some transition elements are given below. Which of these elements will have highest density ?

Element Fe Co Ni Cu

Metallic radii/pm 126 125 125 128

(a) Fe

(b) Ni

(c) Co

(d) Cu

Answer : D

# **Question : Transition metals mostly are**

(a) diamagnetic

(b) paramagnetic

(c) neither diamagnetic nor paramagnetic

(d) both diamagnetic and paramagnetic

Answer : B

Question : Transition metals usually exhibit highest oxidation states in their

(a) chlorides

(b) fluorides

(c) bromides

(d) iodides

## Answer: B

# **Question : Which of the following statements is incorrect?**

(a) Zn,Cd and Hg due to presence of completely filled *d*-orbitals  $[(n-1)d^{10}ns^2]$  are not studied along with other transition metals.

(b) Zn, Cd and Hg have low m.p and are comparitively softer than other transition metals.

(c) Metallic bond made by elements with  $d^{5}$  configuration is stronger as compared to metalic bond made by elements with  $d^{3}$  configuration.

(d) Metals of  $5^d$  series forms strong metallic bonds as compared with metals of  $3^d$  series.

## **Answer : A**

## **Question : Which of the following is incorrect?**

- (a) Mn shows oxidation state of +7 in MnF<sub>7</sub>
- (b) Fe and Co shows +3 oxidation state in FeX<sub>3</sub> and CoF<sub>3</sub>.
- (c) V shows oxidation state of + 5 in VF<sub>5</sub>.
- (d) Cu does not shows +2 oxidation state with I-.

# Answer : A

## Question : Which of the following is not correct about transition metals?

- (a) Their melting and boiling points are high
- (b) Their compounds are generally coloured
- (c) They can form ionic or covalent compounds
- (d) They do not exhibit variable valency

## Answer : D

## **Question : Transition elements**

- (a) have low melting point
- (b) exhibit variable oxidation states
- (c) do not form coloured ions

# Answer : B

Question : Which one of the following ions is the most stable in aqueous solution?

(a) V<sup>3+</sup>

(b) Ti<sup>3+</sup>

(c) Mn<sup>3+</sup>

(d) Cr<sup>3+</sup>

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# **Answer : D**

# Question : Which one of the following does not correctly represent the correct order of the property indicated against it?

(a) Ti < V < Cr < Mn : increasing number of oxidation states

(b)  $Ti^{3+} < V^{3+} < Cr^{3+} < Mn^{3+}$ : increasing magnetic moment

(c) Ti < V < Cr < Mn : increasing melting points

(d) Ti < V < Mn < Cr : increasing 2nd ionization enthalpy

Answer : C

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Question : The ion or group detected by K<sub>2</sub>[HgI<sub>4</sub>] is

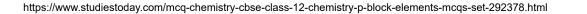
 $\begin{array}{cccc} (a) & NO & (b) & Cl^- \\ (c) & NH_2^- & (d) & NH_4^+ \end{array}$ 

Answer : D

**Question : Which gas cannot be collected over water ?** 

(a) O<sub>2</sub>
(b) PH<sub>3</sub>
(c) N<sub>2</sub>
(d) SO<sub>2</sub>

Answer : D





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#### **Question : The correct order of in creasing oxidising power is :**

- $(a) \quad Cl_2 < Br_2 < F_2 < I_2 \\$
- (b)  $F_2 < Br_2 < Cl_2 < I_2$
- (c)  $F_2 > Br_2 > Cl_2 > I_2$
- $(d) \quad I_2 < Br_2 < Cl_2 < F_2$

#### Answer : D

**Question : H<sub>2</sub>S does not produce metallic sulphide with :** 

(a) CuCl<sub>2</sub>
(b) COCl<sub>2</sub>
(c) CdCl<sub>2</sub>
(d) ZnCl<sub>2</sub>

#### Answer : B

#### Question : The correct order of acid strength of oxyacids is:

- (a) HClO > HClO<sub>2</sub> > HClO<sub>3</sub> > HClO<sub>4</sub>
- (b) HClO<sub>2</sub> > HClO<sub>3</sub> > HClO > HClO<sub>4</sub>
- (c)  $HClO_4 > HClO_3 > HClO_2 > HClO_3$
- (d)  $HClO_3 > HClO_4 > HClO > HClO_2$

#### Answer : C

#### Question : Chlorine acts as a bleaching agent only in presence of:presence of

- (a) Sunlight
- (b) Moisture
- (c) Dry air
- (d) Pure oxygen

#### Answer : B

#### **Question : The laughing gas is:**

(a) NO
(b) N<sub>2</sub>O
(c) N<sub>2</sub>O4
(d) N<sub>2</sub>O5

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#### Answer : B

**Question : The correct decreasing order of basic strength is:** 

- (a)  $AsH_3 > SbH_3 > PH_3 > NH_3$
- $(b) \quad SbH_3 > AsH_3 > PH_3 > NH_3$
- (c)  $NH_3 > PH_3 > AsH_3 > SbH_3$
- $(d) \quad PH_3 > AsH_3 > SbH_3 > NH_3$

Answer : C

#### **Question : lodide of Millon's base is**

(a) HIO<sub>3</sub>
(b) K<sub>2</sub>HgI<sub>4</sub>
(c) NH<sub>2</sub>HgO.HgI
(d) Hg(NH<sub>2</sub>)I

#### Answer : C

#### Question : Which of the following is not the characteristic of interhalogen compounds?

- (a) they are covalent
- (b) more reactive than halogens
- (c) have low B.P. and high volatile
- (d) quite unstable but not explosive

#### Answer : C

#### **Question : The tribasic acid is :**

(a) H<sub>3</sub>PO<sub>4</sub>
(b) H<sub>3</sub>PO<sub>3</sub>
(c) H<sub>3</sub>PO<sub>2</sub>
(d) HPO<sub>3</sub>

#### **Answer : A**

#### **Question : The strongest oxidising agent is**

(a) F<sub>2</sub>
(b) Cl<sub>2</sub>
(c) l<sub>2</sub>
(d) Br<sub>2</sub>

#### **Answer : A**

#### Question : The correct order of solubility in water for He, Ne, Ar, Kr, Xe is

(a) He > Ne > Ar > Kr > Xe
(b) Xe > Kr > Ar > Ne > He
(c) Ne > Ar > Kr > He > Xe
(d) Ar > Ne > He > Kr > Xe

#### **Answer : B**

#### Question : Which of the following compound is a tribasic acid?

(a) H<sub>3</sub>PO<sub>2</sub>(b) H<sub>3</sub>PO<sub>4</sub>

(c) H<sub>3</sub>PO<sub>3</sub> (d) H<sub>4</sub>P<sub>2</sub>O<sub>7</sub>

#### Answer : B

#### Question : The mixture of concentrated HCl and HNO<sub>3</sub> made in 3 : 1 ratio contains

(a) CIO<sub>2</sub>
(b) NOCI
(c) NCI<sub>3</sub>
(d) N<sub>2</sub>O<sub>4</sub>

#### Answer : B

#### Question : The element which forms oxides in all oxidation states varying from +1 to +5 is

(a)	Ν
(b)	Ρ
(c)	As
(d)	Sb

**Answer : A** 

#### Question : The statement true for $N_3^-$ is

(a) it has non-linear structure

(b) it is called pseudohalogen

(c) the formal oxidation state of nitrogen in this anion is -1

(d) it is isoelectronic with  $NO_2$ 

#### **Answer : D**

#### Question : For electron affinity of halogens which of the following is correct?

(a) Br > F
(b) F > CI
(c) Br > CI
(d) F > I

#### **Answer : D**

Question : The number of P–O–P bridges in the structure of phosphorus pentoxide and phosphorus trioxide are respectively

(a) 6, 6 (b) 5, 5 (c) 5, 6 (d) 6, 5

#### Answer : A

#### **Question : Tincture of iodine is**

(a) Aqueous solution of  $I_2$ 

(b) Solution of  $I_2$  in aqueous KI

(c) Alcoholic solution of  $I_2$ 

(d) Aqueous solution of KI

#### Answer : C

Question : Which two of the following salts are used for preparing iodized salt?

(i) KIO<sub>3</sub> (ii) KI (iii) I<sub>2</sub> (iv) HI

(a) (i) and (ii)
(b) (i) and (iii)
(c) (ii) and (iv)
(d) (iii) and (iv)

Answer : A

Question : Ionic radii (in Å) of As<sup>3+</sup>, Sb<sup>3+</sup> and Bi<sup>3+</sup> follow the order

(a)  $As^{3+} > Sb^{3+} > Bi^{3+}$ (b)  $Sb^{3+} > Bi^{3+} > As^{3+}$ (c)  $Bi^{3+} > As^{3+} > Sb^{3+}$ (d)  $Bi^{3+} > Sb^{3+} > As^{3+}$ 

#### **Answer : AD**

#### Question : Which of the following statements is not correct for nitrogen?

- (a) Its electronegativity is very high
- (b) *d*-orbitals are available for bonding
- (c) It is a typical non-metal
- (d) Its molecular size is small

#### **Answer : B**

#### Question : Collectively the elements of group 15 are called -

(a) pnicogens

- (b) pnicopens
- (c) nicopen
- (d) None of these

#### Answer : A

#### Question : Which one of the following elements is most metallic?

(a) P

- (b) As
- (c) Sb
- (d) Bi

#### Answer : D

#### **Question : Which of the following statement is incorrect for group 15 elements ?**

(a) Order of ionization enthalpies is  $\triangle iH_1 < \triangle iH_2 < \triangle iH_3$ 

- (b) The boiling point and melting point increases from top to bottom in the group
- (c) Dinitrogen is a gas while all others are solids
- (d) All statements are correct

#### **Answer : B**

#### Question : Which of the follow group 15 element forms metallic bonds in elemental state ?

(a) As

(b) P

(c) Sb

(d) Bi

#### Answer : D

Question : The three important oxidation states of phosphorus are

(a) -3, +3 and +5

- (b) -3, +3 and -5
- (c) -3, +3 and +2
- (d) -3, +3 and +4

#### Answer : A

#### **Question : Nitrogen is relatively inactive element because**

(a) its atom has a stable electronic configuration

(b) it has low atomic radius

(c) its electronegativity is fairly high

(d) dissociation energy of its molecule is fairly high

#### Answer : D

#### Question : Which of the following has the highest $p\Pi - p\Pi$ bonding tendency ?

(a) N

(b) P

(c) As

(d) Sb

#### Question : Pick out the wrong statement.

(a) Nitrogen has the ability to form  $p\pi$ - $p\pi$  bonds with itself.

(b) Bismuth forms metallic bonds in elemental state.

(c) Catenation tendency is higher in nitrogen when compared with other elements of the same group.

(d) Nitrogen has higher first ionisation enthalpy when compared with other elements of the same group.

#### **Answer : C**

#### Question : Nitrogen forms N<sub>2</sub>, but phosphorus is converted into P<sub>4</sub> from P, the reason is

- (a) Triple bond is present between phosphorus atom
- (b)  $p\mathbf{\pi} p\mathbf{\pi}$  bonding is strong
- (c)  $p\pi p\pi$  bonding is weak
- (d) Multiple bond is formed easily

#### **Answer : C**

#### Question : What causes nitrogen to be chemically inert?

- (a) Multiple bond formation in the molecule
- (b) Absence of bond polarity
- (c) Short internuclear distance
- (d) High bond energy

#### Answer : D

Question : Among the 15<sup>th</sup> group elements, as we move from nitrogen to bismuth, the pentavalency becomes less pronounced and trivalency becomes more pronounced due to

(a) Non metallic character

(b) Inert pair effect

(c) High electronegativity (d) Large ionization energy

#### **Answer: B**

#### Question : Pentavalence in phosphorus is more stable when compared to that of nitrogen even though they belong to same group. This is due to

(a) dissimilar electronic configuration

(b) due to presence of vacant d-orbitals

(c) reactivity of phosphorus

(d) inert nature of nitrogen

#### Answer: B

#### **Question : Which one has the lowest boiling point ?**

(a) NH<sub>3</sub>
(b) PH<sub>3</sub>
(c) AsH<sub>3</sub>
(d) SbH<sub>3</sub>

#### **Answer : B**

#### Question : Most acidic oxide among the following is -

(a) N<sub>2</sub>O<sub>5</sub>
(b) P<sub>2</sub>O<sub>5</sub>
(c) N2O<sub>4</sub>

(d) As2O<sub>3</sub>

#### **Answer : A**

#### Question : Which of the following species has the highest dipole moment?

(a) NH<sub>3</sub>

(b) PH<sub>3</sub>

(c)  $AsH_3$ 

(d)  $SbH_3$ 

#### **Question : The correct decreasing order of basic strength is:**

(a)  $AsH_3 > SbH_3 > PH_3 > NH_3$ 

(b)  $SbH_3 > AsH_3 > PH_3 > NH_3$ 

(c)  $NH_3 > PH_3 > AsH_3 > SbH_3$ 

(d)  $PH_3 > AsH_3 > SbH_3 > NH_3$ 

#### Answer : C

#### Question : Which of the following fluorides does not exist?

(a) NF<sub>5</sub>
(b) PF<sub>5</sub>
(c) AsF<sub>5</sub>
(d) SbF<sub>5</sub>

#### Answer : A

#### Question : The p-block element of group 15 that forms predominantly basic oxide is

(a) N

(b) P

(c) As

(d) Bi

Answer : D

#### Question : With respect to protonic acids, which of the following statements is correct ?

(a)  $PH_3$  is more basic than  $NH_3$ 

(b)  $PH_3$  is less basic than  $NH_3$ 

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(c)  $PH_3$  is equally basic as  $NH_3$ 

(d) PH<sub>3</sub> is amphoteric while NH<sub>3</sub> is basic

#### Answer : B

#### **Question : PCI<sub>5</sub> is possible but NCI<sub>5</sub> does not exist :**

- (a) in N, d-sub-shell is absent
- (b) ionization energy of N is very high
- (c) it does not like Cl
- (d) None of these

#### **Answer : A**

#### Question : Maximum covalency of nitrogen is \_\_\_\_\_.

(a) 3

(b) 5

(c) 4

(d) 6

#### Answer : C

Question : Elements of group-15 form compounds in +5 oxidation state. However, bismuth forms only one well characterised compound in +5 oxidation state. The compound is

(a) Bi<sub>2</sub>O<sub>5</sub>
(b) BiF<sub>5</sub>
(c) BiCl<sub>5</sub>
(d) Bi<sub>2</sub>S<sub>5</sub>

#### Answer : B

#### Question : Pure nitrogen is prepared in the laboratory by heating a mixture of

(a)  $NH_4OH + NaCl$ (b)  $NH_4NO_3 + NaCl$ (c)  $NH_4Cl + NaOH$ (d)  $NH_4Cl + NaNO_2$ .

#### Answer : D

#### Question : On heating ammonium dichromate and barium azide separately we get

(a)  $N_2$  in both cases

(b)  $N_2$  with ammonium dichromate and NO with barium azide

(c)  $N_2O$  with ammonium dichromate and  $N_2$  with barium azide

(d)  $N_2O$  with ammonium dichromate and  $NO_2$  with barium azide

#### Answer : A

**Question : In Haber's process for the manufacture of NH<sub>3</sub> :** 

(a) finely divided nickel is used as a catalyst

(b) finely divided iron is used as a catalyst

- (c) finely divided molybdenum is used as a catalyst
- (d) no catalyst is necessary

#### Answer: B

#### **Question : Ammonia on reaction with hypochlorite anion can form :**

(a) NO (b)  $N_2H_4$ (c)  $NH_4CI$ (d) Both (b) and (c)

#### Answer : D

#### Question : $NH_3$ gas is dried over :

(a) CaO (b) HNO<sub>3</sub> (c) P<sub>2</sub>O<sub>5</sub> (d) CuSO<sub>4</sub>

#### **Answer : A**

#### Question : The shape of ammonia molecule is

(a) tetrahedral

(b) pyramidal

(c) planar triangle

(d) octahedral

#### **Answer : B**

#### Question : When ammonia is heated with cupric oxide, a molecule of ammonia will

- (a) gain 3 electrons
- (b) lose 3 electrons
- (c) gain 2 electrons
- (d) lose 2 electrons

#### Answer : B

#### Question : In which the NH<sub>3</sub> is not used ?

(a) Cold storage

(b) Anaesthetic

(c) Manufacture of rayon and plastic

(d) None of these

Answer : B

Question : Liquid ammonia bottles are opened after cooling them in ice for sometime. It is because liquid NH3

(a) Brings tears to the eyes

(b) Has a high vapour pressure

- (c) Is a corrosive liquid
- (d) Is a mild explosive

#### Answer : B

#### Question : Ammonia is generally manufactured for fertilizers by the reaction

(a)  $2NH_4CI + Ca(OH)_2 \rightarrow CaCI_2 + 2H_2O + 2NH_3$ 

- (b) By passing an electric discharge in a mixture of  $N_2$  and  $H_2$
- (c) By passing a mixture of N<sub>2</sub> and H<sub>2</sub> under high pressure and moderate temperature over a catalyst
- (d) None of these

#### Answer : B

#### **Question : Nitrogen dioxide cannot be obtained by heating :**

(a) KNO<sub>3</sub>
(b) Pb(NO<sub>3</sub>)<sub>2</sub>
(c) Cu(NO<sub>3</sub>)<sub>2</sub>
(d) AgNO<sub>3</sub>

#### **Answer : A**

#### Question : Which of the following oxides is neutral ?

(a) N<sub>2</sub>O<sub>3</sub>
(b) N<sub>2</sub>O<sub>4</sub>
(c) N<sub>2</sub>O<sub>5</sub>
(d) N<sub>2</sub>O

#### Answer : B

#### Question : The bonds present in $N_2O_5$ are :

(a) only ionic(b) covalent and coordinate(c) only covalent(d) covalent and ionic

#### Answer: B

**Question : Which of the following oxides of nitrogen is a coloured gas?** (a)  $N_2O$ 

(d) N2 (b) N0 (c) N2O5 (d) NO2

Answer : D

#### Question : Which one of the following is not an use of ammonia ?

(a) To produce various nitrogenous fertilizers.

(b) In manufacture of nitric acid

(c) As a refrigerate

(d) In the pickling of stainless steel

#### Answer : D

#### Question : In which one of the following oxides of nitrogen, one nitrogen atom is not directly linked

to oxygen?
(a) NO
(b) N <sub>2</sub> O <sub>4</sub>
(c) N <sub>2</sub> O
(d) N <sub>2</sub> O <sub>3</sub>

#### Answer : C

Question : Which of the following oxides of nitrogen reacts with FeSO <sub>4</sub> to form a dark brown
compound
(a) N <sub>2</sub> O
(b) NO
(c) NO <sub>2</sub>
(d) N2O <sub>3</sub>

#### Answer : B

#### Question : Which oxide of nitrogen is obtained on heating ammonium nitrate at 250°C?

- (a) Nitric oxide
- (b) Nitrous oxide
- (c) Nitrogen dioxide
- (d) Dinitrogen tetraoxide

#### Answer: B

#### Question : Which of the following can be used as an anaesthesia ?

(a) N<sub>2</sub>O
(b) NO
(c) NCl<sub>3</sub>
(d) NO<sub>2</sub>

#### **Answer : A**

#### Question : A deep brown gas is formed by mixing two colourless gases which are

(a) NO<sub>2</sub> and O<sub>2</sub> (b) N<sub>2</sub>O and NO (c) NO and O<sub>2</sub> (d) NH<sub>3</sub> and HCl

#### Answer : C

#### Question : Which of the following elements does not form stable diatomic molecules ?

(a) lodine

(b) Phosphorus

(c) Nitrogen

(d) Oxygen

Answer : B

#### Question : The catalyst used in the manufacture of HNO<sub>3</sub> by Ostwald's process is :

(a) platinum gauze(b) vanadium pentoxide(c) finely divided nickel

(d) platinum black .

Question : Concentrated nitric acid, upon long standing, turns yellow brown due to the for	nation of
(a) NO	
(b) NO <sub>2</sub>	

(c) N<sub>2</sub>O

(d) N<sub>2</sub>O<sub>4</sub>

#### Answer : B

<b>Question</b> :	Which	of the	following	trihalide is	unstable?
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(a) NF<sub>3</sub>`

(b) AsCl<sub>3</sub>

(c) SbBr<sub>3</sub>

(d) NCl<mark>3</mark>

#### Answer : D

<b>Question</b> :	Which of the	following	element will	form acidic	oxides of	type E <sub>2</sub> O <sub>3</sub> ?
						-J <u>-</u> - J - J -

(a) As

(b) Sb

(c) Bi

(d) P

#### Answer : D

#### Question : Which of the following is the strongest reducing agent ?

(a) NH<sub>3</sub>

(b) PH<sub>3</sub>

(c) BiH<sub>3</sub>

(d)  $SbH_3$ 

#### Answer : C

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Question : Chottky defect defines imperfection in the lattice structure of



a) solid b) gas c) liquid d) plasma

**Answer: A** 

Question : An AB<sub>2</sub> type structure is found in

a) NaCl b) CaF<sub>2</sub> c)  $AI_2O_3$ d) N<sub>2</sub>O

#### **Answer: B**

https://www.studiestoday.com/mcq-chemistry-cbse-class-12-chemistry-solid-state-mcqs-292181.html



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### Question : An element (atomic mass 100 g/mol) having bcc structure has unit cell edge 400 pm. The density of element is (No. of atoms in bcc, Z = 2).

a) 2.144 g/cm<sup>3</sup>
b) 7.289 g/cm<sup>3</sup>
c) 5.188 g/cm<sup>3</sup>
d) 10.376 g/cm<sup>3</sup>

#### Answer : C

#### Question : What is the coordinatio number of sodium in Na<sub>2</sub>O?

a)	6
a)	6

b) 4

c) 8

d) 2

#### **Answer : B**

#### Question : The compound, found in nature in gas phase but ionic in solid state is

a) PCl<sub>5</sub> b) CCl₄

c) PCl<sub>3</sub>

d) POCl<sub>3</sub>

#### **Answer : A**

### Question : The Ca<sup>2+</sup> and F<sup>-</sup> are located in CaF<sub>2</sub> crystal, respectively at face centred cubic lattice points and in

#### -

a) Tetrahedral voids

- b) Half of tetrahedral voids
- c) Octahedral voids
- d) Half of octahedral voids

#### Answer : B

#### Question : The coordination number in hcp is

a) 6

b) 12

c) 18

- d) 24
- Answer : B

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#### **Question : The space lattice of graphite is**

a) Cubic

b) Tetragonal

c) Rhombic

d) Hexagonal

#### Answer : D

Question : Coordination numbers of Zn<sup>2+</sup> and S<sup>2-</sup> in the crystal structure of wurtzite are

a) 4, 4 b) 6, 6 c) 8, 4 d) 8, 8

#### Answer : A

Question : Gold has a face centred cubic lattice with an edge length of the unit cube of 407 pm. Assuming the closest packing, the diameter of the gold atom is

a) 576.6 pmb) 287.8 pmc) 352.5 pm

d) 704.9 pm

Answer : B

#### **Question : Which is not correct about the Schottky defects?**

a) Both cations and anions are missing from their lattice sites without affecting the stoichiometry of the compound

b) Because of presence of holes the lattice energy decreases.

c) The presence of holes causes the density of the crystal to decrease.

d) The defect increases the electrical of the ions into the holes.

#### Answer : B

#### Question : The existence of a substance in more than one solid modifications is known as

- a) isomorphism
- b) Polymorphism
- c) Amorphism
- d) Allotropy

#### Answer: B

Question : An element (atomic mass = 100 g / mol) having bcc structure has unit cell edge 400 pm. Then, density of the element is

a) 10.376 g/cm<sup>3</sup>
b) 5.188 g/cm<sup>3</sup>
c) 7.289 g/cm<sup>3</sup>
d) 2.144 g/cm<sup>3</sup>

#### **Answer : B**

Question : If AgI crystallises in zinc blende structure with I- ions at lattice points. What fraction of

tetrahedral voids is occupied by Ag+ ions?

a) 25%

b) 50%

c) 100%

d) 75%

**Answer : B** 

Question : Which set of following characteristics for ZnS crystal is correct?

- a) Coordination number (4 : 4); ccp;  $Zn^{2+}$  ion in the alternate tetrahedral voids
- b) Coordination number (6 : 6); hcp;  $Zn^{2+}$  ion in all tetrahedral voids.
- c) Coordination number (6 : 4); hcp;  $Zn^{2+}$  ion in all octahedral voids
- d) Coordination number (4 : 4); ccp;  $Zn^{2+}$  ion in all tetrahedral voids.

#### **Answer : A**

#### Question : Which one of the following statements about packing in solids is incorrect ?

- a) Coordination number in bcc mode of packing is 8.
- b) Coordination number in hcp mode of packing is 12.
- c) Void space in hcp mode of packing is 32%.
- d) Void space is ccp mode of packing is 26%.

#### Answer : C

### Question : Sodium metal crystallizes in a body centred cubic lattice with a unit cell edge of 4.29Å. The radius of sodium atom is approximately :

a) 5.72Å b) 0.93Å c) 1.86Å d) 3.22Å

#### Answer : C

#### **Question : Choose the correct option**

#### Assertion :No compound has both Schottky and Frenkel defects. Reason : Both defects change the density of the solid.

- a) If both Assertion and Reason are correct and Reason is the correct explanation of Assertion.
- b) If both Assertion and Reason are correct, but Reason is not the correct explanation of Assertion.
- c) If Assertion is correct but Reason is incorrect.
- d) If both the Assertion and Reason are incorrect.

#### Answer : D

#### **Question : Choose the correct option**

#### Assertion : Stability of a crystal is reflected in the magnitude of its melting. Reason : The stability of a crystal depends upon the strength of the interparticle attractive force.

- a) If both Assertion and Reason are correct and Reason is the correct explanation of Assertion.
- b) If both Assertion and Reason are correct, but Reason is not the correct explanation of Assertion.
- c) If Assertion is correct but Reason is incorrect.
- d) If both the Assertion and Reason are incorrect.

#### Answer : A

#### **Question : Choose the correct option**

#### Assertion : Due to Frenkel defect, there is no effect on the density of the crystalline solid. Reason : In Frenkel defect, no cation or anion leaves the crystal.

a) If both Assertion and Reason are correct and Reason is the correct explanation of Assertion.b) If both Assertion and Reason are correct, but Reason is not the correct explanation of Assertion.c) If Assertion is correct but Reason is incorrect.d) If both the Assertion and Reason are incorrect.

#### **Answer : A**

**Question : Choose the correct option** 

#### Assertion: On heating ferromagnetic or ferrimagnetic substances, they become paramagnetic. Reason: The electrons change their spin on heating.

- a) If both Assertion and Reason are correct and Reason is the correct explanation of Assertion.
- b) If both Assertion and Reason are correct, but Reason is not the correct explanation of Assertion.
- c) If Assertion is correct but Reason is incorrect.
- d) If both the Assertion and Reason are incorrect.

#### **Answer : A**

#### Question : Which of the following is not a characteristic property of solids?

- a) Intermolecular distances are short.
- b) Intermolecular forces are weak.
- c) Constituent particles have fixed positions.
- d) Solids oscillate about their mean positions.

#### Answer : B

#### Question : Most crystals show good cleavage because their atoms, ions or molecules are

- a) weakly bonded together
- b) strongly bonded together
- c) spherically symmetrical
- d) arranged in planes

#### Answer : D

### Question : "Crystalline solids are anisotropic in nature. What is the meaning of anisotropic in the given tement?

a) A regular pattern of arrangement of particles which repeats itself periodically over the entire crystal.

b) Different values of some of physical properties are shown when measured along different directions in the same crystals.

c) An irregular arrangement of particles over the entire crystal.

d) Same values of some of physical properties are shown when measured along different directions in the same crystals.

#### Answer: B

#### **Question : A crystalline solid**

a) changes abruptly from solid to liquid when heated

b) has no definite melting point

c) undergoes deformation of its geometry easily

d) has an irregular 3-dimensional arrangements

#### Answer : A

#### Question : Which of the following is not a characteristic of a crystalline solid ?

- a) Definite and characteristic heat of fusion.
- b) Isotropic nature.

c) A regular periodically repeated pattern of arrangement of constituent particles in the entire crystal.

d) A true solid

#### **Answer : B**

#### **Question : Which of the following is not a crystalline solid?**

- a) KCl
- b) CsCl
- c) Glas
- d) Rhombic S

#### Answer : C

Question : Which of the following statements about amorphous solids is incorrect ?

- a) They melt over a range of temperature
- b) They are anisotropic
- c) There is no orderly arrangement of particles
- d) They are rigid and incompressible

#### Answer : B

#### Question : Which of the following is not a crystalline solid?

- a) KCl
- b) CsCl
- c) Glass
- d) Rhombic S

#### Answer : C

a) Graphite (C)

b) Quartz glass (SiO<sub>2</sub>)

c) Chrome alum

d) Silicon carbide (SiC)

Answer : B

#### Question : Which of the following statement is not true about amorphous solids ?

a) On heating they may become crystalline at certain temperature.

b) They may become crystalline on keeping for long time.

c) Amorphous solids can be moulded by heating.

d) They are anisotropic in nature.

#### **Answer : D**

**Question :** The sharp melting point of crystalline solids is due to \_\_\_\_\_.

- a) a regular arrangement of constituent particles observed over a short distance in the crystal lattice.
- b) a regular arrangement of constituent particles observed over a long distance in the crystal lattice.
- c) same arrangement of constituent particles in different directions.
- d) different arrangement of constituent particles in different directions.

#### **Answer: B**

Question : Why some glass objects from ancient civilisations are found to become milky in appearance?

- a) Glass is a crystalline solid, milky appearance is due to its crystalline nature.
- b) Glass is amorphous but on heating it become crystalline at some temperature.
- c) Because of reaction of glass with impurities present in the atmosphere.
- d) None of these.

#### **Answer: B**

Question : Which of the following amorphous solid is used as photovoltaic material for conversion of unlight into electricity?

- a) Quartz glass
- b) Quartz
- c) Silicon

d) Both a) and b)

#### **Answer : C**

#### **Question : Solid CH<sub>4</sub> is**

- a) ionic solid
- b) covalent solid
- c) molecular solid

d) does not exist

#### **Answer : C**

Question : An example of a covalent crystalline solid is:

a) Si

b) Al

c) NaF

d) Ar

#### Question : Among solids, the highest melting point is exhibited by

- a) Covalent solids
- b) lonic solids
- c) Pseudo solids
- d) Molecular solids

#### Answer : A

#### Question : Which of the following exists as covalent crystals in the solid state ?

- a) lodine
- b) Silicon
- c) Sulphur
- d) Phosphorus

#### Answer : B

#### Question : The major binding force of diamond, silicon and quartz is

- a) electrostatic force
- b) electrical attraction
- c) covalent bond force
- d) non-covalent bond force

#### Answer : C

#### **Question : In graphite electrons are :**

- a) localised on each carbon atom
- b) spread out between the sheets
- c) localised on every third carbon atom
- d) present in antibonding orbital.

#### **Answer : B**

Question : Which one of the following forms a molecular solid when solidified?

a) Silicon carbide

b) Calcium fluoride

c) Rock salt

d) Methane

Answer : D

Question : Which of the following is a network solid ?

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a) SO<sub>2</sub> (solid)

b) I2

c) Diamond

d) H<sub>2</sub>O (Ice)

#### Answer : C

#### Question : Which of the following solids is not an electrical conductor?

a) Mg (s)

b) TiO (s)

c) l<sub>2</sub> (s)

d) H<sub>2</sub>O (s)

#### Answer : C

Question : lodine molecules are held in the crystals lattice by \_\_\_\_\_.

a) london forces

b) dipole-dipole interactions

c) covalent bonds

d) coulombic forces

#### Answer : A

#### Question : Which of the following is not the characteristic of ionic solids?

a) Very low value of electrical conductivity in the molten state.

b) Brittle nature.

c) Very strong forces of interactions.

d) Anisotropic nature.

#### **Answer : A**

**Question :** Graphite is a good conductor of electricity due to the presence of \_\_\_\_\_.

a) lone pair of electrons

b) free valence electrons

c) cations

d) anions

Answer : B

Question : Graphite cannot be classified as \_\_\_\_\_.

a) conducting solid

b) network solid

c) covalent solid

d) ionic solid

#### **Answer : D**

#### Question : Which of the following cannot be regarded as molecular solid ?

(i) SiC (Silicon carbide)

(ii) AIN

(iii) Diamond

(iv) I2

a) (i), (ii) and (iii) b) (ii) and (iii) c) (iv) d) (ii) and (iv)

#### **Answer: A**

Question : Crystals can be classified into basic crystal units, equal to	
a) 7	
b) 4	
c) 14	
d) 2	

#### **Answer : A**

#### Question : How many three dimensional crystal lattice are possible?

a) 20

b) 7

c) 14

d) 10

#### **Answer : C**

#### Question : Which of the following is the correct increasing order of packing efficiency for hcp, bcc and simple cubic lattice?

#### a) hcp < bcc < simple cubic

b) bcc < hcp < simple cubic

c) simple cubic < bcc < hcp

d) simple cubic < hcp < bcc

**Answer : C** 

Question : In face-centred cubic lattice, a unit cell is shared equally by how many unit cells

a) 2

b) 4

d) 8

#### Answer : C

Question : Percentages of free space in cubic close packed structure and in body centered packed structure are respectively

a) 30% and 26%

b) 26% and 32%

c) 32% and 48%

d) 48% and 26%

#### Answer : B

#### Question : The empty space in the body centred cubic lattice is

a) 68%

b) 52.4%

c) 47.6%

d) 32%

(e) 74%

#### **Answer : D**

#### Question : Which one of the following statements about packing in solids is incorrect ?

a) Coordination number in bcc mode of packing is 8.

b) Coordination number in hcp mode of packing is 12.

c) Void space in hcp mode of packing is 32%.

d) Void space is ccp mode of packing is 26%.

#### Answer : C

#### Question : The number of atoms contained in a *fcc* unit cell of a monoatomic substance is

a) 1

b) 2

c) 4

d) 6



#### Question : In face-centred cubic lattice, a unit cell is shared equally by how many unit cells

a) 2

#### b) 4

c) 6

#### d) 8

#### Answer : C

#### Question : The number of atoms per unit cell of bcc structure is

a) 1			
b) 2			
c) 4			
d) 6			

#### Answer : B

Question : When molten zinc is converted into solid state, it acquires *hcp* structure. The number of nearest neighbours of Zn will be

a) 6 b) 12 c) 8 d) 4

#### Answer : B

#### Question : Hexagonal close packed arrangement of ions is described as

a) ABC ABA

b) ABC ABC

c) ABABA

d) ABBAB EBD\_7207

#### Answer : C

#### Question : In which of the following crystals alternate tetrahedral voids are occupied?

- a) NaCl
- b) ZnS
- c) CaF<sub>2</sub>
- d) Na2O

#### Answer : B

Question : Which of the following metal(s) show(s) hexagonal close packed structure (hcp) and which show face centredvcubic (fcc) structure? hcp fcc

a) Ag, Zn Mg, Cu

b) Mg, Zn Ag, Cu

c) Cu, Fe Al, Sn

d) Na, Li Zn, Cu

Answer : B

### Question : The number of octahedral voids present in a lattice is A . The number of closed packed articles, the number of tetrahedral voids generated is B the number of closed packed particles

a) A- equal, B- half

b) A- twice, B- equal

- c) A- twice , B- half
- d) A- equal, B- twice

#### **Answer : A**

### Question : In the hexagonal close packed structure of a metallic lattice, the number of nearest neighbours of a metallic atom i

a) twelve

b) four

- c) eight
- d) six

#### Answer : A

### Question : The Ca2+ and F- are located in CaF2 crystal, respectively at face centred cubic lattice points and in

- a) tetrahedral voids
- b) half of tetrahedral voids
- c) octahedral voids
- d) half of octahedral voids

#### Answer : A

### Question : If Germanium crystallises in the same way as diamond, thenmwhich of the following statement is not correct?

- a) Every atom in the structure is tetrahedrally bonded to 4 atoms.
- b) Unit cell consists of 8 Ge atoms and co-ordination number is 4.
- c) All the octahedral voids are occupied.
- d) All the octahedral voids and 50% tetrahedral voids remain unoccupied.

#### Answer : C

#### Question : The arrangement ABC ABC ..... is referred to as

a) Octahedral close packing

b) Hexagonal close packing

c) Tetrahedral close packing

**d**) Cubic close packing

Answer : D

#### Question : The total number of tetrahedral voids in the face centred unit cell is \_\_\_\_\_.

Answer : B			
d) 12			
c) 10			
b) 8			
a) 6			

#### Question : What is the coordination number in a square close packed structure in two dimensions ?

a) 2			
b) 3			
c) 4			
d) 6			

#### Answer : C

#### Question : In the cubic close packing, the unit cell has \_\_\_\_\_.

- a) 4 tetrahedral voids each of which is shared by four adjacent unit cells.
- b) 4 tetrahedral voids within the unit cell.
- c) 8 tetrahedral voids each of the which is shared by four adjacent unit cells.
- d) 8 tetrahedral voids within the unit cells.

#### **Answer : D**

#### Question : In which of the following arrangements octahedral voids are formed ?

(i) *hcp* (ii) *bcc* 

- (iii) simple cubic (iv) fcc
- a) (i), (ii)
- b) (i), (iv)
- c) (iii)

d) (ii), (iv)

#### Answer : B

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